Thermal decomposition of Solid Fuels
Fire Dynamics

Thomas Rogaume, Sophie Duquesne

and many persons

ESIA
Ecole des Sciences de l'Incendie et Applications

Thermal decomposition of solid fuels

- Introduction
- Thermal decomposition aspect – Problems in the condensed phase
- Thermal decomposition description
  - 1st phase: development of the pyrolysis model
  - 2nd phase: thermochemical and radiative properties
  - 3rd phase: heat feedback and oxygen diffusion
  - 4th phase: validation of the thermal decomposition model
- Fire retardancy of polymers
- Conclusion and perspective

Introduction

Global Introduction

Improvement and validation of the numerical tools

Experimental and numerical multi-step approach

Objective:

- ✓ To furnish the data required for the numerical tools development and validation
  - Initial and boundary conditions
  - Chemical, physical and thermal properties (as a function of temperature and environmental conditions)
  - Experimental data for the comparison with the numerical predictions and the validation of the model
- ✓ To define the laws of evolution of the fire (its development and the characteristics) and the key parameters as a function of time and/or temperature

Global Introduction

⇨ **Improvement and validation of the numerical tools**

Composition of a CFD model:

- ✓ Different sub-models, for example:
  - Heat transfer
  - Radiation
  - Combustion: EDC, EBU, flamelet, etc.
  - Turbulence
  - etc.
  - Pyrolysis model
- ✓ A correlation between those sub-models: FDS, Firefoam, etc.

Global Introduction

⇨ **Improvement and validation of the numerical tools**

Composition of a CFD model:

✓ Different sub-models, for example:

- Heat transfer
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- Combustion: EDC, EBU, flamelet, etc.
- Turbulence
- etc.
- **Pyrolysis model**

Our subject… (partially…)

✓ A correlation between those sub-models: FDS, Firefoam, etc.

Introduction – What is a pyrolysis model

<u>Model of pyrolysis:</u>

- ❑ Relation between temperature, atmosphere and kinetic of thermal decomposition of a material
- ❑ Limit: volume element d*V* wich is considered at each step homogeneous in term of temperature and of composition.

So, a **pyrolysis model permits to describe the chemical evolution (reactions) of a small volume element as a function of temperature and of the atmosphere (%O_2)**

Then, the pyrolysis model:

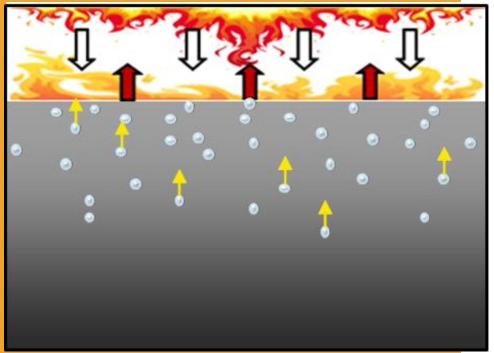
- Must be coupled to other ones to describe heat transfers (temperature), mass transfer (%O_2, devolatilization transfer, reactive mixture, etc.) and the boundary conditions of the volume element.
- It is **a part** of the description of thermal decomposition process.

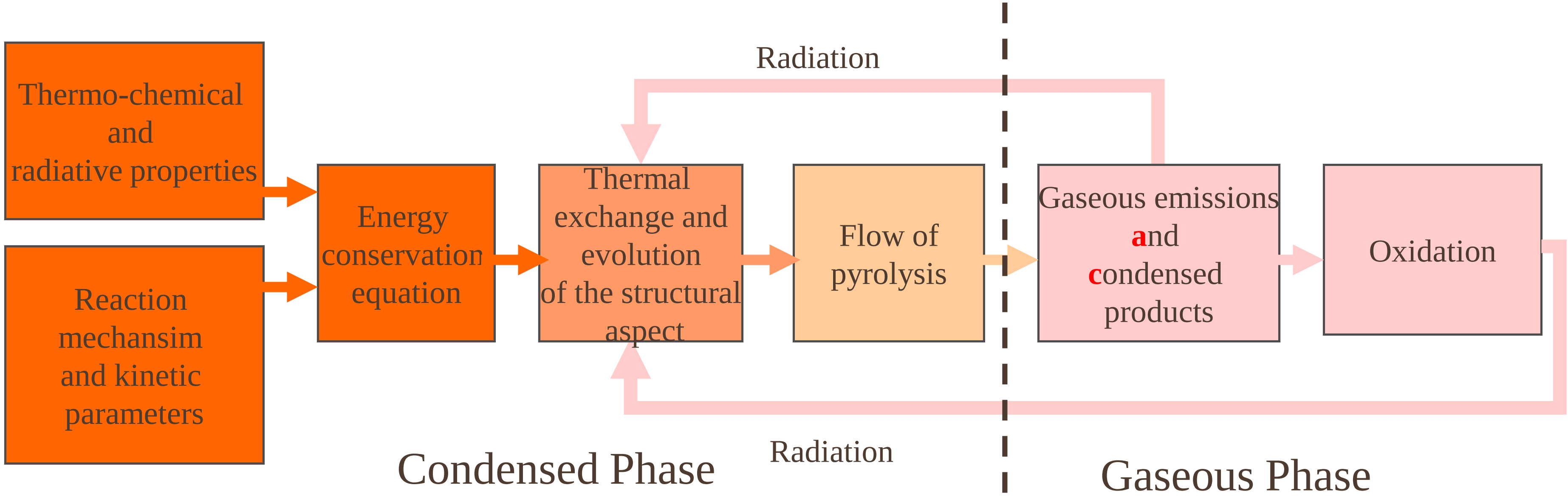
Introduction – Thermal decomposition

The description of the Thermal decomposition of a solid fuel reclaims:

- A model of pyrolysis
- The determination of physical, chemical and thermal properties
- Its coupling with other submodels

<u>Simplified representation</u>





--> equ. of mass conservation
--> equ. of movement quantity conservation (rate)
--> equ. of energy conservation (T)
--> equ. of species conservation (mass fractions Y_i)

Introduction – Thermal decomposition

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Our Challenge

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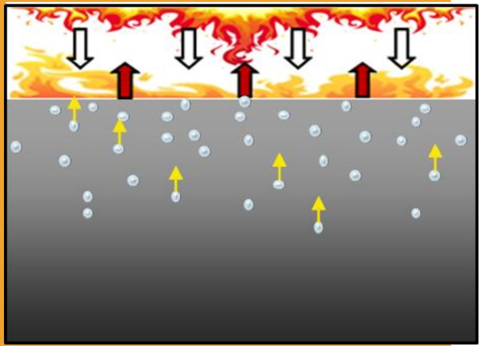

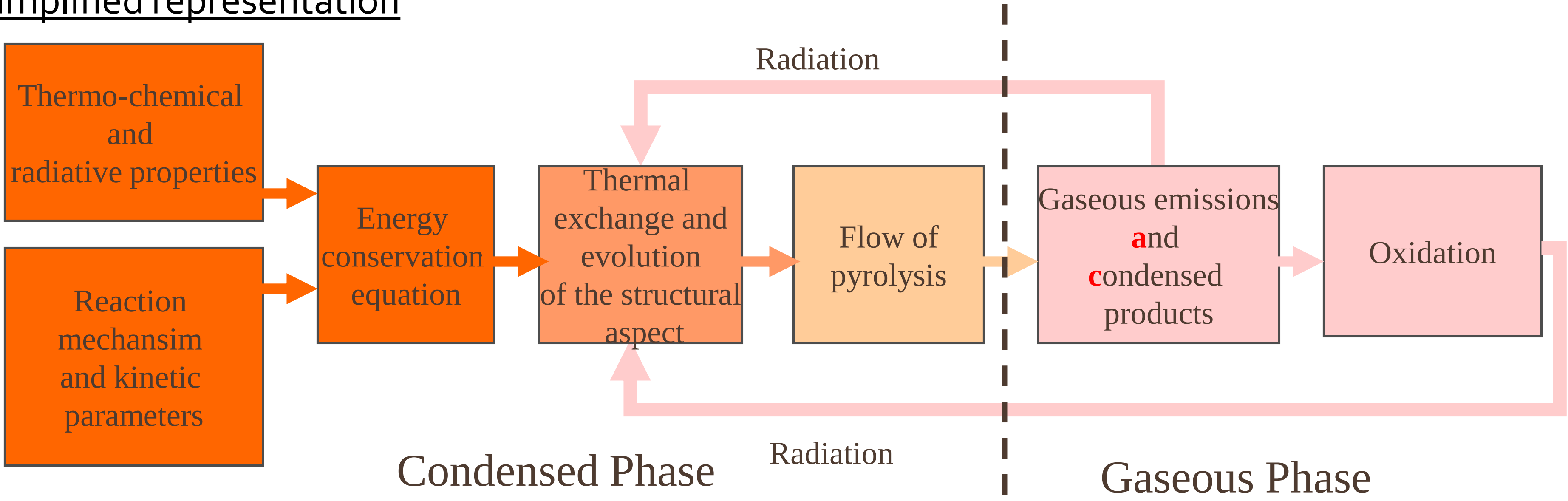

--> equ. of mass conservation
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Introduction – Challenge of the thermal decomposition

The modeling of thermal decomposition permits to describe:

- The mass loss and the mass loss rate of the solid
- So, the devolatilization flowrate: gaseous combustible products kept into the gas phase (quantity and composition)

→ **Source Term**

It influences:

- The ignition process (time)
- The flame structure (height, growing, etc.)
- The temperatures of the flame – Heat Release Rate
- The flame propagation
- Etc.

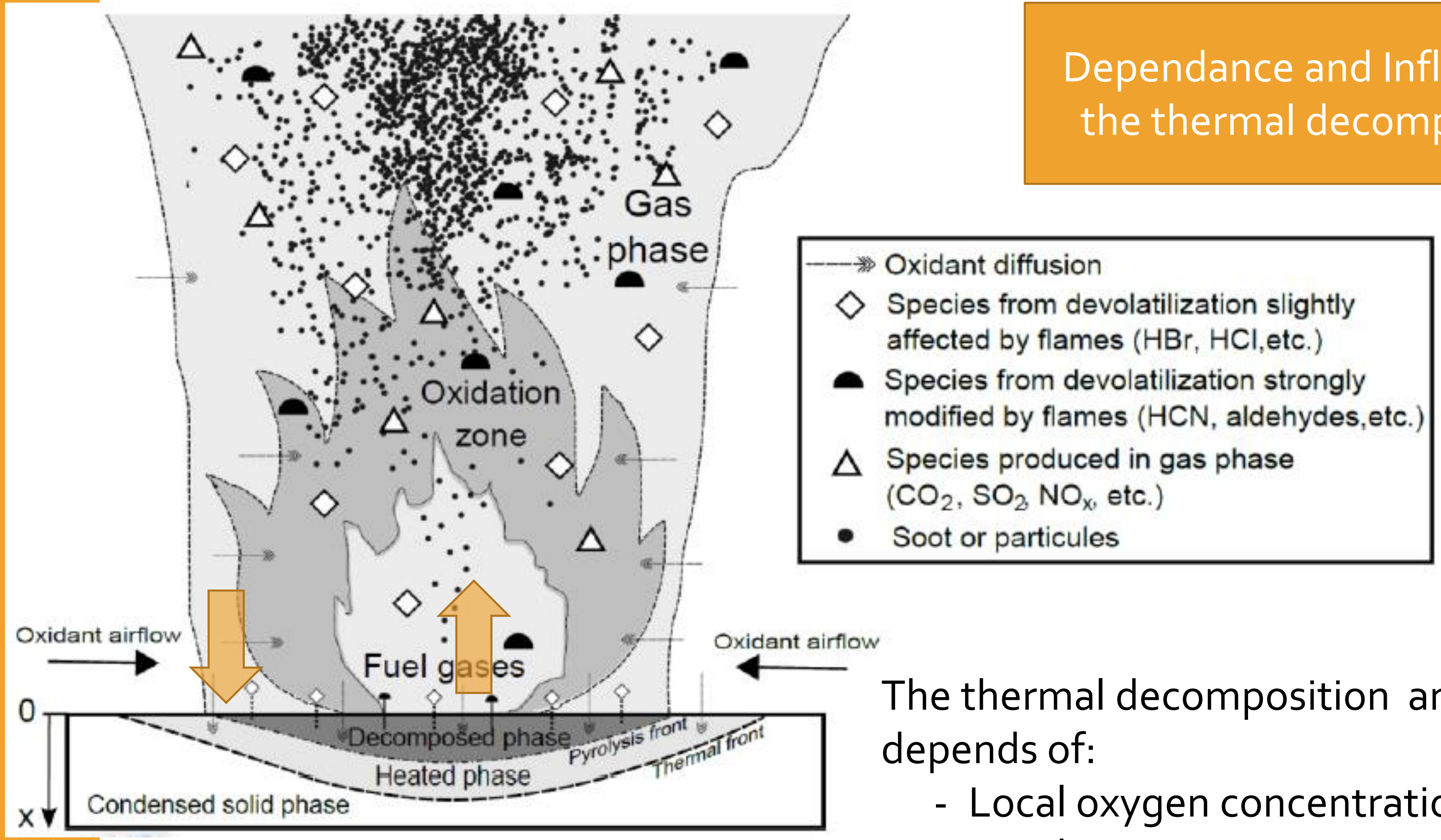
→ **So the Dynamics of the fire**

Are depending of the thermal decomposition

↓

Big challenge now to describe with accuracy the thermal decomposition of the materials involved into a fire

Introduction – Challenge of the thermal decomposition

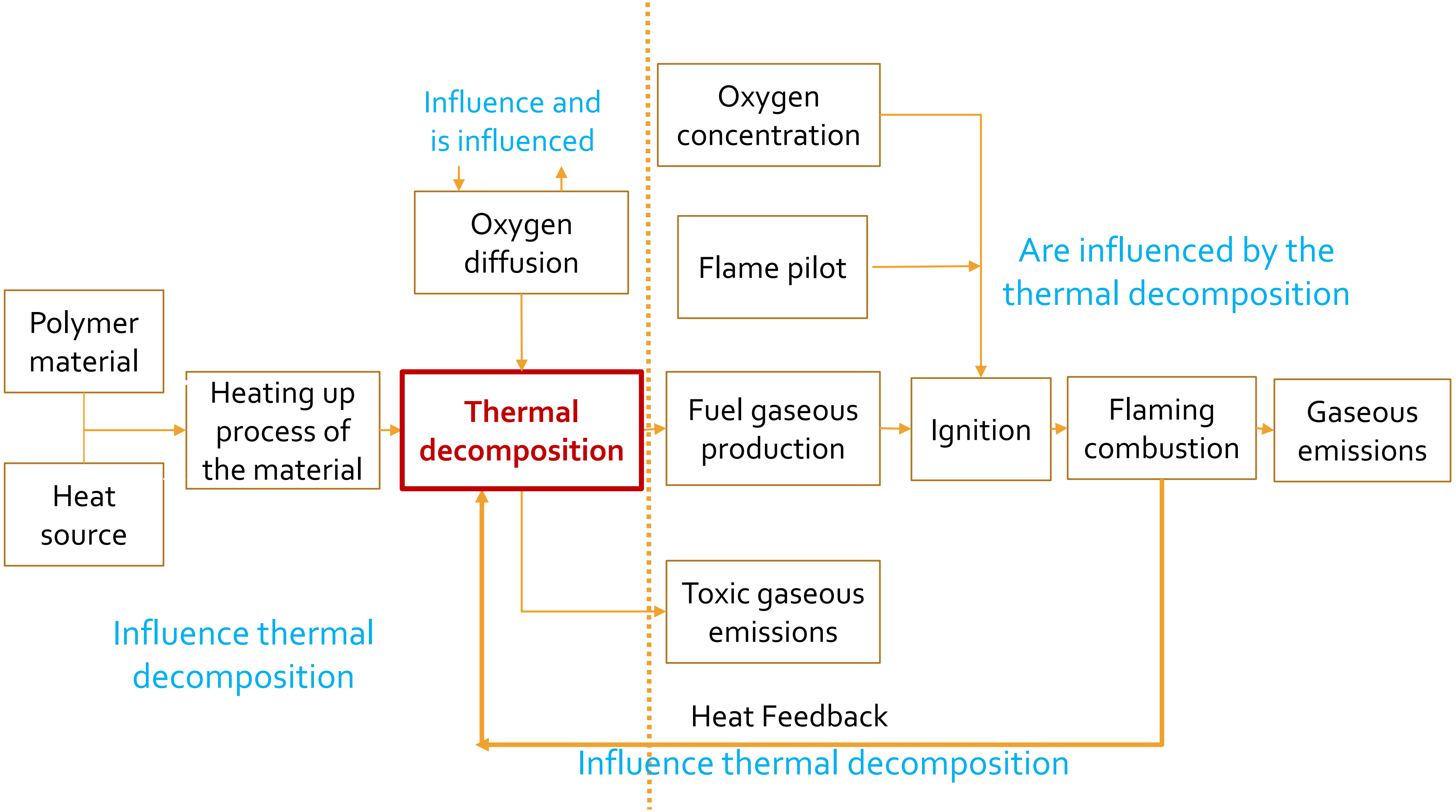


Dependance and Influence of the thermal decomposition

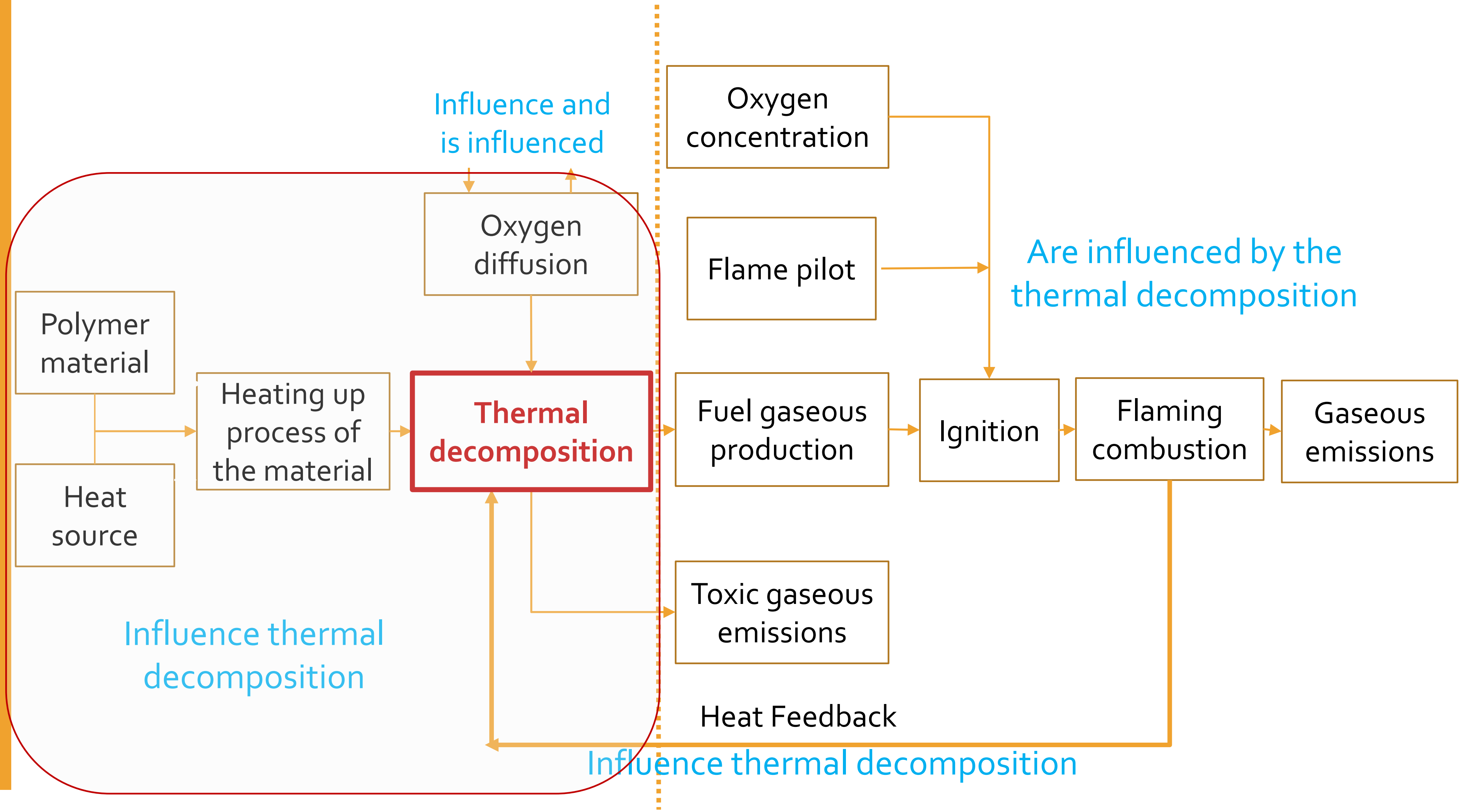
The thermal decomposition and gases production depends of:

- Local oxygen concentration
- Local temperature
- Structure and properties of the material (Intrinsic chemical and physical properties)

Introduction – Challenge of the thermal decomposition



Introduction – Challenge of the thermal decomposition



Challenge and complexity of thermal decomposition description

Thermal decomposition aspects

Nomenclature

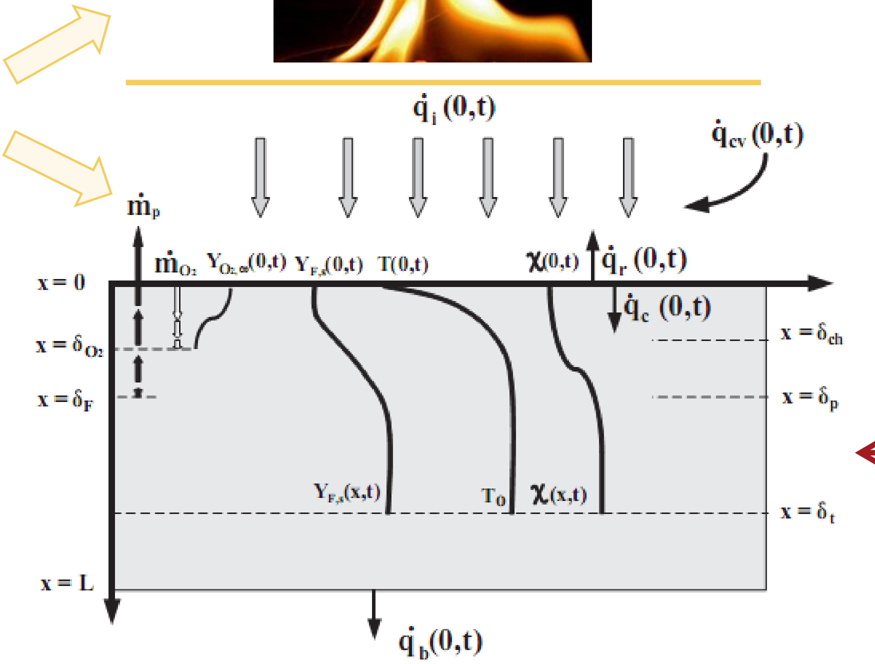
A	Pre-exponential factor	[s^{-1}]
E or Ea	Apparent activation energy	[$kJ \cdot mol^{-1}$]
E_{O_2}	Heat of combustion per unit mass of oxygen consumed (13.1 in this work)	[$MJ \cdot kgO_2^{-1}$]
EHC	Effective heat of combustion	[$kJ \cdot kg^{-1}$]
h_c	Convective heat transfer coefficient	[$W \cdot m^{-2} \cdot K^{-1}$]
HRR	Heat release rate per unit area	[$kW \cdot m^{-2}$]
k_s	Thermal conductivity	[$W \cdot m^{-1} \cdot K^{-1}$]
K_i	Solid mass fraction of the reaction i	[$g \cdot g^{-1}$]
m	Mass	[kg]
m_b	Mass-flow rate of species b	[$g \cdot s^{-1}$]
m''_b	Mass flux of species b	[$g \cdot s^{-1} \cdot m^{-2}$]
MLR	Mass Loss Rate	[$g \cdot s^{-1}$]
n	Reaction order	[-]
P	Pressure	[atm]
R	Universal constant of gases equal to 0.082	[$l \cdot atm \cdot mol^{-1} \cdot K^{-1}$]
$SMLR$	Specific mass-loss rate (per unit area)	[$g \cdot s^{-1} \cdot m^{-2}$]
T	Temperature	[°C] or [K]
ΔH	Enthalphy of the reaction	[$kJ \cdot kg^{-1}$]

Nomenclature

Greek symbols

α	Degree of conversion	[-]
β	Heating rate	[°C·min^{-1}]
Φ	Oxygen depletion factor	[-]
δ	Reaction order for oxygen mass fraction	[-]
λ	Wavelength	[m^{-1}]
ϕ	Fitness factor between curves	[-]
ρ	Density	[kg·m^{-3}]
υ_i	Stoichiometric coefficient of a solid or liquid product of reaction i	[-]
ω_i	Arrhenius reaction rate of reaction i	[s^{-1}]
ψ	Scale factor	[-]

Problem – Strong coupling between the condensed and the gas phases

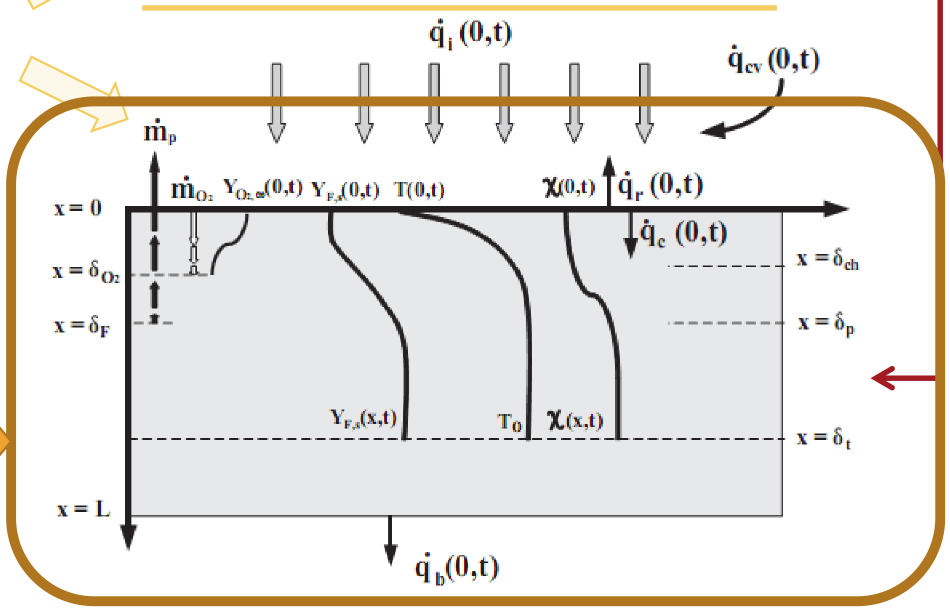


Simplified processus (1D) of thermal degradation of a solid - Torero.

Coupling
Description of the interactions

Problem – Actual approach, separation of the condensed and the gas phases

Objective of the present class



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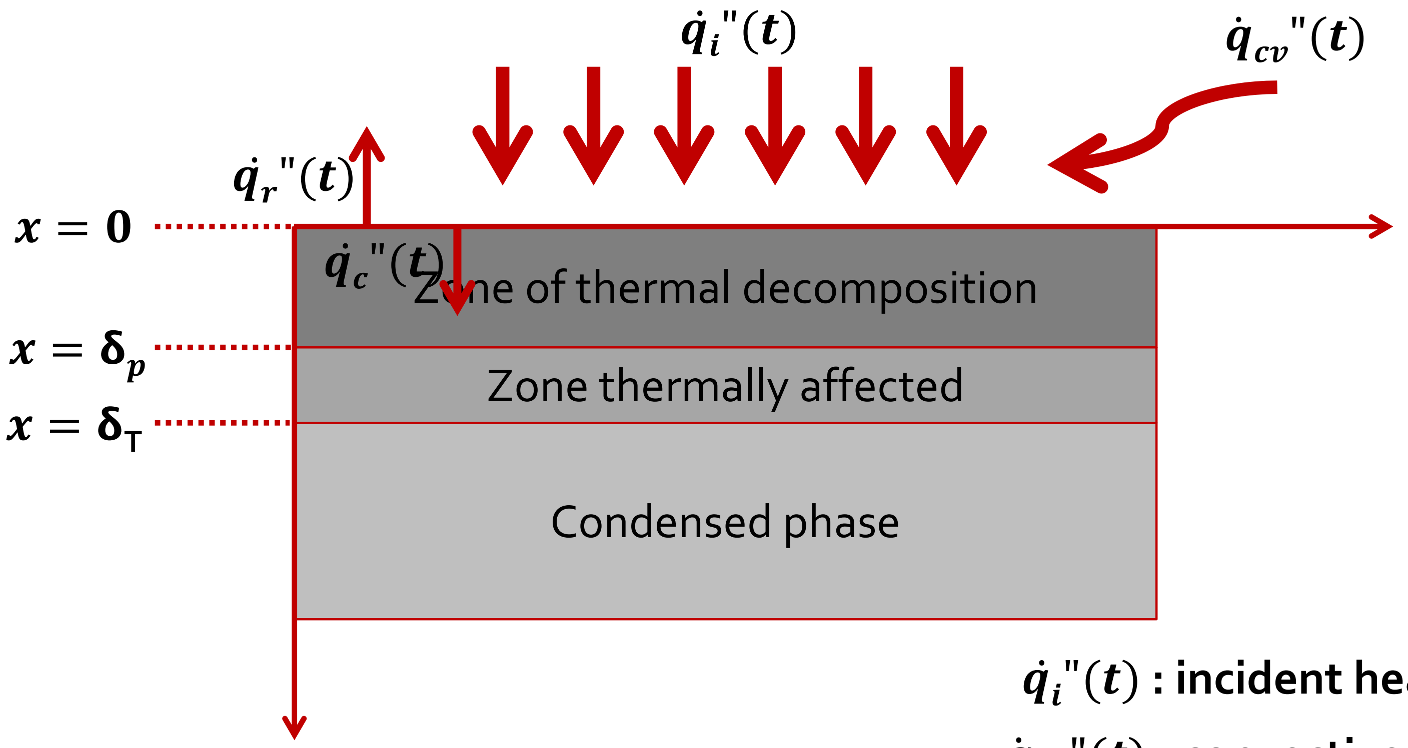
Problems into the Condensed phase



$x = \delta_p$: Position of the front of decomposition

$x = \delta_t$: Position of the heat front

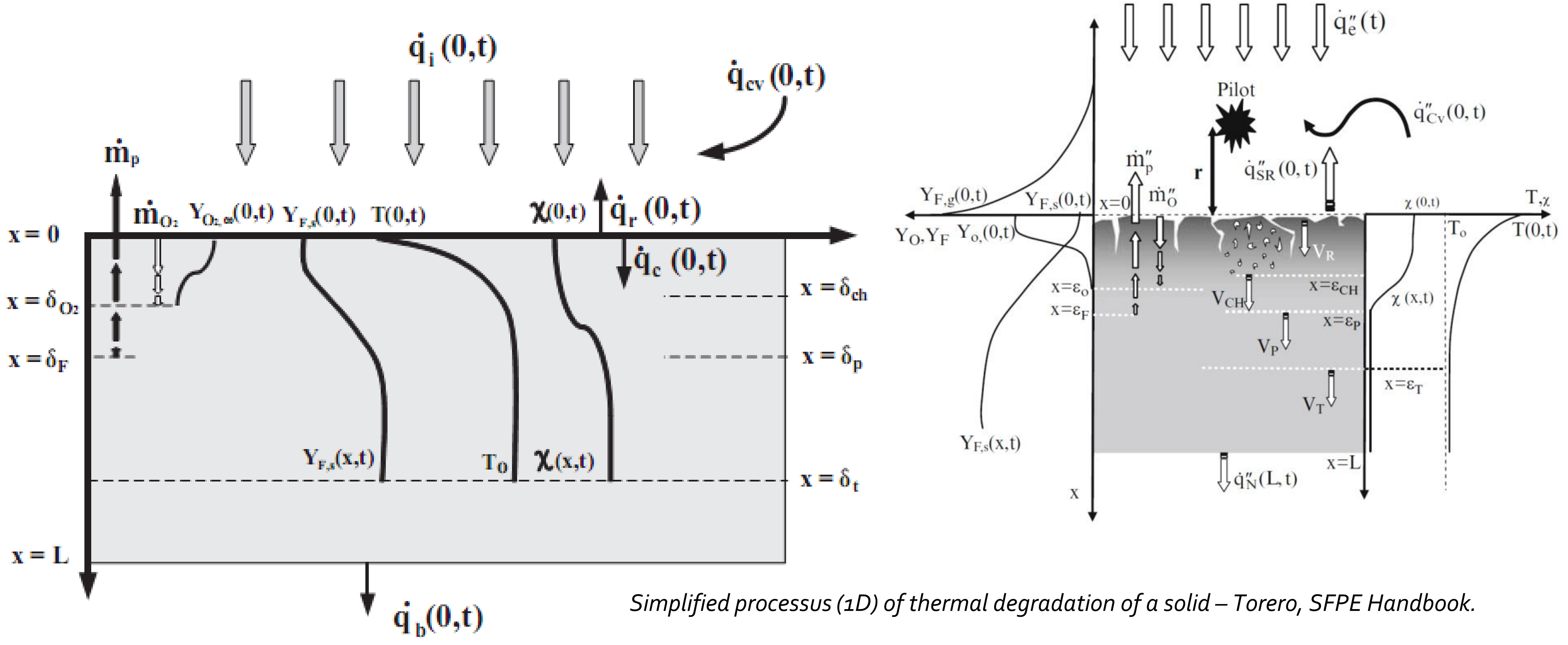
$\dot{q}_i''(t)$: incident heat flux

$\dot{q}_{cv}''(t)$: convective heat flux

$\dot{q}_c''(t)$: conductive heat flux

$\dot{q}_r''(t)$: re-radiative heat flux

Problems into the Condensed phase



Simplified processus (1D) of thermal degradation of a solid – Torero, SFPE Handbook.

- External radiation
- Thermal transfer into the solid
- Mass transfer (gaseous emissions and air)
- Kinetic reactions of thermal decomposition and of combustion of the solid fuel
- Char production…

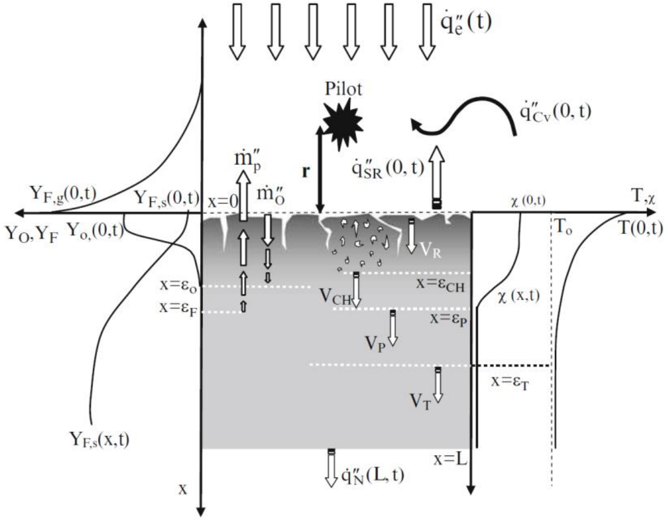
Energy conservation

$$\frac{\partial[\rho_S C_S T]}{\partial t} = \frac{\partial}{\partial x}\left[k_S \frac{\partial T}{\partial x}\right] + \frac{\partial[\dot{m}''_P C_{P,P} T_P]}{\partial x} - \frac{\partial[\dot{m}''_O C_{P,O} T_O]}{\partial x} + \frac{\partial[\rho_S V_R C_S T]}{\partial x} + \dot{q}'''_{RAD} + \sum_{i=1}^{i=N} \Delta H_{P,i}\rho_S\left[A_i Y_O^{m_i} Y_S^{n_i} e^{-E_i/RT}\right]$$

Chemical reactions

[J.L. Torero, SFPE Handbook]

Problems into the Condensed phase

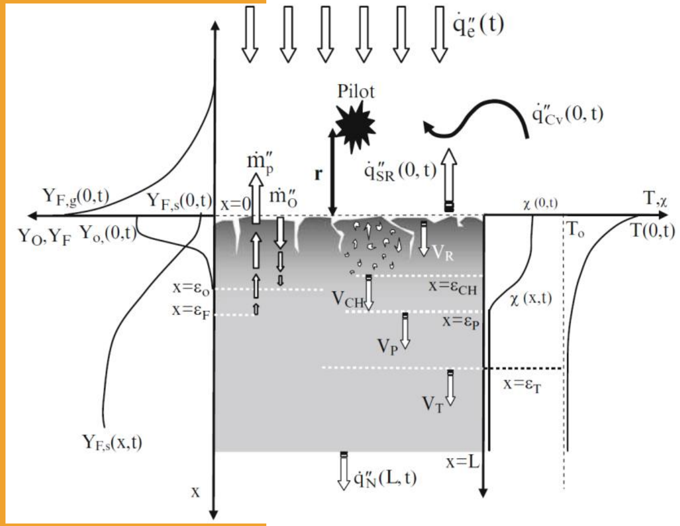


<u>The thermal degradation depends of - 1D approach - :</u>

- Temperature $T(x,t)$.
- Local mass fraction of combustible, $Y_s(x,t)$.
- Local mass fraction of oxygen, $Y_{O2}(x,t)$.
- Mass fraction of residual solid fuel, $Y_{F,s}(x,t)$.
- Permeability, $\chi(x,t)$.
- Thickness of oxygen diffusion into the solid, $\delta_{O2}(t)$.
- Thickness of the reactive zone, $\delta_F(t)$.
- Kinetic parameters values of each reaction, A_i, n_i, m_i, E_i.

[J.L. Torero, SFPE Handbook]

Problems into the Condensed phase



The mass loss rate is determined (sum of the i reactions):

- Locally by:

$$\dot{m}'''_p(x,t) = Y_{F,s}(\mathbf{x},t)\sum_{i=1}^{N}\left[A_i Y_{0_2}^{m_i}(\mathbf{x},t)\,Y_s^{n_i}(\mathbf{x},t)\,e^{-E_i/RT(\mathbf{x},t)}\right]$$

- For a surface unit, taking into account the thickness and so the permeability (volumic element dV) by:

$$\dot{m}''_p(\mathbf{x},t) = \int_0^L \chi(\mathbf{x},t)\left[Y_{F,s}(\mathbf{x},t)\sum_{i=1}^{N}\left(A_i Y_{0_2}^{m_i}(\mathbf{x},t)\,Y_s^{n_i}(\mathbf{x},t)\,e^{-E_i/RT(\mathbf{x},t)}\right)\right]dx$$

[J.L. Torero, SFPE Handbook]

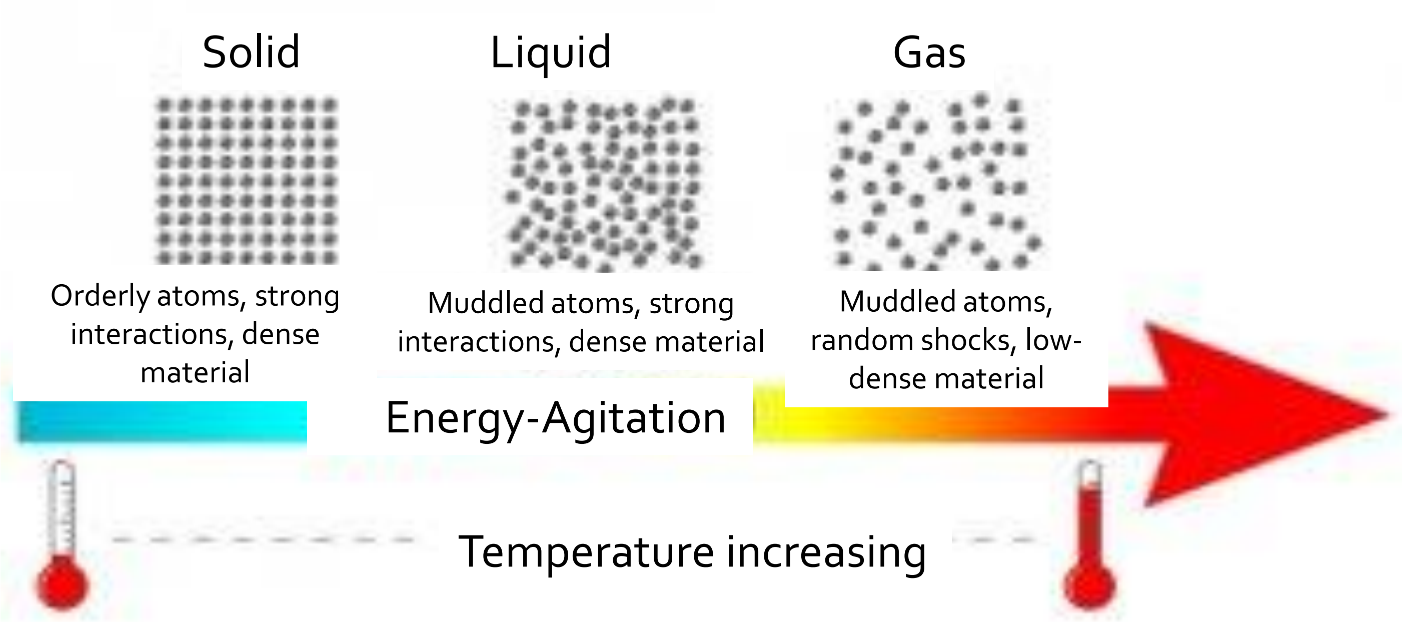
Problems into the Condensed phase

Thermal decomposition: transformation of the solid into gases by the breaking of the molecules into smaller ones – breaking of the chemical linkage

In a molecule, the atoms are linked the ones to the others by an energy = linkage energy

It is necessary to provide a sufficient energy in order to break this linkage

Energy of activation



Required parameters for the modeling of thermal decomposition
(Initial and boundary Conditions)

- Ambiant conditions: temperature, humidity, flows (rate), pressure…)

- Conditions of ventilation: Y_{O_2}

- Properties of the materials: physical, chemical, thermal (ρ, C_p, k, ε…) ; for each condensed phase

- Heat of each reaction and of combustion (ΔH_i)

- Kinetic model of thermal decomposition

- Kinetic parameters of each reaction: μ, A, Ea, n

- Heat flux received

- …

In space (x, y, z)

And

As a function of time (t) / temperature (T)

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(Initial and boundary Conditions)

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In space (x, y, z)

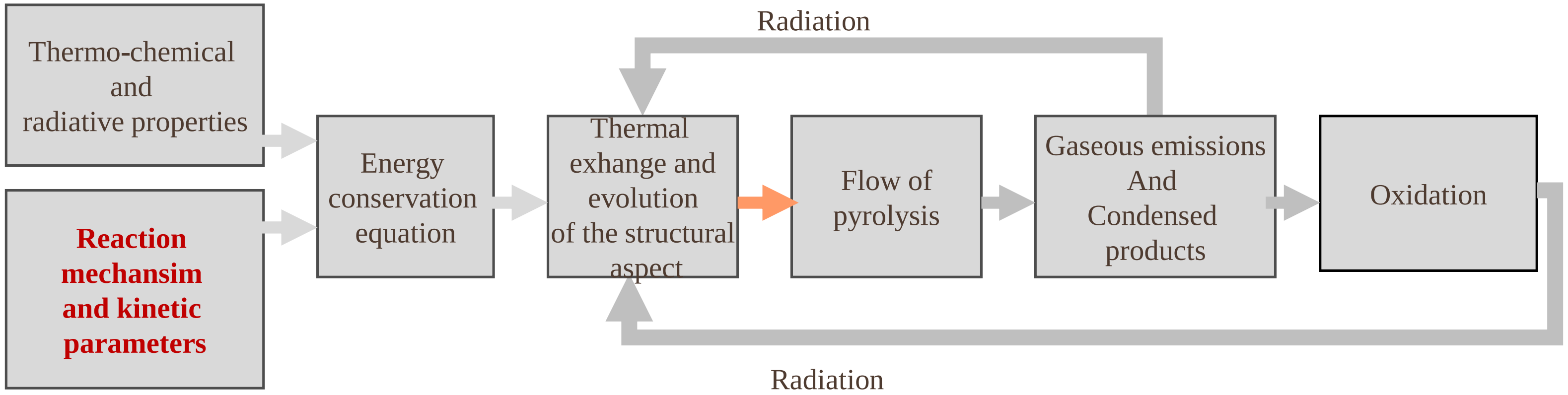
And

As a function of time (t)

How to determine Them ?

- Experimental investigations
- Inverse method of optimization
- Numerical approaches
- Different empirical laws

Thermal decomposition 1st phase: Development of the pyrolysis model



The pyrolysis model

Pyrolysis is a thermochemical decomposition process of organic material at elevated temperatures without the participation of oxygen. It involves the simultaneous change of chemical composition and physical phase and is irreversible. The word come from the Greek-derived word **pyr** « fire » and **lysis** « separating »

The different approaches to determine the model of pyrolysis:

- Modelistic approach – Model fitting method:
  - Uses a define reaction mechanism with an Arrhenius formulation
  - Requires the definition of A, E_a and n for each reaction. A, E_a and n are defined as properties of the reaction → use of optimization inverse methods
- Isoconvertionnal approach – Free model method:
  - Permits to determine the evolution of the activation energy as a function of the degree of conversion of the reaction. E_a is dependant of α and T.
  - Does not use a reaction mechanism (just one reaction) but is based on an Arrhenius form. The evolution of E_a permits to represent the MLR
  - Are available in the case of 1 reaction of thermal decomposition, or when the steps are clearly separated and chronologics (not parallel)
- Hybrid approach: combination of the modelistic and the isoconversionnal ones. Each peak of MLR is treated with a Kissinger method

The pyrolysis model

The different approaches to determine the model of pyrolysis:

Software	Modelistic approach	Isoconvertionnal approach
Component Kinetic		X
Thermokinetics	X	X
Thermo-Calc	X	X
Gpyro	X	
Thermakin	X	
FDS (V6)	X	

- **Special Focus Here** on the Modelistic approach – Model fitting method

Thermal decomposition
Development of the pyrolysis model

Special focus on the experimental benchscales

Focus: Experimental Investigations – Thermal decomposition

4 scales are classicaly used

Matter scale	mm mg	Thermal analysis: Thermo Gravimétric Analysis (TGA), Differencial Scanning Calorimetry (DSC), TDA, etc.
Small scale	cm g	Calorimetry: Cone Calorimeter (CC) and Fire Propagation Apparatus (FPA)
Product scale	10^n cm - m kg	IMO-LIFT, Medium Burning Item, Single Burning Item, room corner test
Real size scale	m kg - t	Rooms, House, real objects (train waggon, plane…)

The objective is to determine real properties and models avalaible for each scale and conditions

Focus: Experimental Investigations – Thermal decomposition

Multi-scale appoach

Knowledge of the phenomenon and control of the conditions (simplification)

Building scale

Room scale

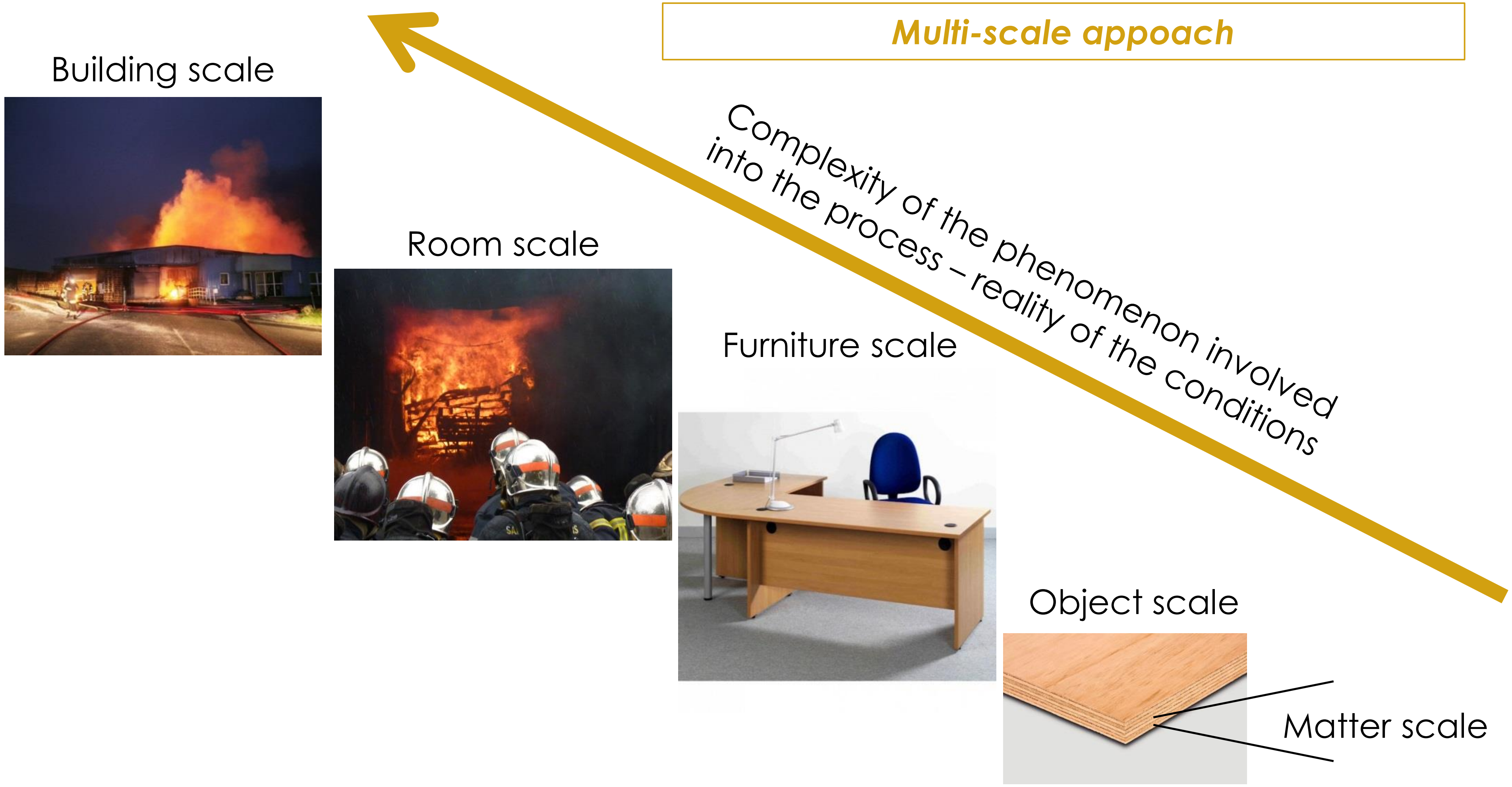
Furniture scale

Object scale

Matter scale

Focus: Experimental Investigations – Thermal decomposition

Multi-scale appoach



Focus: Experimental Investigations – Thermal decomposition

Matter Scale | Small Scale | Product scale | Real Scale

0D process

- No thermal diffusion
- No mass transfer
- « Perfect » control of the boundary conditions
- Controlled conditions: imposed by the user

- Intrinsic properties

1D process

- 1D Thermal transfer
- Radiative properties
- 1D mass transfer
- We try to control the boundary and the experimental conditions
- No flame transfer-propagation

- Simplified characterization of the diffusion aspects

2D or 3D process

- Thermal transfer
- Radiative properties
- Mass transfer
- Uncontrolled or difficulty to control the boundary conditions and the experimental ones
- Flame transfer-propagation

- Real diffusion aspects but… many process invoved and difficulty to determine the « conditions »

Focus: Experimental Investigations – Thermal decomposition

Matter Scale	Small Scale	Product scale	Real Scale

Matter Scale

- Mass Loss
- Mass loss rate

- Enthalpy of the reactions
- Heat of combustion
- Specific heat

- Model of pyrolysis (mechanisms and parameters)

Small Scale

- Delays of inflammation, extinction
- Mass loss
- Mass loss rate
- HRR
- Temperatures
- Gaseous emissions concentrations

- Thermal inertia
- Effusivity

Product scale / Real Scale

- Delays of inflammation, extinction, flash-over
- Mass loss
- Mass loss rate
- Gaseous emissions
- HRR
- Flame propagation
- Temperatures

Thermal decomposition
Development of the pyrolysis model

The pyrolysis model

Determination of a model of pyrolysis - Matter scale investigations

The modeling of the kinetic of thermal decomposition at matter scale

Choice of an approach **with an imposed model**

- Represents the detail kinetic
- Is applicable with complex mechanisms of kinetic of degradation
- Requires the definition of an homogeneous control volum

Difficulties of the method

- *Heterogeneity of the materials and multi-material fuels*
- *Thousands of kinetic reactions (ex. PE, more than 1500)*

Kinetic law

$$\dot{\omega} = A\, e^{-\frac{E_a}{RT}}\, m^n$$

→ Function of conversion ($f(\alpha)$)

↓ Law of variation of the rate ($k(T)$)

Reaction mechanism

$$+ \quad m_1 \xrightarrow{\dot{\omega},\nu} m_2 \Longrightarrow$$

Kinetic model – Model of Pyrolysis

Determination of a model of pyrolysis - Matter scale investigations

Considered here

The **model-fitting (modelistic) method** consists in selecting from a list of models the one that best fits TGA non-isothermal experimental curves.

The model-fitting method is expressed in terms of the degree of conversion: equal to 0 at the beginning of the test and to 1 when all the mass has been decomposed.

The degree of conversion is defined as:

$$\alpha = \frac{m_0 - m_t}{m_0 - m_f}$$

Where, m_o is the mass of the sample at the beginning of the process, m_t is the mass of the sample at an arbitrary time, m_f is the mass of the sample at the end of the process.

Determination of a model of pyrolysis - Matter scale investigations

Solid reaction rate = [k(T), the rate constant + f(α) the differential conversion function]:

- The rate constant is the Arrhenius equation

$$\frac{d\alpha}{dt} = k(T)\,f(\alpha) = A\exp\left(-\frac{E}{RT}\right)f(\alpha)$$

With A pre-exponential factor, E apparent activation energy, R universal gas constant, T absolute temperature

- The conversion function is the reaction model.

Determination of a model of pyrolysis - Matter scale investigations

The modeling of the kinetic of thermal decomposition at matter scale

Kinetic law

$$\dot{\omega} = A\, e^{-\frac{E_a}{RT}}\, m^n$$

Law of variation of the rate ($k(T)$)

Function of conversion ($f(\alpha)$)

Reaction mechanism

$$+ \quad m_1 \xrightarrow{\dot{\omega},\nu} m_2 \Longrightarrow$$

Kinetic model – Model of Pyrolysis

The pyrolysis model is formed by:

- A kinetic mechanism of thermal decomposition: kinetic reactions
- A kinetic model: rate of the reactions descrition
- Kinetic parameters for each reactions considered

How to describe them ?
How to determine them ?

- Intrinsic properties
- Intrinsic kinetic reactions and parameters
- In known and controlled conditions

I work at the matter scale – In 3 steps

Determination of a model of pyrolysis - Matter scale investigations

The modeling of the kinetic of thermal decomposition at matter scale

2nd step

Kinetic law

$$\dot{\omega} = A\, e^{-\frac{E_a}{RT}}\, m^n$$

Law of variation of the rate ($k(T)$)

3rd step

Function of conversion ($f(\alpha)$)

+

1st step

Reaction mechanism

$$m_1 \xrightarrow{\dot{\omega},\, \nu} m_2$$

⟹ Kinetic model – Model of Pyrolysis

Focus: Experimental Investigations – Thermal decomposition

Matter Scale | Small Scale | Product scale | Real Scale

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Determination of a model of pyrolysis - Matter scale investigations

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Hypothesis of the analysis:

- The material is thermally thin
- Temperature and concentration around the sample are homogeneous
- Surface thermal decomposition
- No gaseous diffusion and mass transfers
- The sample does not affect the fluid flow (vector gas)
- No heat transfer: equilibium between the furnace and the particle
- Thermodynamical equilibrium between solid and gas phases
- No local pressure gradient
- Etc.

Matter Scale

Controled parameters

- Heating rate
- Temperature
- Atmosphere
- Diffusion and transfer

Unknown parameters

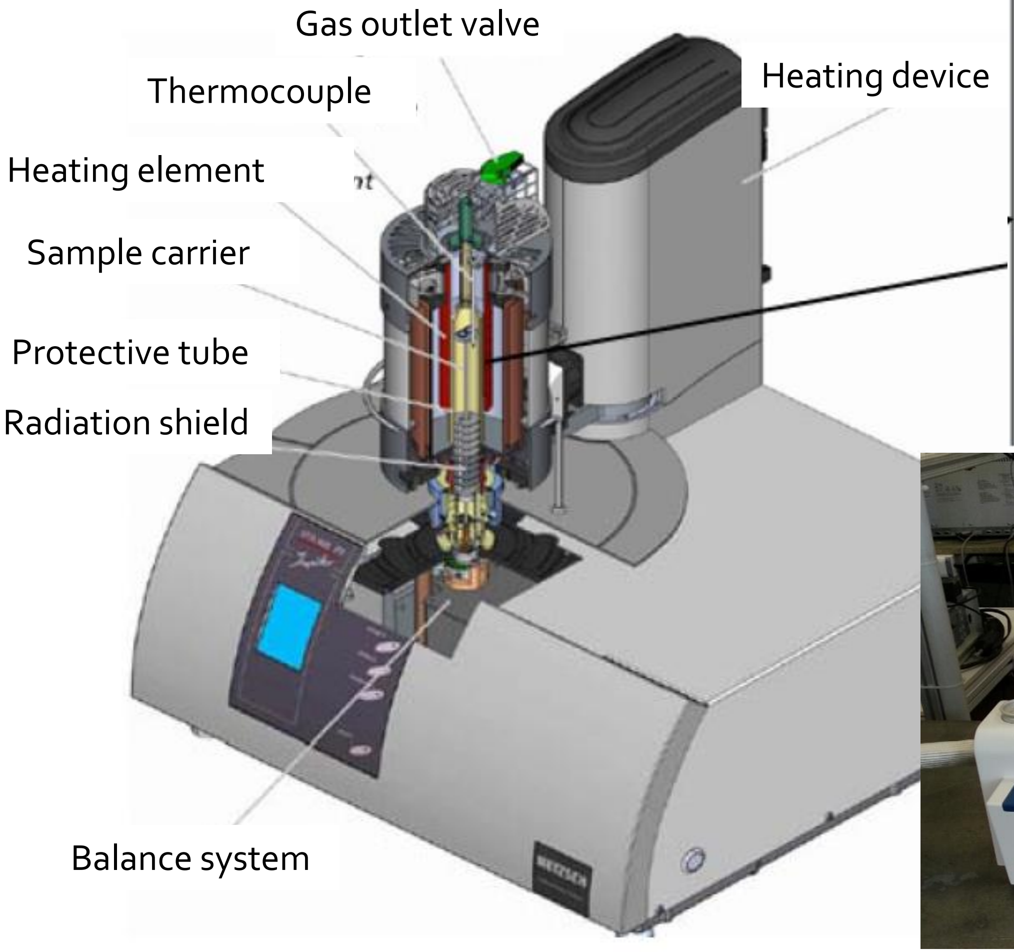
- /

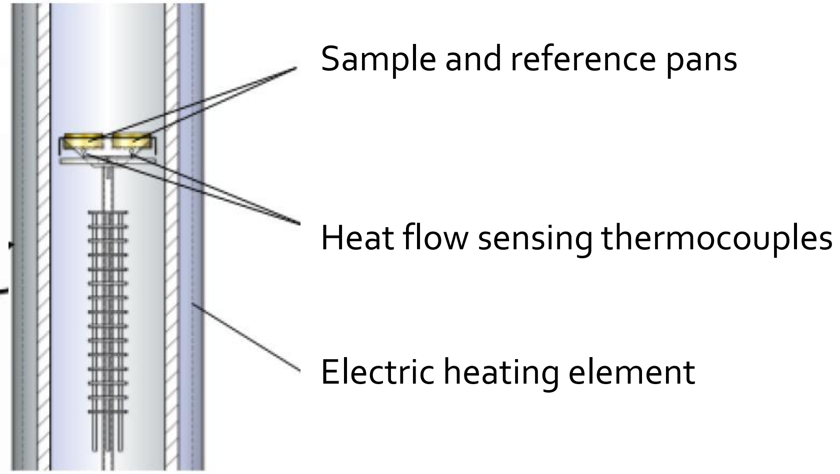
0D Investigation

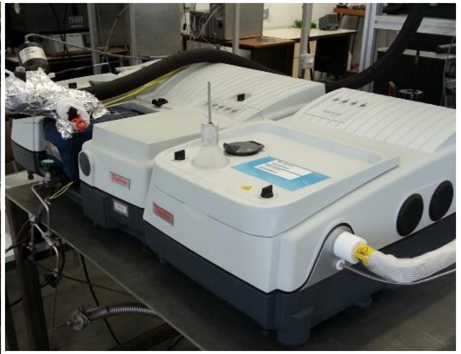
Determination of a model of pyrolysis - Matter scale investigations

1. A kinetic mechanism of thermal decomposition: kinetic reactions

TGA apparatus







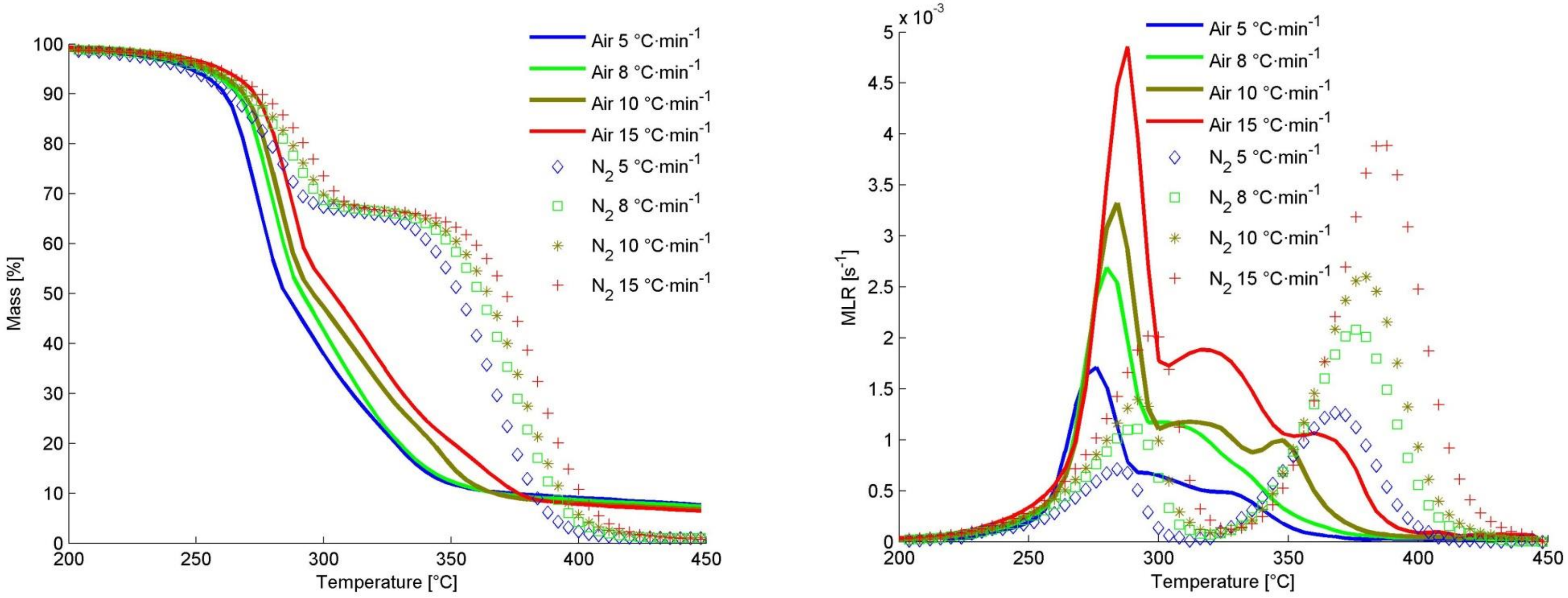
(TGA) – FTIR Apparatus

Determination of a model of pyrolysis - Matter scale investigations

TGA analysis:

- Examples of results

1. A kinetic mechanism of thermal decomposition: kinetic reactions



Thermal degradation of a PU foam in TGA

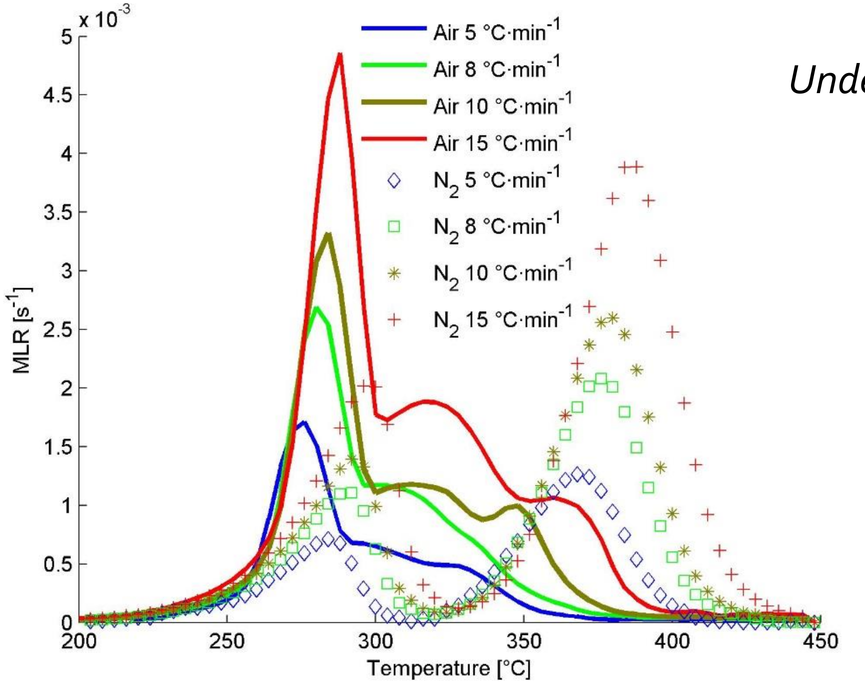
[*T. Rogaume & al. Development of the thermal decomposition mechanism of polyether polyurethane foam using both condensed and gas phase release data. Combustion Science and Technology, 2011*]

Determination of a model of pyrolysis - Matter scale investigations

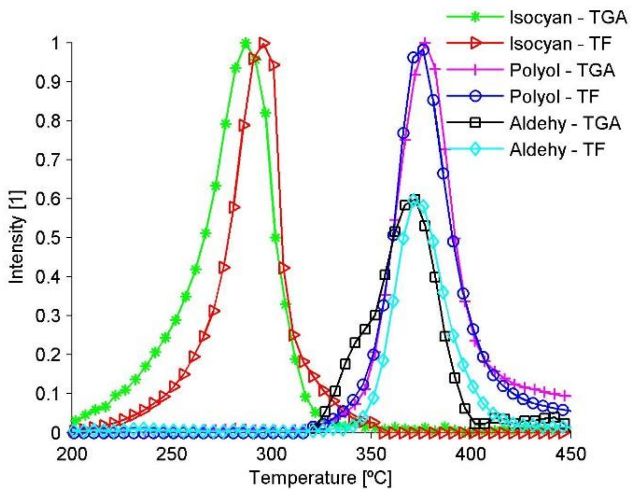
Results of TGA + $FTIR_{qlt}$

1. A kinetic mechanism of thermal decomposition: kinetic reactions

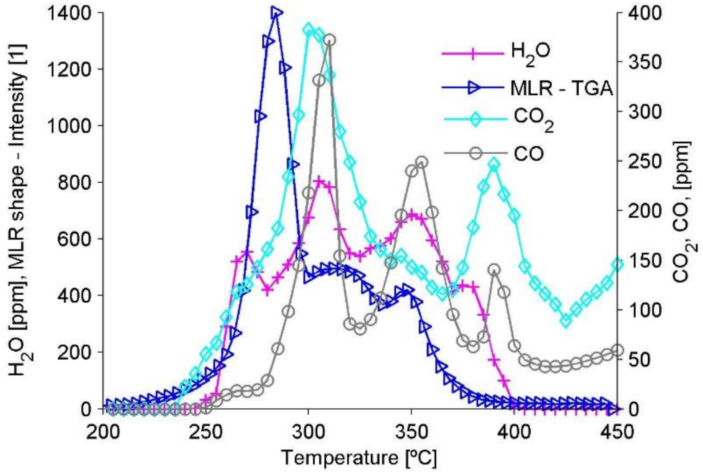
- Also possible with GC and MS system: knowledge of the gaseous emissions = tracker of what happen in the solid phase



Under inert atmosphere



Under air



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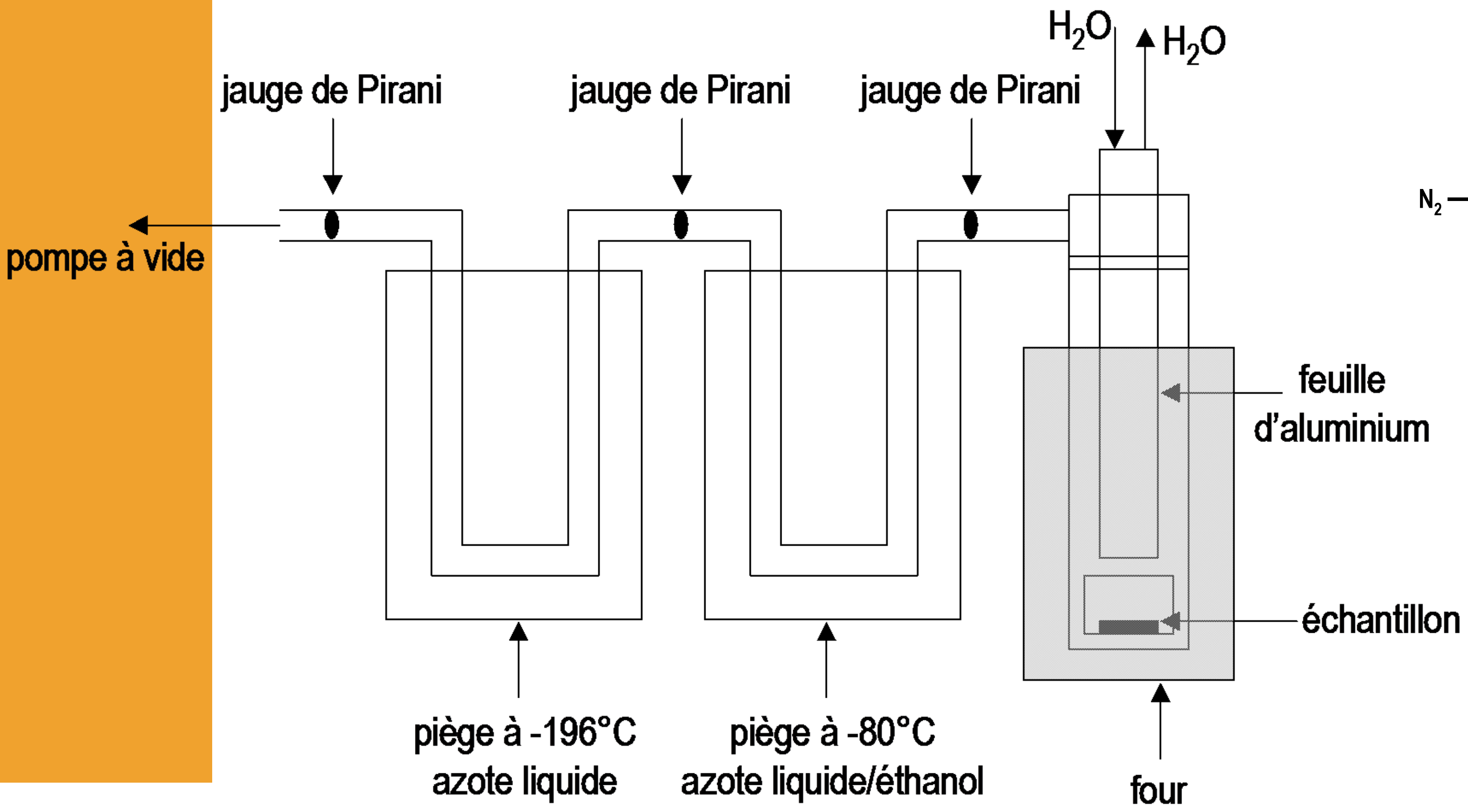
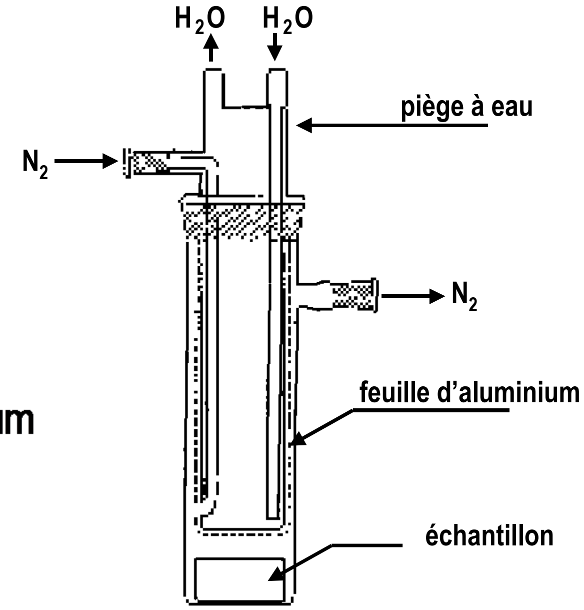
Determination of a model of pyrolysis - Matter scale investigations

Thermal Volatilisation Analyses (TVA)

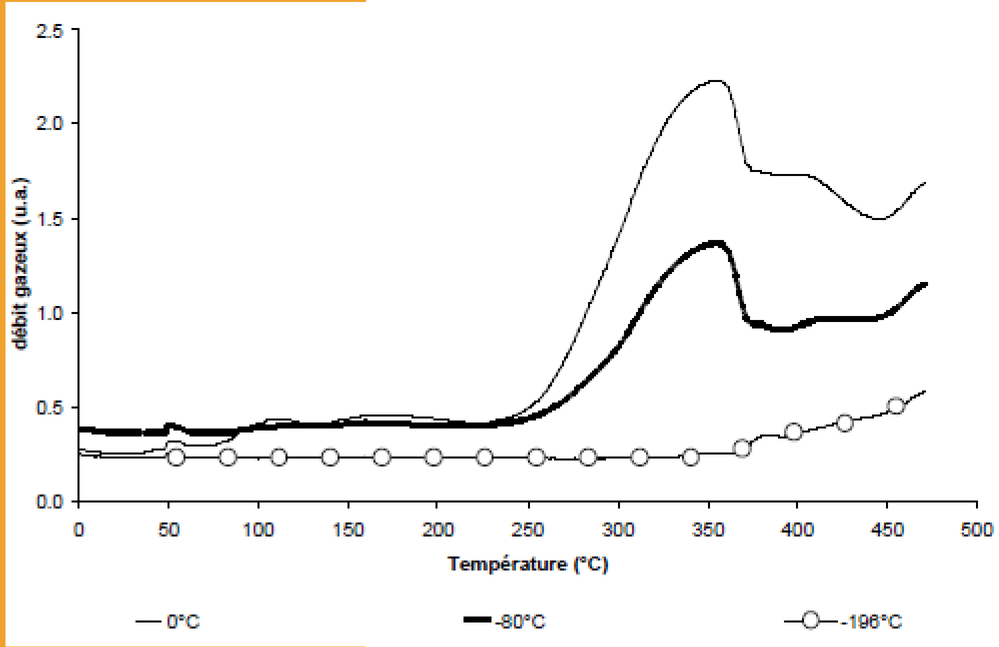


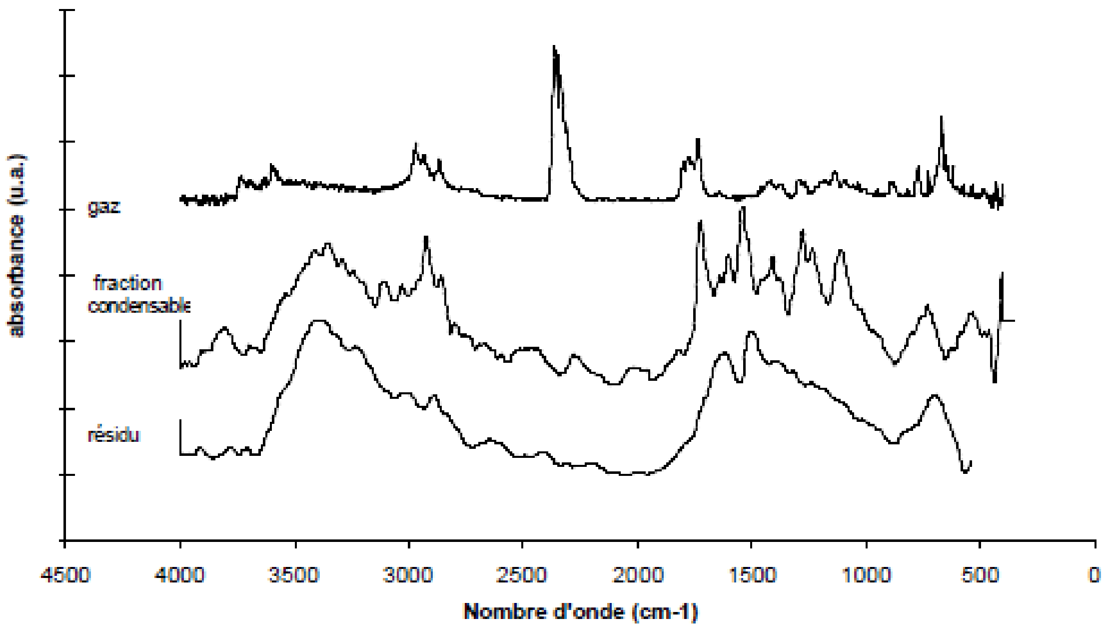


Determination of a model of pyrolysis - Matter scale investigations

TVA :

- Examples of results : PU



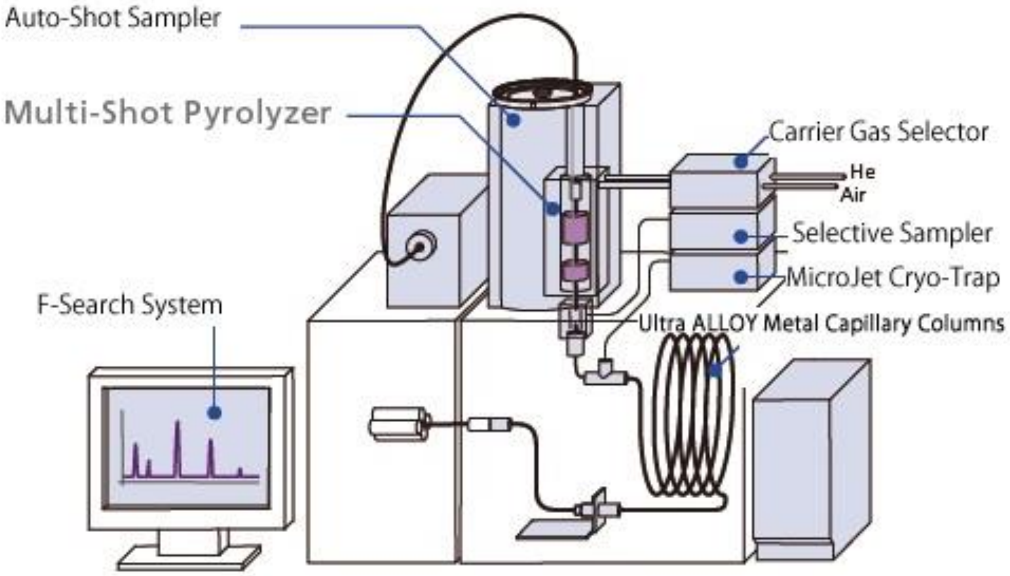


Spectre FTIR des différentes fractions collectées

Determination of a model of pyrolysis - Matter scale investigations

Pyrolyser – GC/MS

https://www.ssi.shimadzu



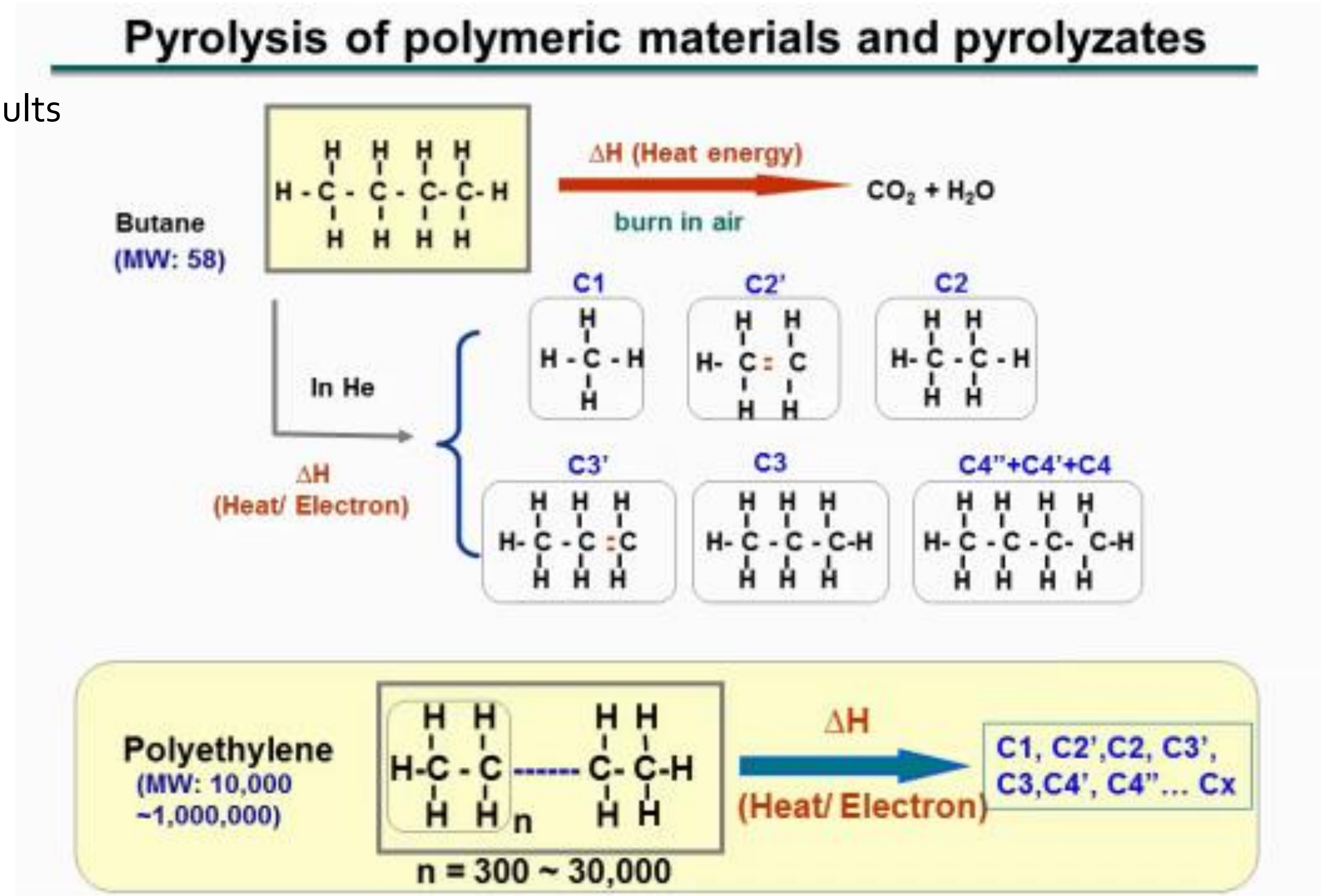
Py-GC/MS System (Multi-Shot Pyrolyzer and peripherals)

- Pyrolysis has many modes:
  - Thermal desorption (TD)
  - Flash pyrolysis (PY)
  - Temperature programming such as EGA

Determination of a model of pyrolysis - Matter scale investigations

Py-GC/MS:

- Examples of results

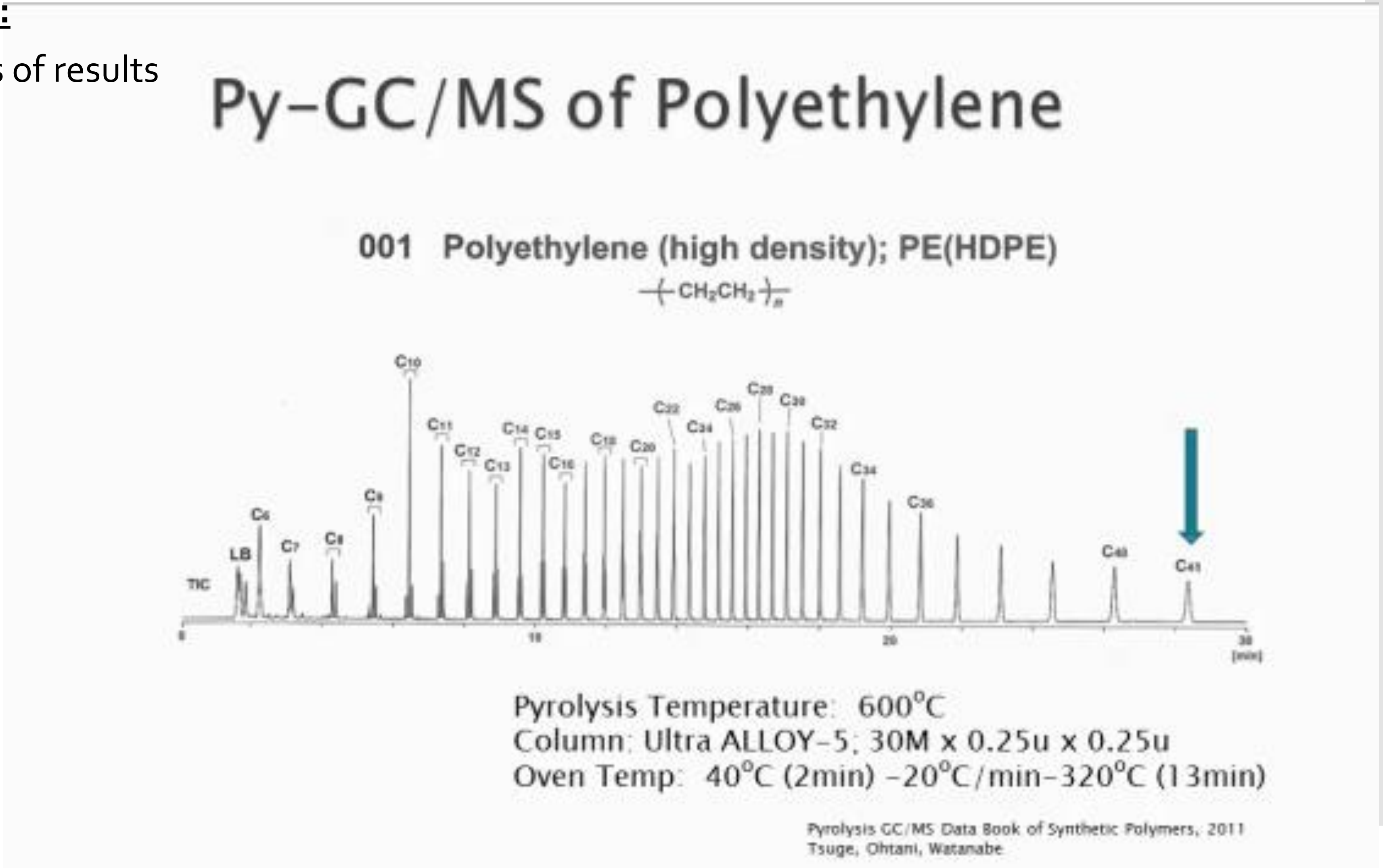


https://www.ssi.shimadzu.com

Determination of a model of pyrolysis - Matter scale investigations

Py-GC/MS:

- Examples of results



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Determination of a model of pyrolysis - Matter scale investigations

1. A kinetic mechanism of thermal decomposition: kinetic reactions

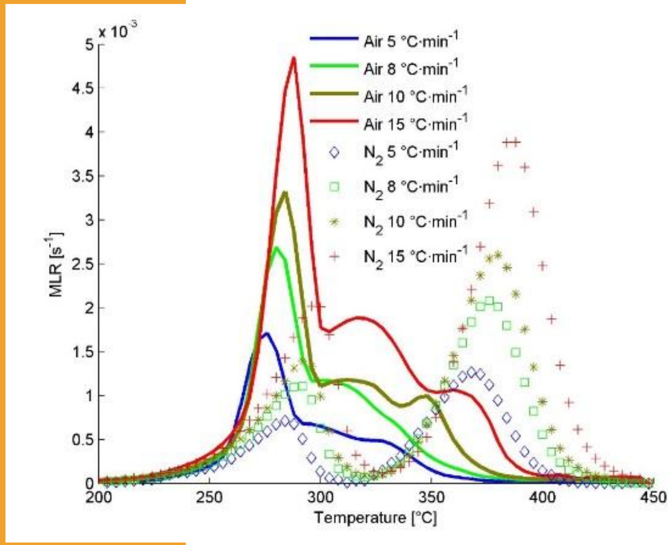
From the TGA, TGA-FTIR, TGA-GC analysis → reaction mechanism

Different approaches are possible:

- ***Lumped parameter approach (LPA):*** considers that the material is one, homogeneous. The thermal décomposition is the one of this material.
- ***The constituant approach***: considers that a material is composed of several constituants and that the thermal decomposition is the sum of the one of each component – *Example: wood is composed of cellulose, hemicellulose and lignin.*
- ***The functional approach:*** the thermal decomposition is described as functional groups, due to the well known fragmentation of the polymers – *Done for some « simple » plastic polymers.*

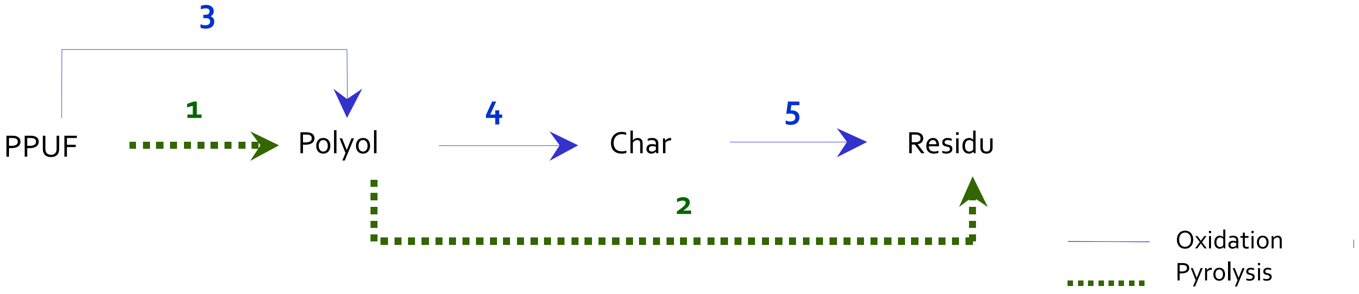
Determination of a model of pyrolysis - Matter scale investigations

Proposition of a reaction mechanism of thermal decomposition (LPA)



1. A kinetic mechanism of thermal decomposition: kinetic reactions

No	Type of reaction	Temp. [°C]	Reactives		Products solid or liquid		Products gas
1	Pyrolysis	200 – 340	PPUF	→	$\upsilon_1 \cdot$Polyol	+	$\tau_1 \cdot$[Isocyanate]
2	Pyrolysis	340 – 450	Polyol	→	$\upsilon_2 \cdot$Residue	+	$\tau_2 \cdot$[Polyol + H_2CO + H_2O + CH_4]
3	Oxidation	200 – 275	PPUF + O_2	→	$\upsilon_3 \cdot$Polyol	+	$\tau_3 \cdot$[Polyol + CO_2 + H_2O]
4	Oxidation	220 – 300	Polyol + O_2	→	$\upsilon_4 \cdot$Char	+	$\tau_4 \cdot$[Polyol + H_2CO + CH_4+ CO + CO_2 + H_2O]
5	Oxidation	300 – 450	Char + O_2	→	$\upsilon_5 \cdot$Residue	+	$\tau_5 \cdot$[Polyol + H_2CO + CH_4+ CO + CO_2 + H_2O]



[*T. Rogaume & al. Development of the thermal decomposition mechanism of polyether polyurethane foam using both condensed and gas phase release data. Combustion Science and Technology, 2011*]

Determination of a model of pyrolysis - Matter scale investigations

Rate of the reactions

2. A kinetic model: law of variation of the rate of the reactions

Steady rate in the gaseous phase

$$k_i(t) = A_i e^{\frac{-E_i}{RT(t)}}$$

k: Steady rate, s^{-1}
A: pré-exponentiel factor, s^{-1}
E: activation energy, $J.kg^{-1}$
R: Constant of the perfect gases

Arrhenius

Rate of variation of a concentration A, B and C

$$A \xrightarrow{k_1, \upsilon_1} B \xrightarrow{k_2, \upsilon_2} C$$

$$\upsilon_1 = k_1 \cdot [A] \qquad \upsilon_2 = k_2 \cdot [B]$$

$$-\frac{d[A]}{dt} = \upsilon_1 = k_1 \cdot [A]$$

$$\frac{d[B]}{dt} = \upsilon_1 - \upsilon_2 = k_1 \cdot [A] - k_2 \cdot [B]$$

$$-\frac{d[C]}{dt} = \upsilon_2 = k_2 \cdot [B]$$

[B. Batiot, Étude et modélisation de la cinétique de décomposition thermique des matériaux solides. Application à la dégradation du bois. Thèse de Doctorat ISAE-ENSMA, 2014]

Determination of a model of pyrolysis - Matter scale investigations

Rate of the reactions (solid phase)

2. A kinetic model: law of variation of the rate of the reactions

Rate of the reaction

$$\dot{\omega}_i = k Y_j^{n_i} Y_{O_2}$$

$$k_i(t) = A_i e^{\frac{-E_i}{RT(t)}}$$

$\dot{\omega}_i$ reaction rate (s^{-1}),
A : pré-exponential factor (s^{-1}),
E_i : activation energy, (J/kg),
n_i : order of the reaction
Y_j : mass fraction of a specie j into the reaction i,
T : temperature,
R perfect gas constant,
: mass fraction of oxygen into the reaction zone
δ : equal to 1 under oxidative atmosphere and to 0 under inert atmosphere

Mass loss rate of a specie j

Difference between the rate of production and of consumption of the specie j

$$\frac{d}{dt} Y_j = \sum_{\gamma \in H_j} Y_{r_\gamma} \dot{\omega}_i - \sum_{\xi \in G_j} \dot{\omega}_\xi$$

Y_r Residual mass fraction
H_j all the reactions producting j and G_j, the ones consuming j

Total mass loss rate

$$\frac{d}{dt} Y_t = \sum_{j=1}^{M} \frac{d}{dt} Y_j$$

[Rein G. *Computational model of forward and opposed smoldering combustion with improved chemical kinetics*. Thèse de doctorat, University of California, Berkley, 2005]

Determination of a model of pyrolysis - Matter scale investigations

2. A kinetic model: law of variation of the rate of the reactions

Mass balance of the species:

Total mass at time t, is the sum of remaining mass of each condensed phase species

$$MLR_i = \frac{dm_i}{dt} = \upsilon_i \cdot \omega_i$$

Each equation has an Arrhenius reaction rate

$$\dot{\omega}_i = A_i e^{-\frac{E_i}{RT}} \left(\frac{m_i}{m_o} \right)^{n_i} Y_{O_2}^{\delta}$$

Mass balance is expressed in terms of reaction rates and stoichiometric coefficients (v_i)

$$\frac{dm(t)}{dt} = \sum_{i=1}^{n} (v_i - 1)\dot{\omega}_i$$

Calculation results are compared to experiments

Determination of a model of pyrolysis - Matter scale investigations

Different conversion function

Model		f(α)
Reactional Order	ordre 0	1
	1er ordre	$1 - \alpha$
	2nd ordre	$(1 - \alpha)^2$
	ordre 3	$(1 - \alpha)^3$
	ordre *n*	$(1 - \alpha)^n$
Exponential law		α
Power function	$P = 3/2$	$\frac{3}{2}\alpha^{-1/3}$
	$P = 2$	$2\alpha^{1/2}$
	$P = 3$	$3\alpha^{2/3}$
	$P = 4$	$4\alpha^{3/4}$
	$P = n$	$n\alpha^{(1-1/n)}$

← Classically used (1er ordre, $1 - \alpha$)

Determination of a model of pyrolysis - Matter scale investigations

4. The determination of the kinetic parameters

Different conversion function

$$\frac{dm}{dt} = \sum_{b=1}^{i} MLR_b$$

$$\dot{\omega}_i = A_i e^{-\frac{E_i}{RT}} \left(\frac{m_i}{m_o} \right)^{n_i} Y_{O_2}^{\delta}$$

$$\frac{dm(t)}{dt} = \sum_{i=1}^{n} (v_i - 1)\dot{\omega}_i$$

k: Steady rate, s^{-1}
A: pré-exponentiel factor, s^{-1}
E: activation energy, $J.kg^{-1}$
R: Constant of the perfect gases
n: reaction order
υ: stoichiometric coefficient
δ=0 under inert, and =1 under air

Are Unknown parameters – How to determine them?

- Experimentally, it is not possible
- Then, we use inverse optimization methods: Genetic Algorithms, Particle Swarm Optimisation, Shuffled Complex Evolution, etc.

Determination of a model of pyrolysis - Matter scale investigations

Inverse methods of optimization

- Use an evaluation function
- This one defines a fitness ϕ which evaluates the adequation between the experimental and the numerical results of ML or/and MLR.
- Different evaluation functions:

$$\phi = \sum_{\beta=1}^{c}\left[\left(\int\left|\frac{dm}{dt}^{Calc} - \frac{dm}{dt}^{Exp}\right|dT\right)^{-1} + \psi\left(\int\left|m^{Calc} - m^{Exp}\right|dT\right)^{-1}\right]_{\beta}$$ [Rein & al.]

$$\phi = \sum_{\beta=1}^{c}\sum_{j=1}^{k}\left(\frac{dm_j}{dt}^{Calc} - \frac{dm_j}{dt}^{Exp}\right)_{\beta}^{-2}$$ [Esperanza & al.]

$$\phi = \sum_{\beta=1}^{c}\left[\cos\left(\widehat{(\vec{x},\vec{y})}\right).\left(\frac{\|\vec{x}-\vec{y}\|}{\|\vec{x}\|}\right)^{-1}\right]_{\beta} \qquad \vec{x} = \frac{dm^{exp}}{dT} \qquad \vec{y} = \frac{dm^{cal}}{dT}$$ [Bustamante Valencia & al.]

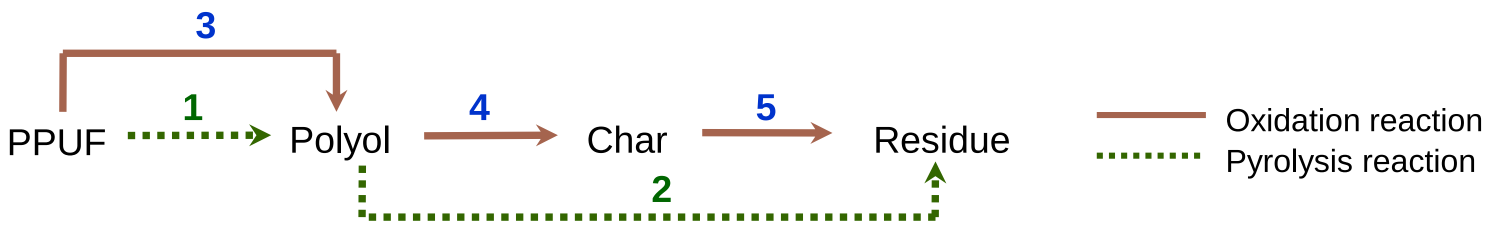
Determination of a model of pyrolysis - Matter scale investigations

One example of application: PU Foam

The mechanism includes the species of solid and gas phases

No	Type of reaction	Temp. [°C]	Reactives		Products solid or liquid		Products gas
1	Pyrolysis	200 – 340	PPUF	→	υ_1 ·Polyol	+	τ_1 ·[Isocyanate]
2	Pyrolysis	340 – 450	Polyol	→	υ_2 ·Residue	+	τ_2 ·[Polyol + H_2CO + H_2O + CH_4]
3	Oxidation	200 – 275	PPUF + O_2	→	υ_3 ·Polyol	+	τ_3 ·[Polyol + CO_2 + H_2O]
4	Oxidation	220 – 300	Polyol + O_2	→	υ_4 ·Char	+	τ_4 ·[Polyol + H_2CO + CH_4+ CO + CO_2 + H_2O]
5	Oxidation	300 – 450	Char + O_2	→	υ_5 ·Residue	+	τ_5 ·[Polyol + H_2CO + CH_4+ CO + CO_2 + H_2O]

The **"chemically correct"** kinetic mechanism allow prediction of MLR and gas releasing.



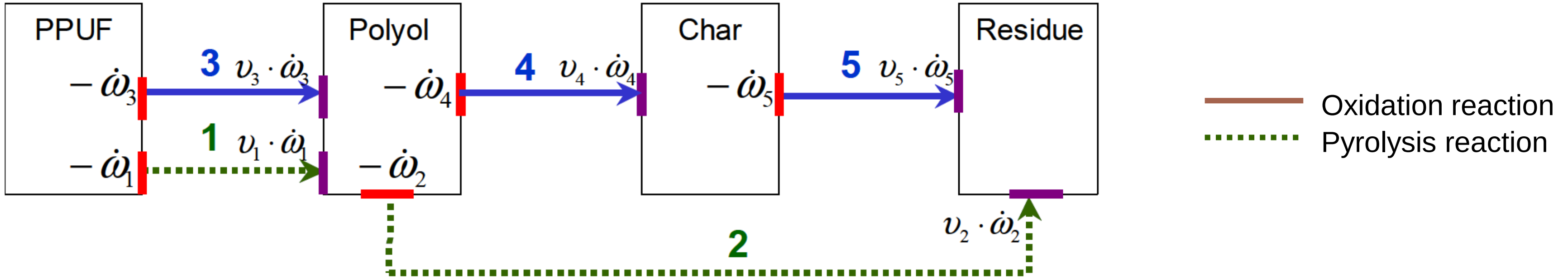
[Bustamante Valencia & al.]

Determination of a model of pyrolysis - Matter scale investigations

Mass loss description

One example of application: PU Foam

Mass balance of the species:



$$\frac{dm_{PPUF}}{dt} = -\dot{\omega}_1 - \dot{\omega}_3$$

$$\frac{dm_{Polyol}}{dt} = +\upsilon_1 \cdot \dot{\omega}_1 \; \upsilon_3 \cdot \dot{\omega}_3 - \dot{\omega}_2 - \dot{\omega}_4$$

$$\frac{dm_{Char}}{dt} = \upsilon_4 \cdot \dot{\omega}_4 - \dot{\omega}_5$$

$$\frac{dm_{Residue}}{dt} = \upsilon_2 \cdot \dot{\omega}_2 + \upsilon_5 \cdot \dot{\omega}_5$$

$$\dot{\omega}_i = A_i e^{-\frac{E_i}{RT}} \left(\frac{m_i}{m_o}\right)^{n_i} Y_{O_2}^{\delta}$$

$$\frac{dm}{dt} = \sum_{b=1}^{4} MLR_b = \frac{dm_{PPUF}}{dt} + \frac{dm_{Polyol}}{dt} + \frac{dm_{Char}}{dt} + \frac{dm_{Residue}}{dt}$$

$$\frac{dm}{dt} = (\upsilon_1 - 1)\dot{\omega}_1 + (\upsilon_2 - 1)\dot{\omega}_2 + (\upsilon_3 - 1)\dot{\omega}_3 + (\upsilon_4 - 1)\dot{\omega}_4 + (\upsilon_5 - 1)\dot{\omega}_5$$

[T. Rogaume & al. Development of the thermal decomposition mechanism of polyether polyurethane foam using both condensed and gas phase release data. Combustion Science and Technology, 2011]

Determination of a model of pyrolysis - Matter scale investigations

ESIA – Ecole des Sciences de l'Incendie et Applications – Obernai, 27 mai au 1er juin 2018

The unknown kinetic parameters determined are:

Calculation using Genetic Algorithms and a fitness function

One example of application: PU Foam

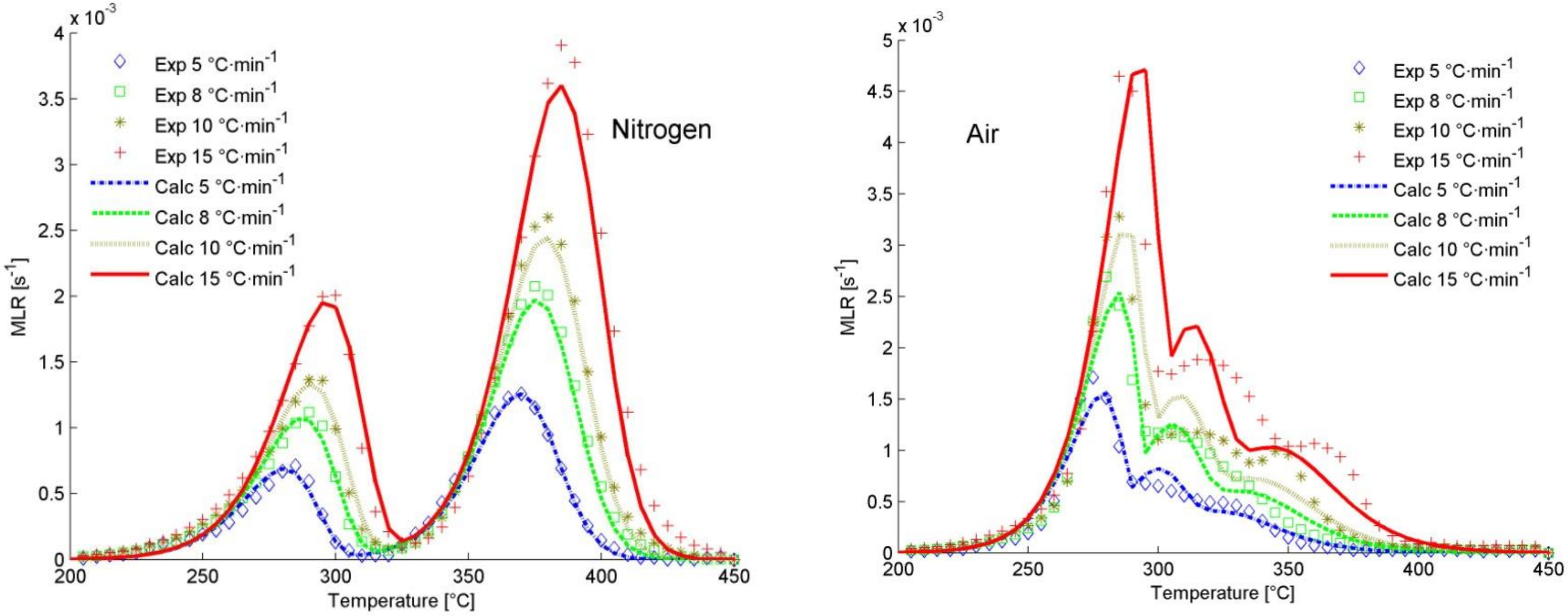
Reaction	Parameter	Value	Range High	Range Low	Units
PPUF pyrolysis	E_1	169.9	190	150	kJ·mol^{-1}
	A_1	6.09 x 10^{13}	1 x 10^{22}	1 x 10^{7}	s^{-1}
	$_1$	0.91	1	0.1	_
	υ_1	0.69	0.9	0.1	Kg·kg^{-1}
	τ_1	_	9 x 10^{9}	1.5 x 10^{9}	_
Polyol pyrolysis	E_2	243.9	260	100	kJ·mol^{-1}
	A_2	4.42 x 10^{17}	1 x 10^{19}	1 x 10^{7}	s^{-1}
	n_2	1.26	1.5	0.1	_
	υ_2	0.10	0.81	0.1	kg·kg^{-1}
	τ_2	4.9 x 10^{9}	9 x 10^{9}	1.5 x 10^{9}	_
PPUF oxidation	E_3	214.1	240	161	kJ·mol^{-1}
	A_3	3.07 x 10^{18}	1 x 10^{20}	1 x 10^{7}	s^{-1}
	n_3	0.48	3	0.2	_
	υ_3	0.44	0.7	0.1	kg·kg^{-1}
	τ_3	8.9 x 10^{4}	1.5 x 10^{5}	3 x 10^{4}	_
Polyol oxidation	E_4	213.6	240	161	kJ·mol^{-1}
	A_4	1.26 x 10^{18}	1 x 10^{22}	1 x 10^{7}	s^{-1}
	n_4	0.95	3	0.3	_
	υ_4	0.56	0.7	0.1	kg·kg^{-1}
	τ_4	8 x 10^{5}	2.2 x 10^{6}	2 x 10^{4}	_
Char oxidation	E_5	160.8	240	160	kJ·mol^{-1}
	A_5	4.30 x 10^{12}	3 x 10^{15}	1 x 10^{11}	s^{-1}
	n_5	1.64	3	0.5	_
	υ_5	0.25	0.8	0.1	kg·kg^{-1}
	τ_5	3.4 x 10^{6}	9 x 10^{6}	1.7 x 10^{5}	_

[*T. Rogaume & al. Development of the thermal decomposition mechanism of polyether polyurethane foam using both condensed and gas phase release data. Combustion Science and Technology, 2011*]

Determination of a model of pyrolysis - Matter scale investigations

One example of application: PU Foam

Validation of the pyrolysis model:



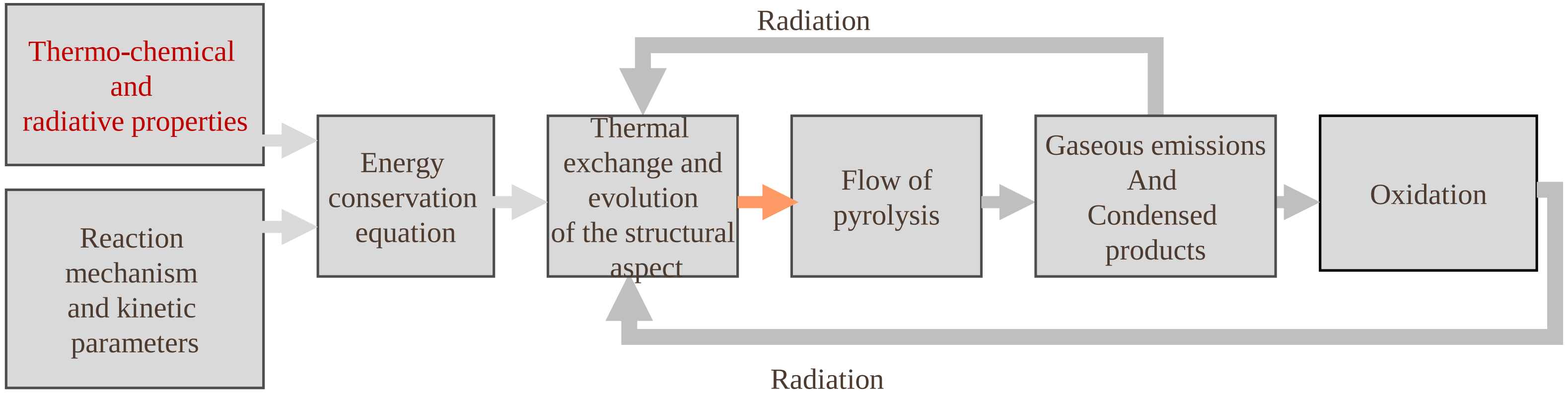
Comparison of numerical and experimental MLR curves in TGA

[*T. Rogaume & al. Development of the thermal decomposition mechanism of polyether polyurethane foam using both condensed and gas phase release data. Combustion Science and Technology, 2011*]

Validation of the pyrolysis model:

- The pyrolysis model has been developed from matter scale experiments
- The mechanism of thermal decomposition is proposed from TGA and gas analysis measurements
- The kinetic is described from a modified Arrhenius law and a conversion function
- The unknown kinetic parameters of each reaction are determined using inverse methods of optimization
- The objective is to fit, at matter scale, the experimental and numerical mass loss and MLR curves using a fitness function.

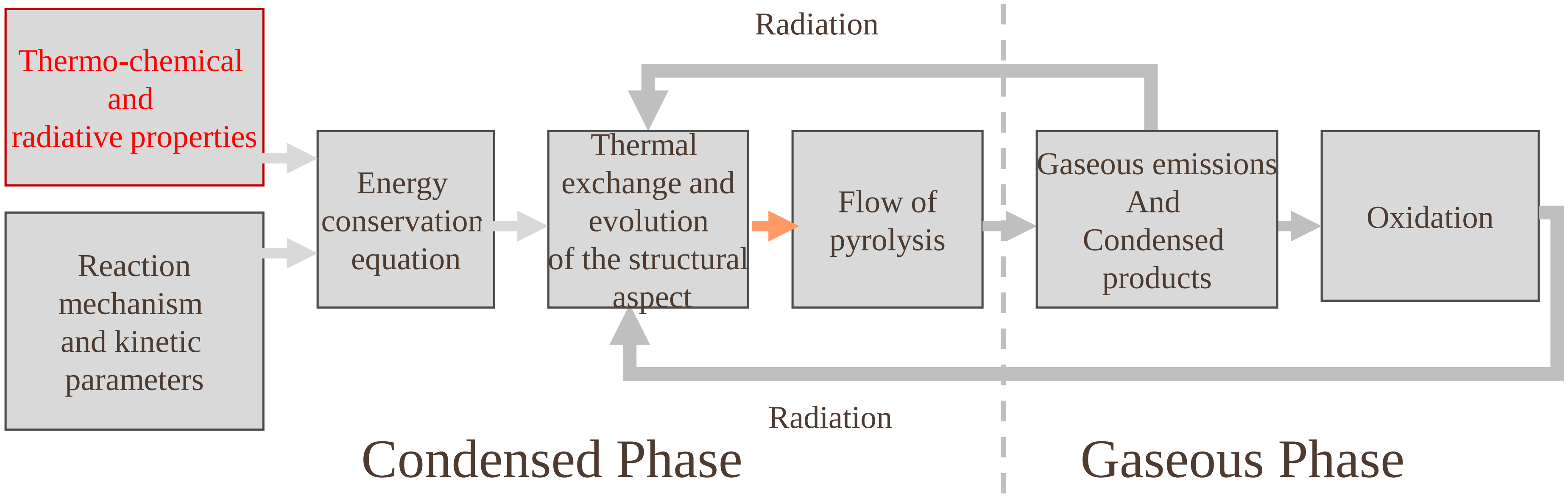
Thermal decomposition 2nd phase: Thermochemical and radiative properties



Determination of the thermochemical properties of the condensed phase

- k, ρ, Cp, ε…
- [C, H, O…] - Y_i
- Enthalpies of the reactions

- Experimental investigations
- Inverse optimization



Determination of the thermochemical properties of the condensed phase

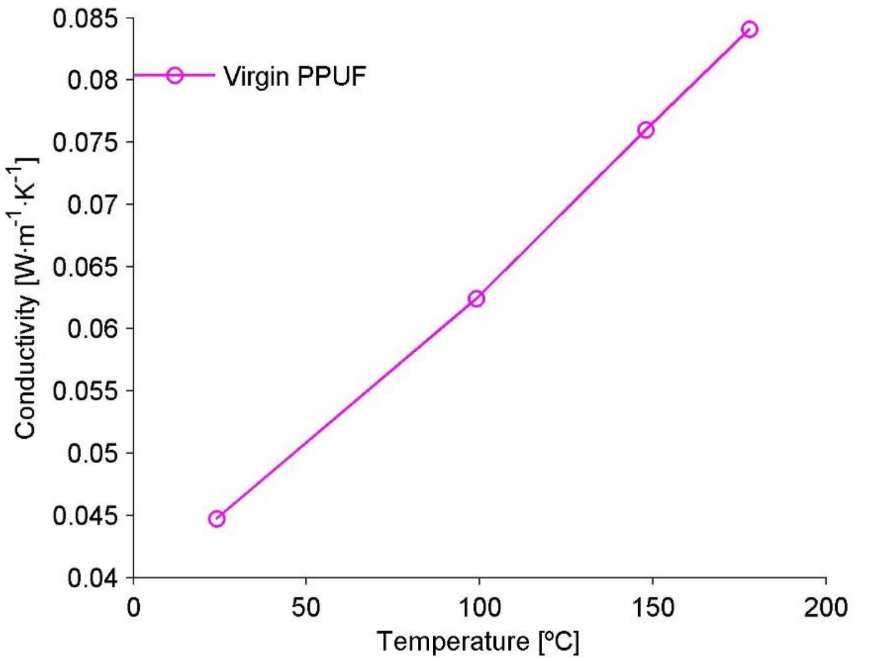
Determination of k – Different methodologies:
- Flash method (with a laser)
- Guarded hot plate

Superior calorific power : bomb calorimeter

Chemical properties:
- Elementary analysis: [C], [H], [O], [N], [S], [Cl],[H_2O], [ash], etc.
- Nuclear Magnetic Resonance: chemical linking.

Determination of Cp and ΔH
- Differencial scanning calorimetry (DSC)

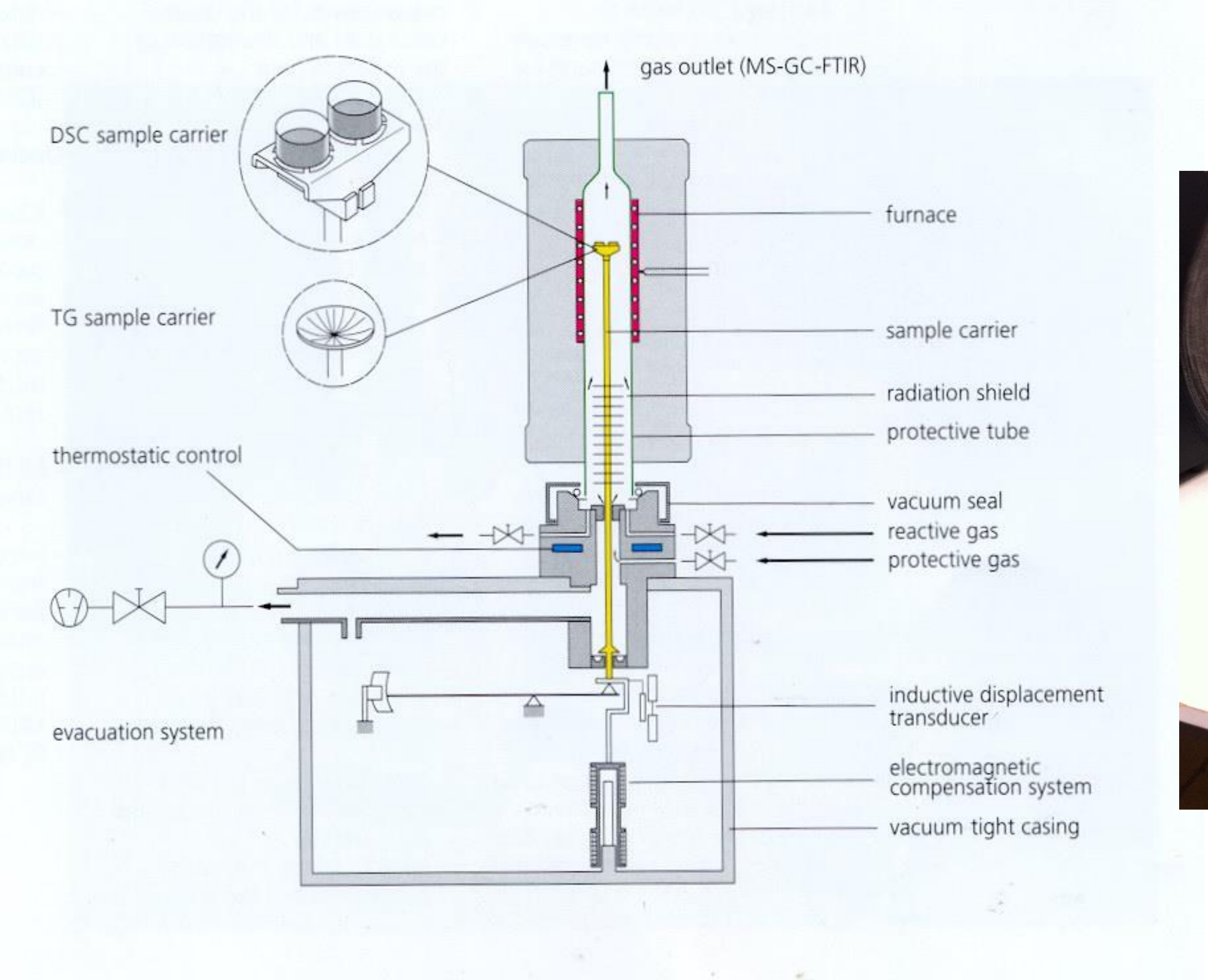
Determination of the radiative properties
- IR and thermal camera approach



Example of k measurement of a PU Foam with a guarded hot plate apparatus

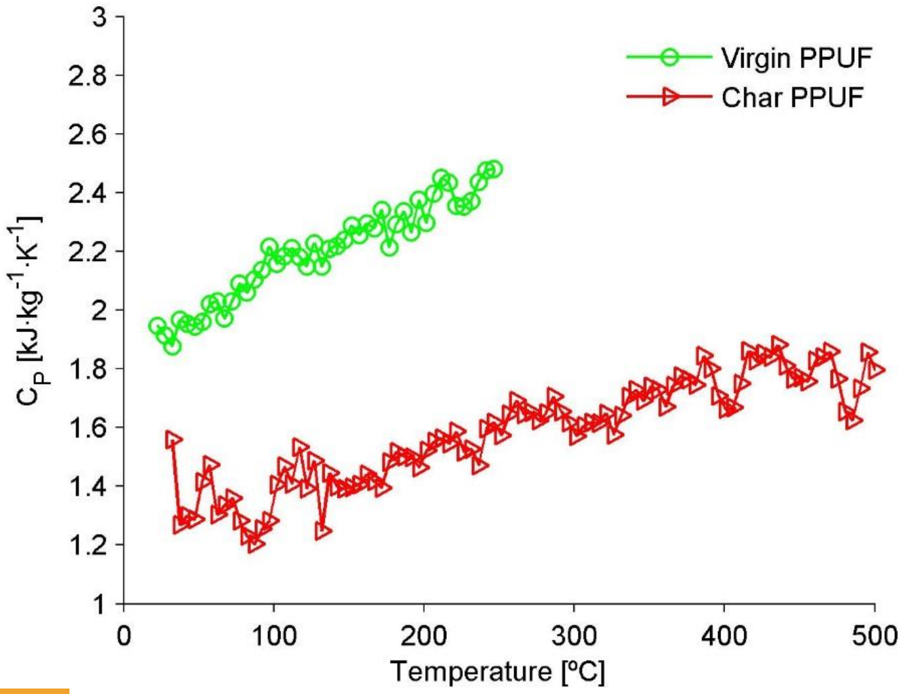
Determination of the thermochemical properties of the condensed phase

Differential Scanning Calorimetry (DSC)

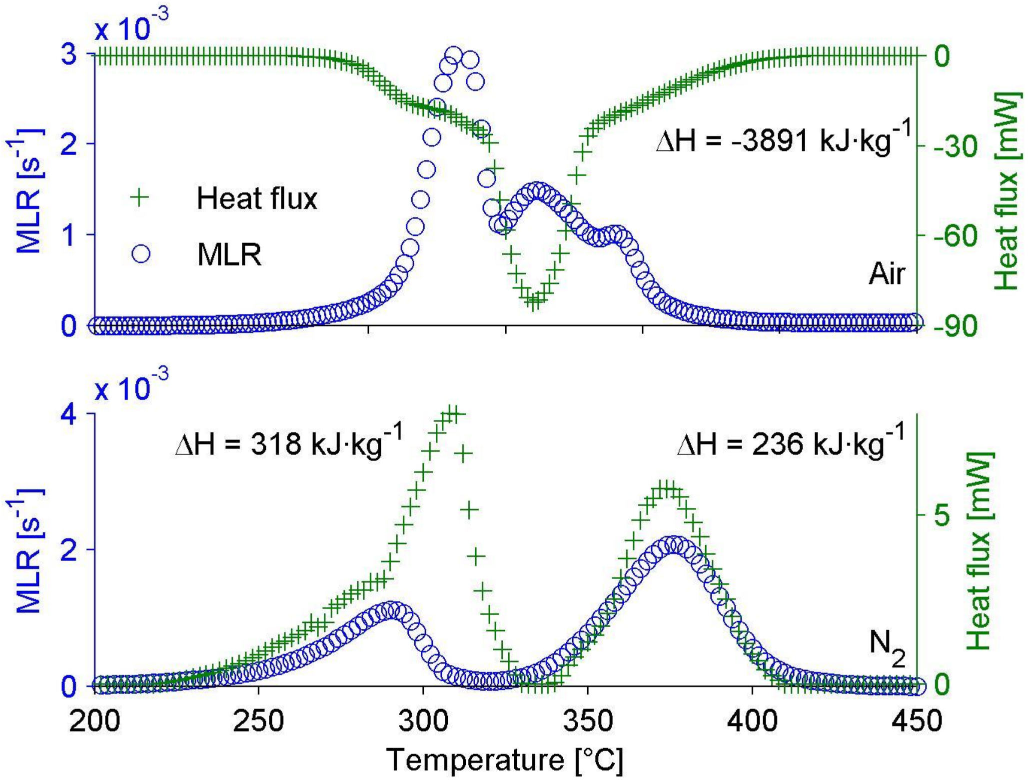


Determination of the thermochemical properties of the condensed phase

Differential Scanning Calorimetry (DSC)



Example of Cp measurement of a PU Foam in DSC

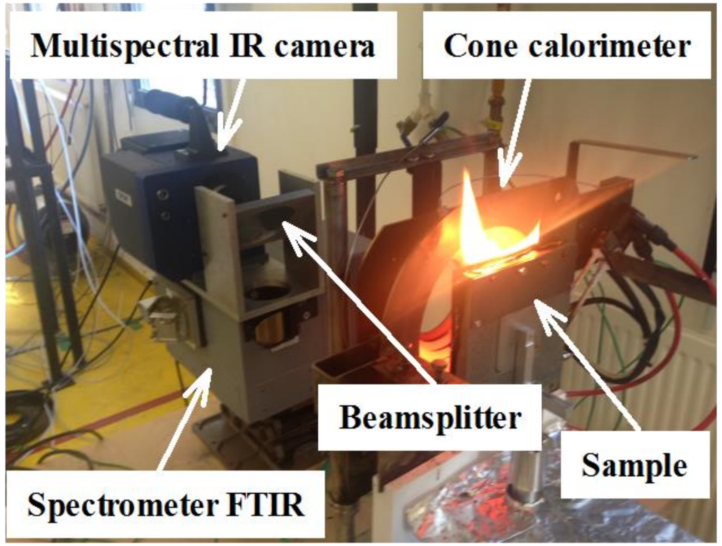


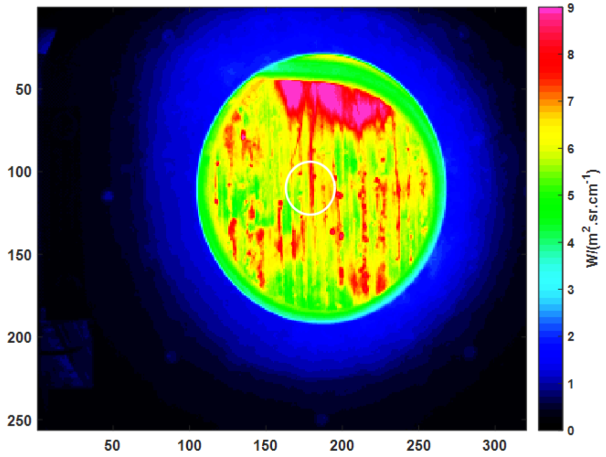
Example of ΔH measurement of a PU Foam in DSC

[Bustamante Valencia & al.]

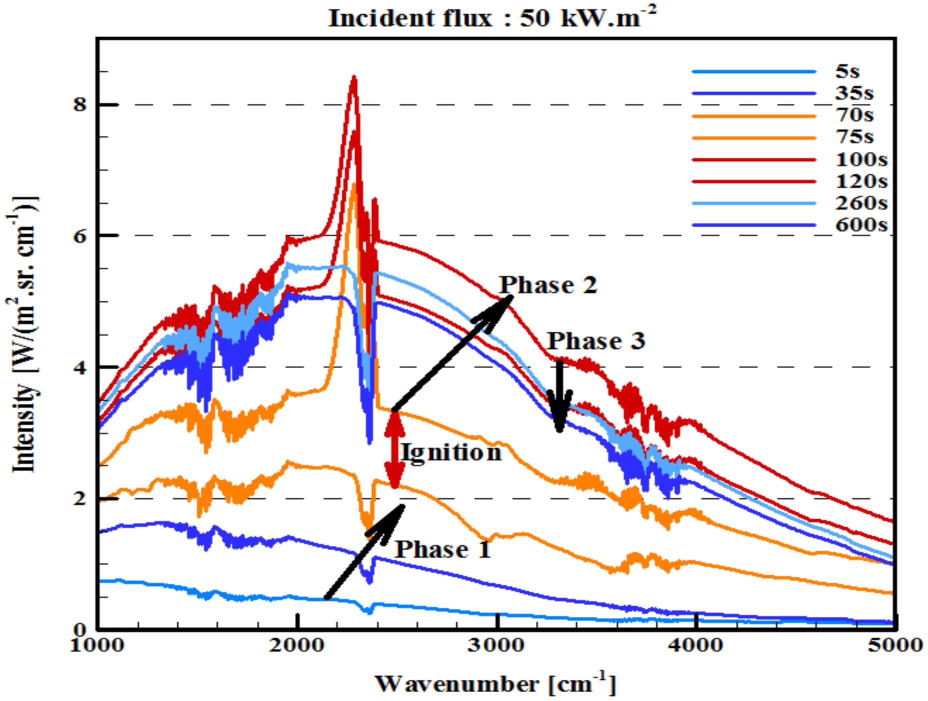
Determination of the thermochemical properties of the condensed phase

Radiative properties – IR spectrometer and thermal Camera





IR image at 2353 cm⁻¹, converted in intensity.
Composite - Incident flux 50 kW/m², time t = 80s

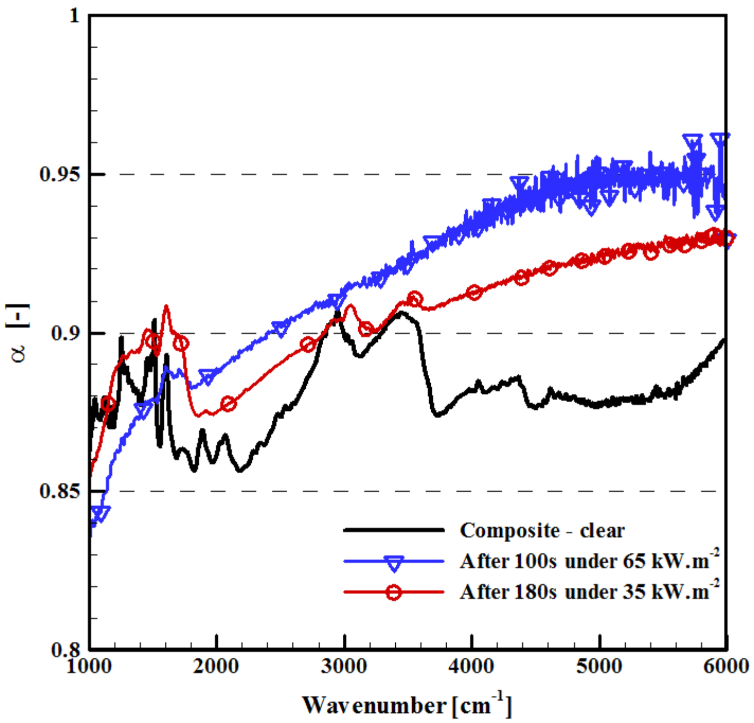


Radiative intensity emitted and reflected by a composite surface at 50 kW/m²

[Boulet & al.]

Determination of the thermochemical properties of the condensed phase

Radiative properties – IR spectrometer and thermal Camera



Radiative intensity emitted and reflected by composite surface at 50 kW/m²

Mean absorptivity of the plywood samples as a function of the irradiation duration in the cone calorimeter. Averages based on Planck's means with reference temperature of 1000 and 1200 K respectively.

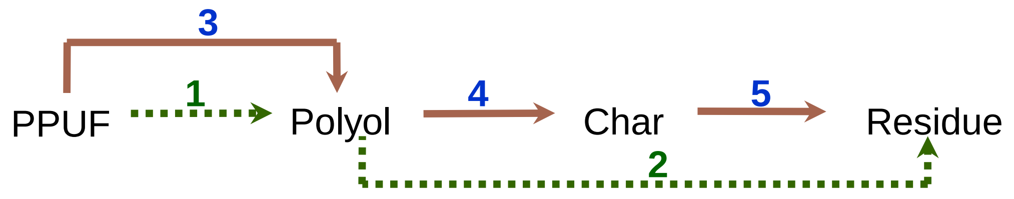
Sample and irradiation	Ref temp.	
	1000 K	1200 K
Virgin M1	0.87	0.84
M1 30 kW/m^2 during 1 min	0.85	0.83
M1 30 kW/m^2 during 10 min	0.94	0.94
M1 50 kW/m^2 during 2 min	0.93	0.93
M1 50 kW/m^2 during 10 min	0.95	0.95
Virgin M3	0.82	0.78
M3 30 kW/m^2 during 1 min	0.74	0.74
M3 30 kW/m^2 during 5 min	0.88	0.88
M3 50 kW/m^2 during 2 min	0.86	0.87
M3 50 kW/m^2 during 5 min	0.88	0.88

Plywood studied

[Boulet & al.]

Determination of the thermochemical properties of the condensed phase

Critical analysis of the experimental investigations of the thermochemical properties

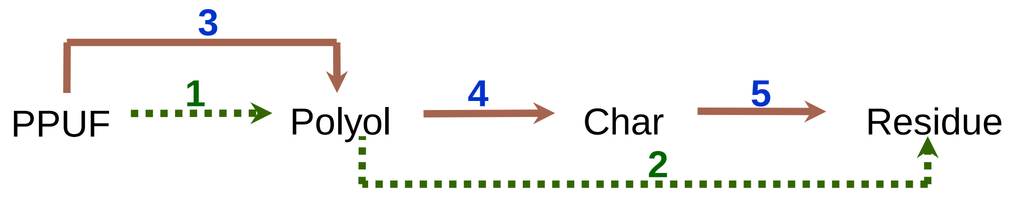


The Thermochemical properties must be determined for each « condensed specie » formed, but, **<u>Considering k</u>**:

- It is very difficult to « isolate » each condensed product in order to do the measurement required. A solution should be to degrade until a certain temperature and to stop the test. But when cooling, modification of the structure of the material… so impact on the k measurement.
- It is classically done on virgin material, between 0 to 200°C
- No thermal degradation can occured in the setup used – destruction
- The measurement when there is some char is not possible (wrong), whatever the technique of measurement

Determination of the thermochemical properties of the condensed phase

Critical analysis of the experimental investigations of the thermochemical properties

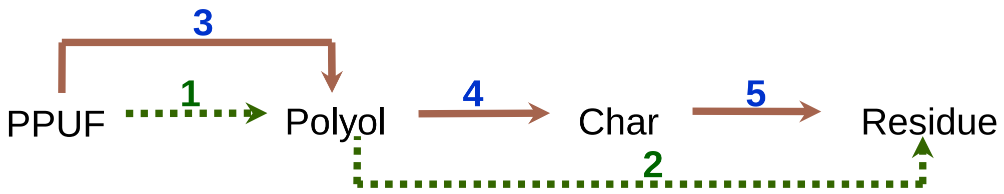


The Thermochemical properties must be determined for each « condensed specie » formed, but, **<u>Considering Cp and ΔH</u>**:

- It is very difficult to « isolate » each condensed product in order to realize the measurement required.
- In TGA-DSC apparatus:
  - The furnace is not enough performant - thermal inertia and sensibility.
  - The Cp value must be corrected by the mass loss
- In DSC apparatus: the test must be stopped at a temperature before the thermal degradation – destruction of the furnace.
- A solution should be to degrade until a certain temperature in TGA and to do some DSC analysis on the materials obtained… But the material characterized is it representative ?

Determination of the thermochemical properties of the condensed phase

Critical analysis of the experimental investigations of the thermochemical properties

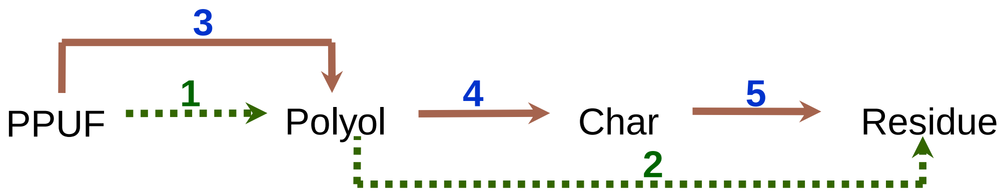


The Thermochemical properties must be determined for each « condensed specie » formed, but, **Considering ρ**:

- It is very difficult to « isolate » each condensed product in order to realize the measurement required.
- A solution should be to degrade until a certain temperature and to stop the test. But when cooling, modification of the structure of the material so of ρ…

Determination of the thermochemical properties of the condensed phase

Critical analysis of the experimental investigations of the thermochemical properties



→The Thermochemical properties must be determined for each « condensed specie » formed

It is very difficult experimentally

What are the solutions:

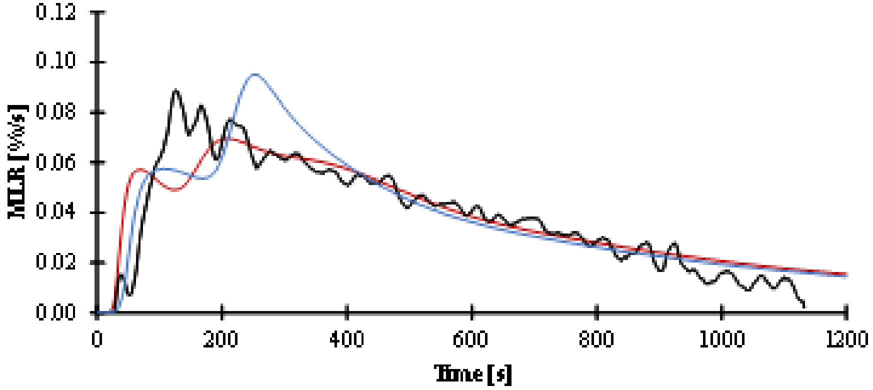
Classically we use a « weighting (average) law », between the properties of the initial material and the final one :

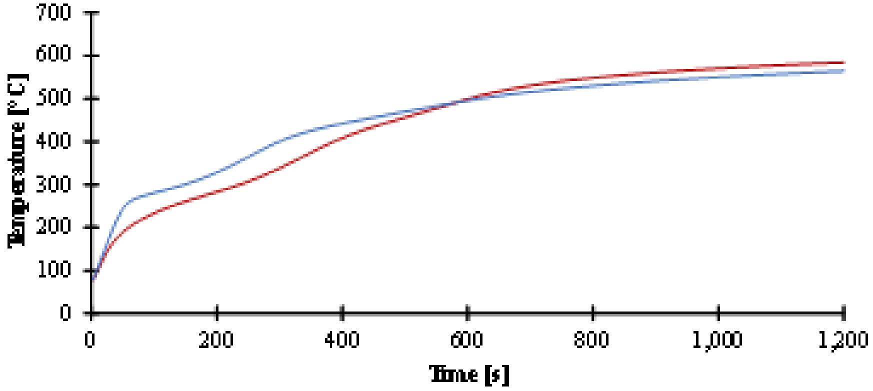
- or, equivalent properties are taken for all the « materials » - as for just 1 equivalent material.
- or, a linear evolution of the properties between the initial and the final materials is considered

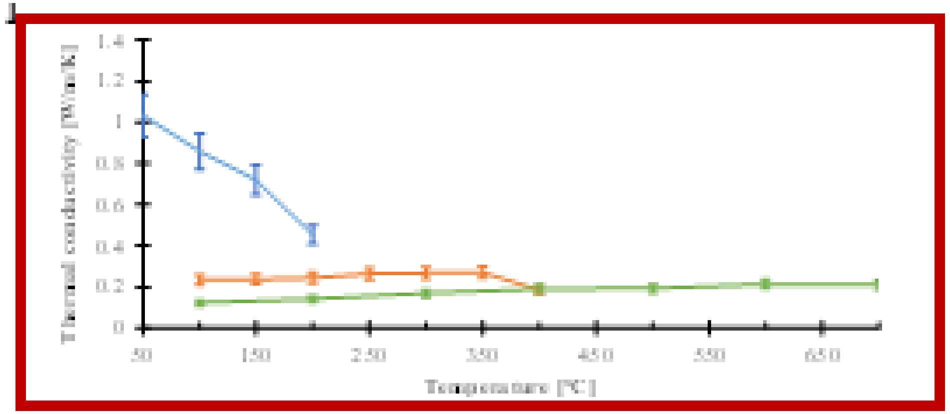
Determination of the thermochemical properties of the condensed phase

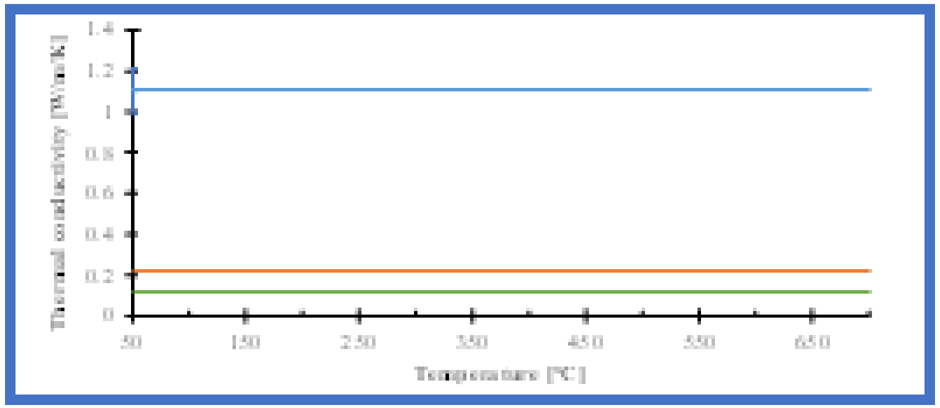
Examples of results:

$$k(T) = \frac{m_{Initial}}{m_{total}} . k_{Initial}(T) + \frac{m_{Intermediate}}{m_{total}} . k_{Intermediate}(T) + \frac{m_{Residue}}{m_{total}} . k_{Residue}(T)$$









Determination of the thermochemical properties of the condensed phase

Other solutions of determination?

- Inverse optimization methods, but…

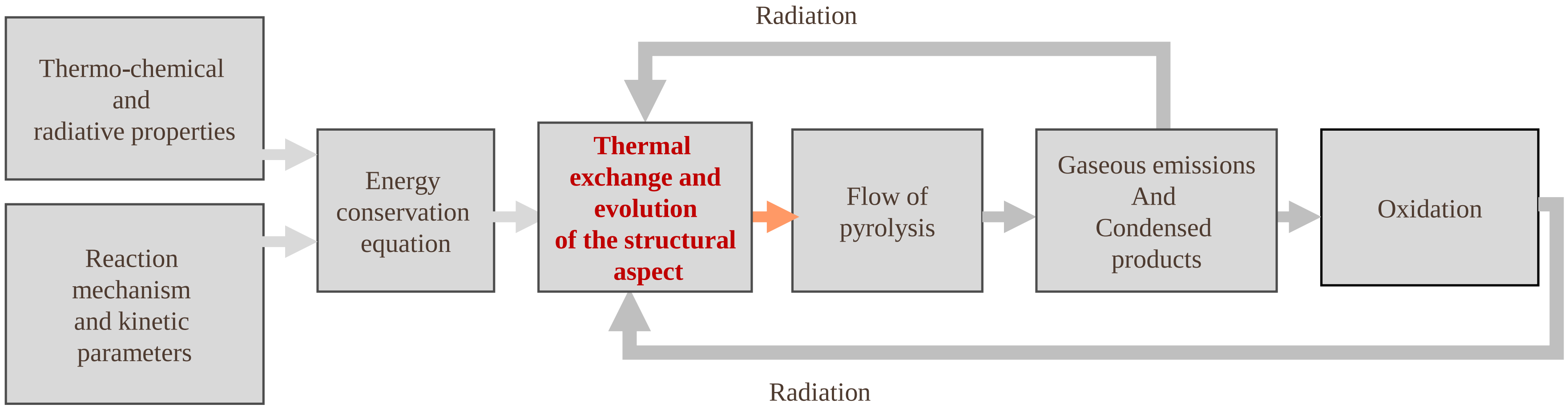
This is like some degrees of liberty, some mathematical fitting methods with more variables = the thermochemical properties

  - The results obtained are not realistics and physics
  - Compensation phenomenon are observed due to the large unknown parameters (the kinetics and the thermochemical ones)

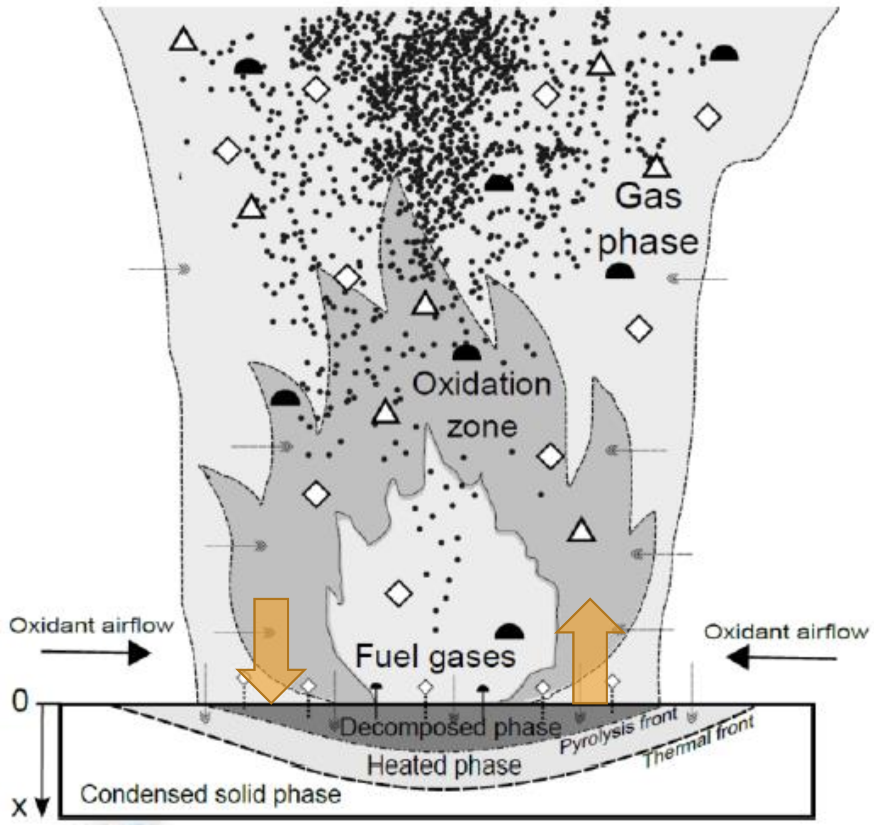
→ The determination of the thermochemical properties is a main challenge

This is the target of numerous actual research work

Thermal decomposition
3rd phase: Heat feedback and oxygen diffusion



Heat feedback and oxygen diffusion - Determination



This description requires the description of :

- What happen in the gas phase: flame, temperature.
- The MLR flow and the species diffusion.
- The mass transfer diffusion, from the condensed phase to the gas one, and from the gas phase to the solid.
- The thermal transfer between the gas phase and the solid one.

How to describe them ?
How to determine Them ?

In controlled and defined conditions

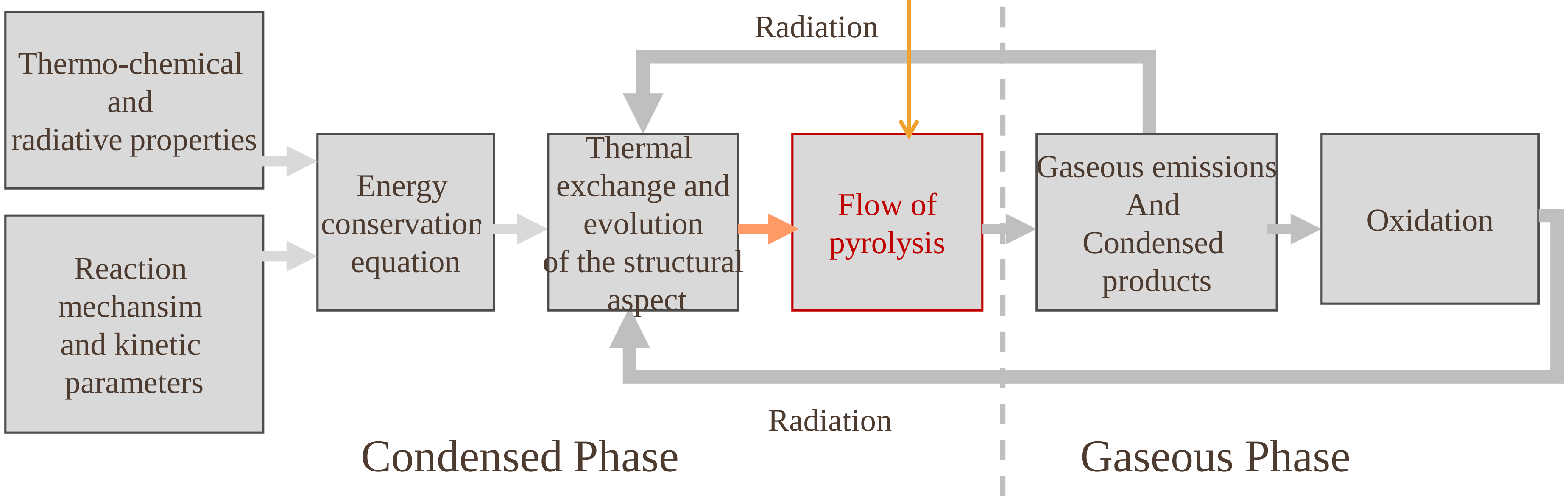
Multi-scale experimental investigations

Thermal decomposition
3rd phase: Heat feedback and oxygen diffusion

Special focus on the experimental benchscales

Modeling of Thermal decomposition – Required parameters

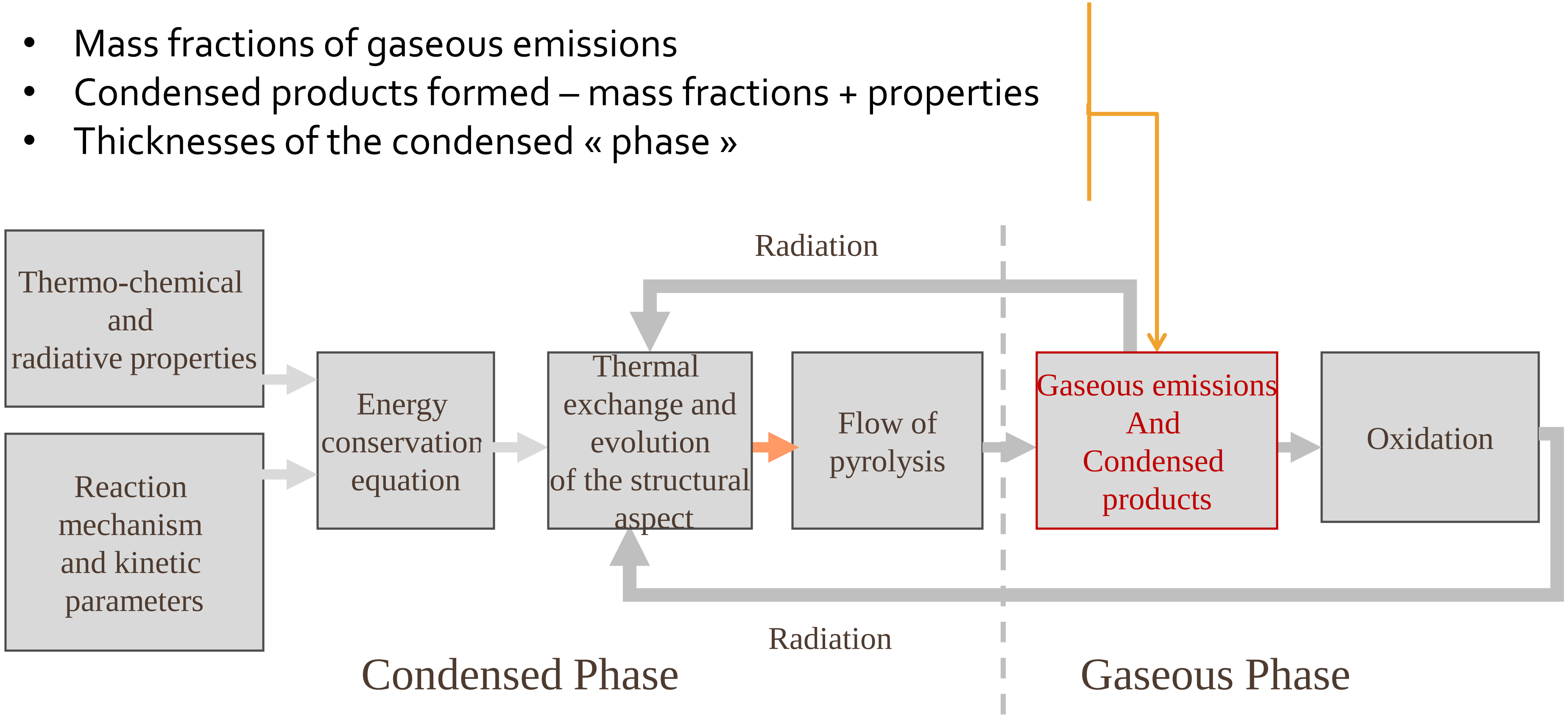
- Mass loss
- Mass loss rate- MLR
- Regression of the surface (thickness)



--> equ. of mass conservation
--> equ. of movement quantity conservation (rate)
--> equ. of energy conservation (T)
--> equ. of species conservation (mass fractions Y_i)

Modeling of Thermal decomposition – Required parameters

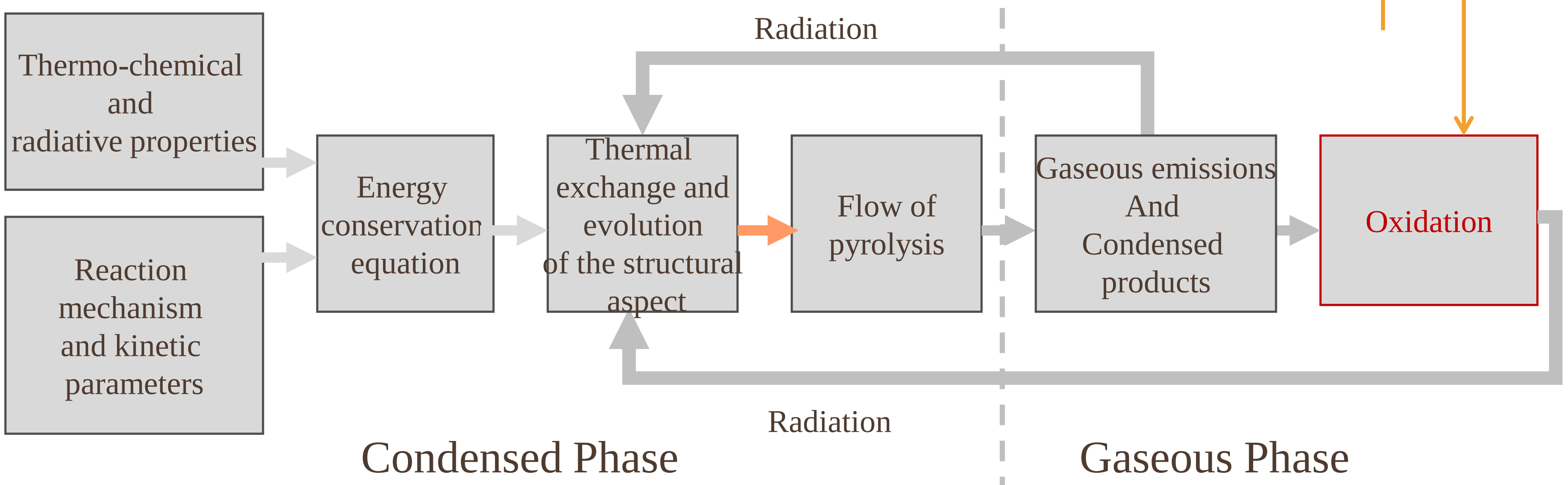
- Mass fractions of gaseous emissions
- Condensed products formed – mass fractions + properties
- Thicknesses of the condensed « phase »



--> equ. of mass conservation
--> equ. of movement quantity conservation (rate)
--> equ. of energy conservation (T)
--> equ. of species conservation (mass fractions Y_i)

Modeling of Thermal decomposition – Required parameters

- Delays (ignition, extinction…)
- Gaseous emissions mass fractions
- Heat release rate HRR
- Flame structure and dynamic
- Temperatures



--> equ. of mass conservation
--> equ. of movement quantity conservation (rate)
--> equ. of energy conservation (T)
--> equ. of species conservation (mass fractions Y_i)

Focus: Experimental Investigations – Thermal decomposition

Matter Scale	Small Scale	Product scale	Real Scale

0D process

- No thermal diffusion
- No mass transfer
- « Perfect » control of the boundary conditions
- Controlled conditions: imposed by the user

- Intrinsic properties

1D process

- 1D Thermal transfer
- Radiative properties
- 1D mass transfer
- We try to control the boundary and the experimental conditions
- No flame transfer-propagation

- Simplified characterization of the diffusion aspects

2D or 3D process

- Thermal transfer
- Radiative properties
- Mass transfer
- Uncontrolled or difficulty to control the boundary conditions and the experimental ones
- Flame transfer-propagation

- Real diffusion aspects but… many process invoved and difficulty to determine the « conditions »

Approach: to simplify the problem and to add complexity as we go along

Focus: Experimental Investigations – Thermal decomposition

Matter Scale

- Mass Loss
- Mass loss rate
- Enthalpy of the reactions
- Heat of combustion
- Specific heat
- Model of pyrolysis (mechanisms and parameters)

Small Scale

- Delays of inflammation, extinction
- Mass loss
- Mass loss rate
- HRR
- Temperatures
- Gaseous emissions concentrations
- Thermal inertia
- Effusivity

Product scale

Real Scale

- Delays of inflammation, extinction, flash-over
- Mass loss
- Mass loss rate
- Gaseous emissions
- HRR
- Flame propagation
- Temperatures

Focus: Experimental Investigations – Thermal decomposition

Cône Calorimeter (CC) or Fire Propagation Apparatus (FPA) :

- Few g and (100*100*Thickness) mm³ samples
- Heat flux: 0 to 100 kW/m²
- Possible controlled atmosphere
- Coupling with gas analysis
- Piloted ignition or auto-ignition

✓ Determination of

- t_{ig}, T_{ig}, CHF
- ML and MLR
- HRR
- Gaseous emissions
- Temperatures into the condensed and the gas phases
- And other parameters can be deducted

Small Scale

Controled parameters

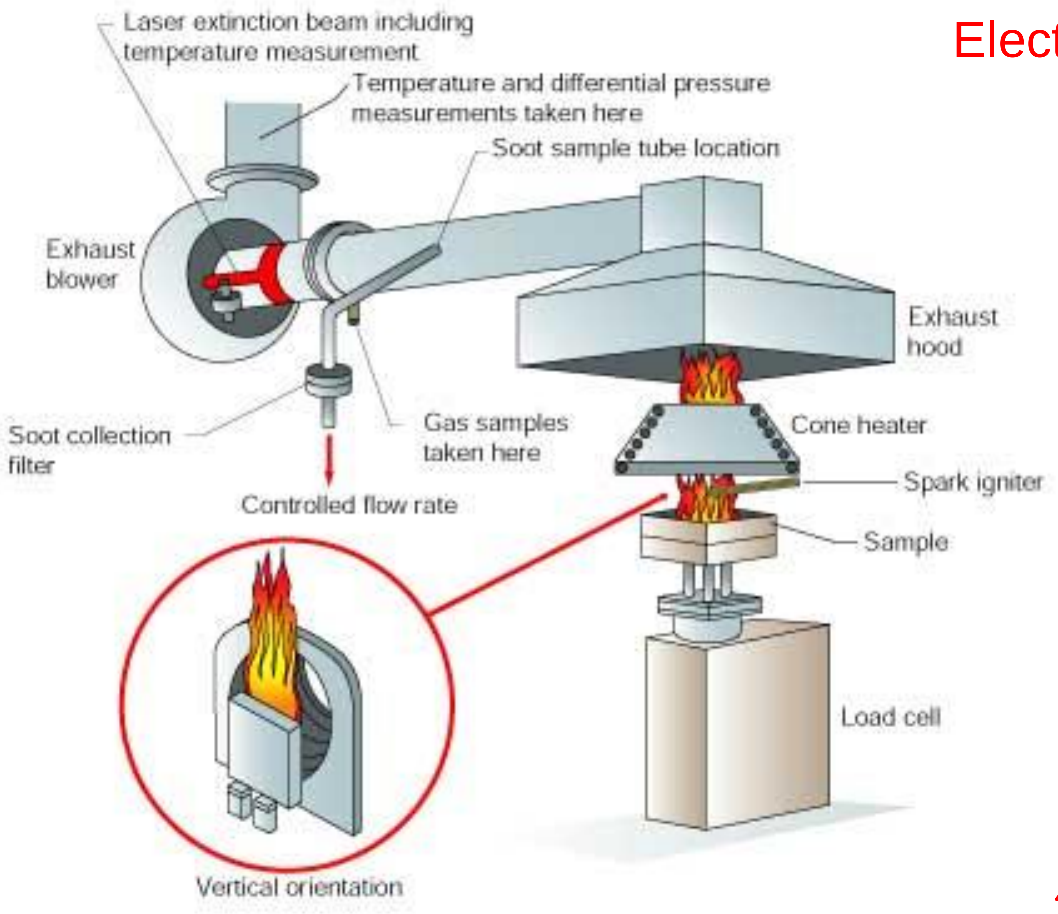
- Heat flux
- Global atmosphere

Unknown parameters

- Temperature
- Heating rate
- Local mass fraction of combustible and of oxygen
- Permeability of the solid
- Thickness of the reactive zone

Simplified « 1D » investigation, with a flame Heat and mass tranfers whithout a flame propagation

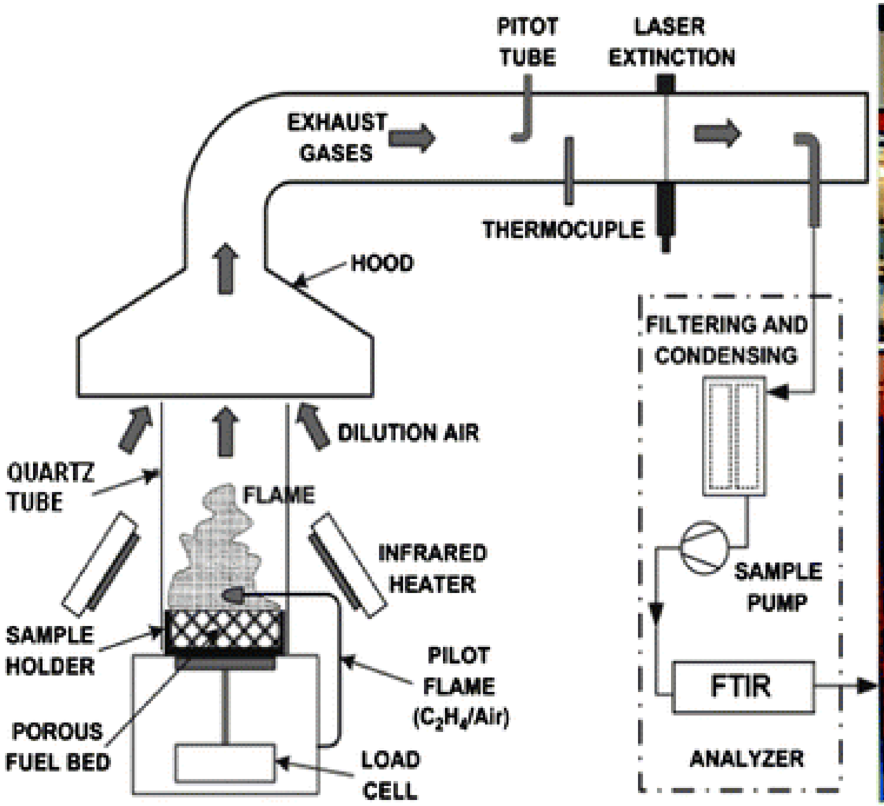
Cone Calorimeter (CC)



Electric heater
Hood

Weighing device
Sample

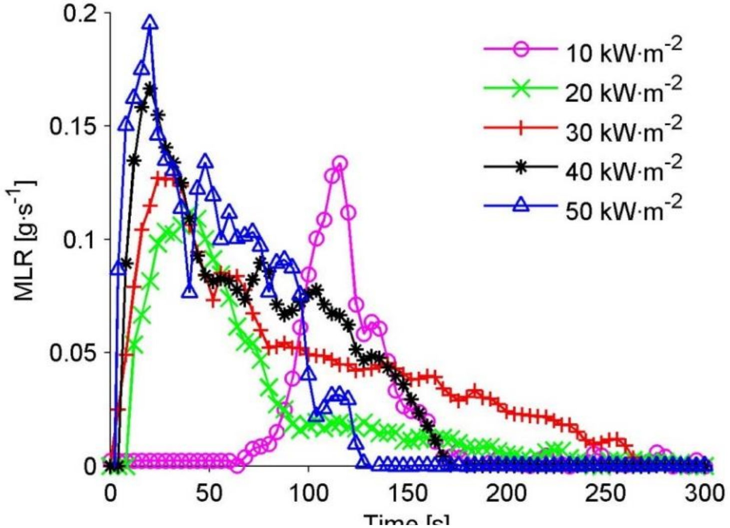
Fire Propagation Apparatus (FPA):

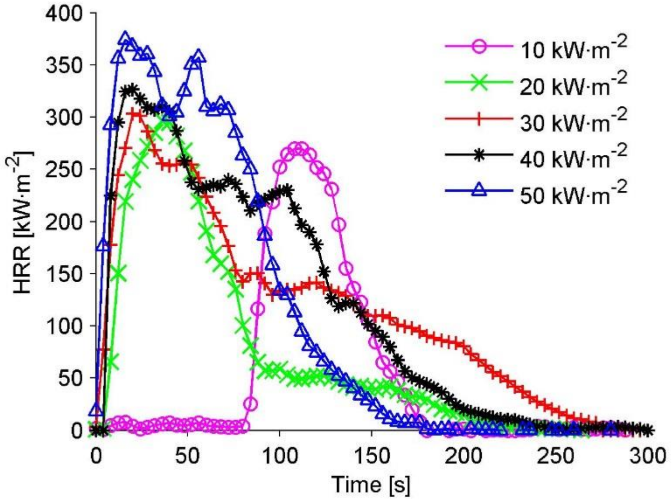


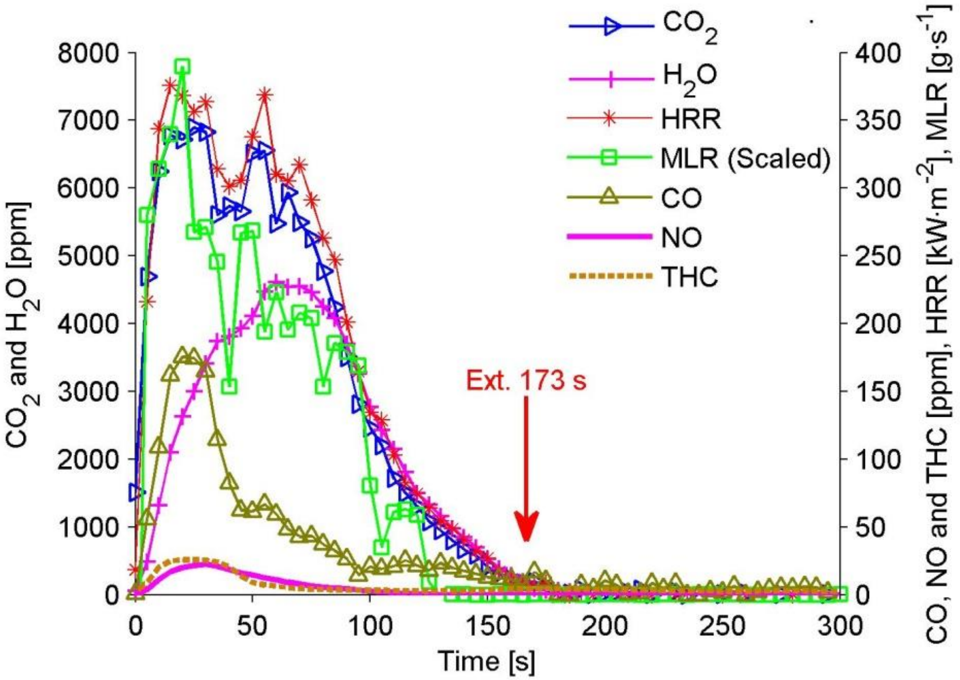
[*Diallo & al. An innovative experimental approach aiming to understand and quantify the actual fire hazards of ionic liquids, Energy and environmental science, 2013*]

Focus: Experimental Investigations – Thermal decomposition

Examples of results:





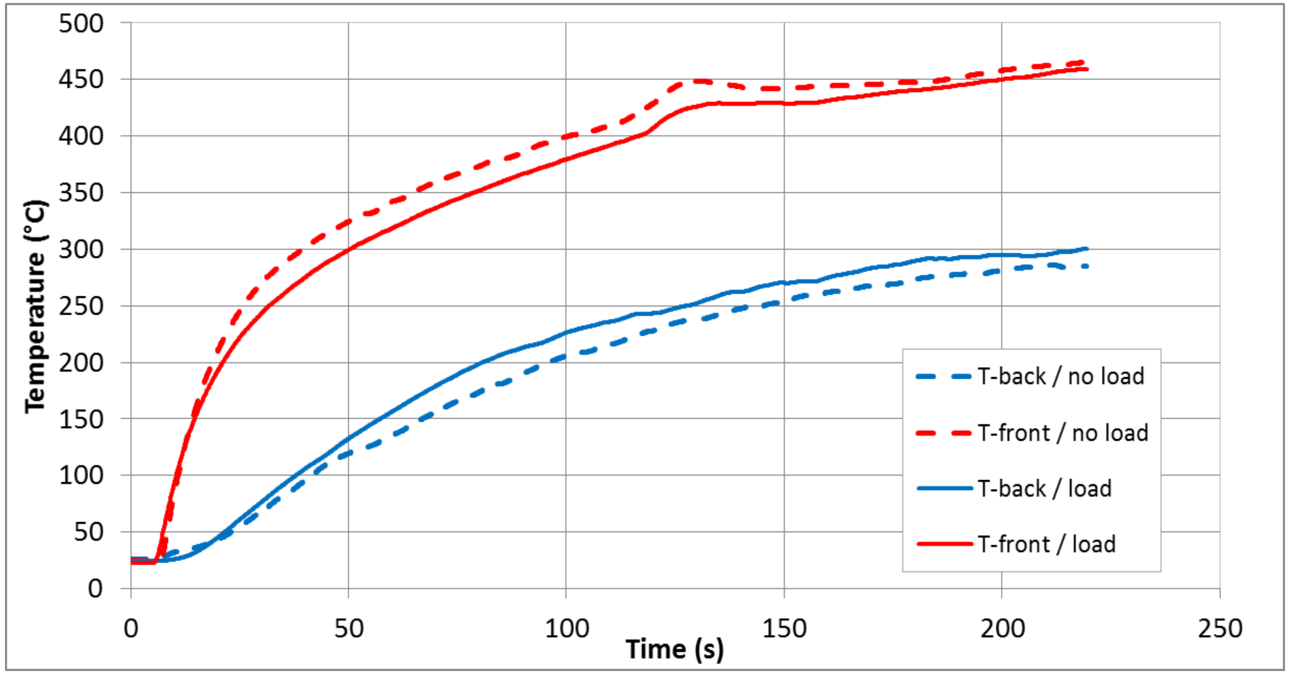


Thermal decomposition of a PU Foam

[L. Bustamante Valencia & al. Analysis of principal gas products during combustion of polyether polyurethane foam at different irradiance levels. Fire Safety Journal, 2009.]

Focus: Experimental Investigations – Thermal decomposition

Examples of results:



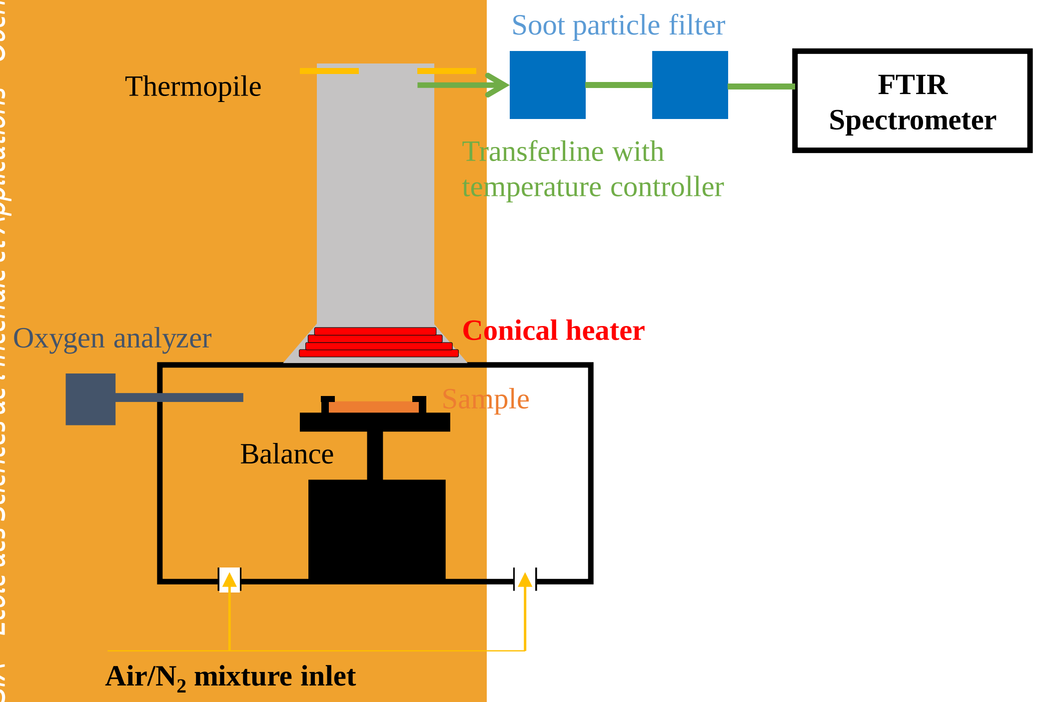
Temperatures evolution during the thermal decomposition of a composite material in CC

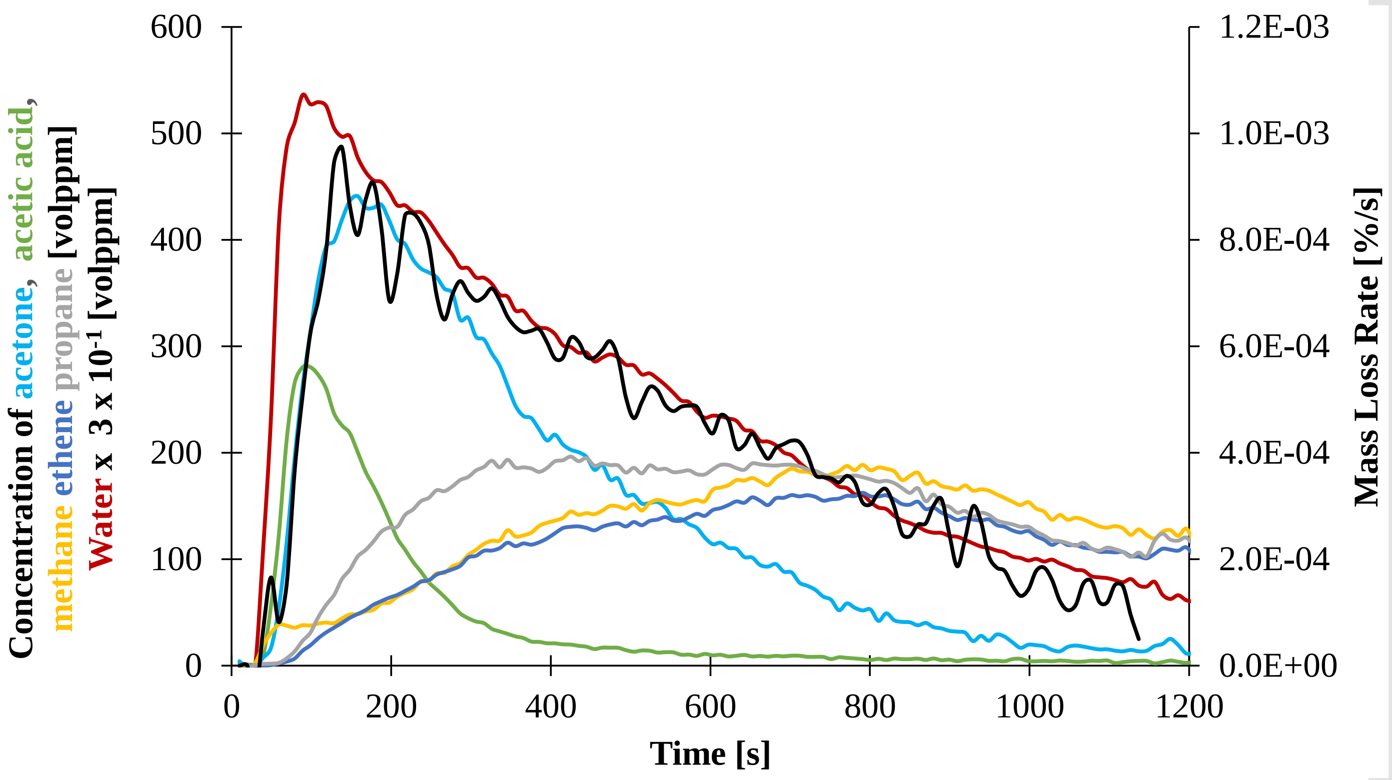
[*A. Benelfellah & al. Effect of a coupled thermomechanical loading on the residual mechanical strength and on the surface temperature of wound carbon/epoxy composite, Journal of Composite Material, 2017*]

Focus: Experimental Investigations – Thermal decomposition

Examples of results:

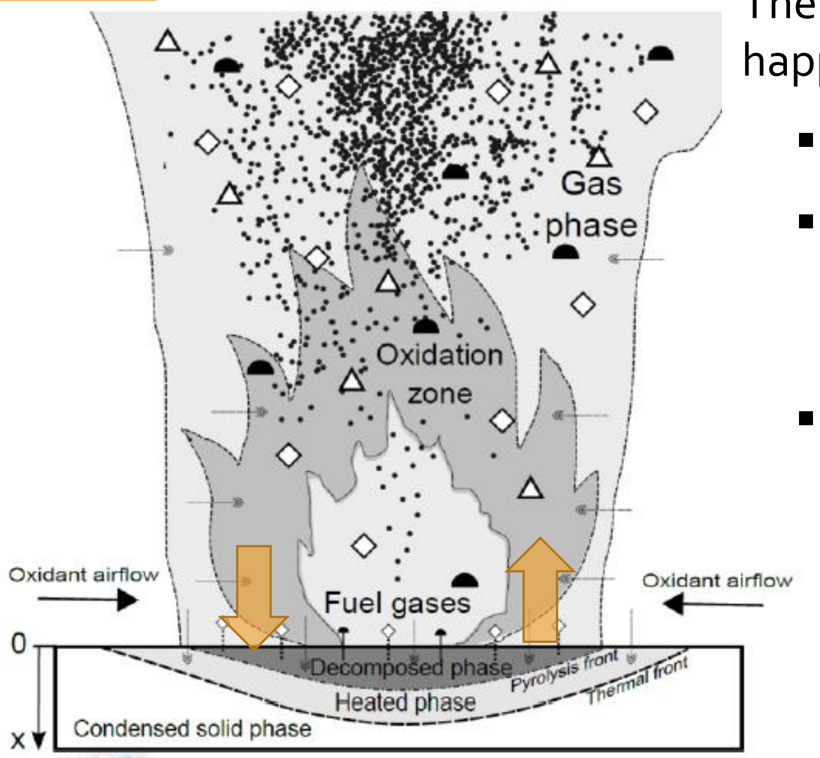
Quantification of the decomposition gases : Controlled-Atmosphere Mass Loss Calorimeter coupled FTIR





Example of results obtained for EVA/ATH

Heat feedback and oxygen diffusion - Modelling

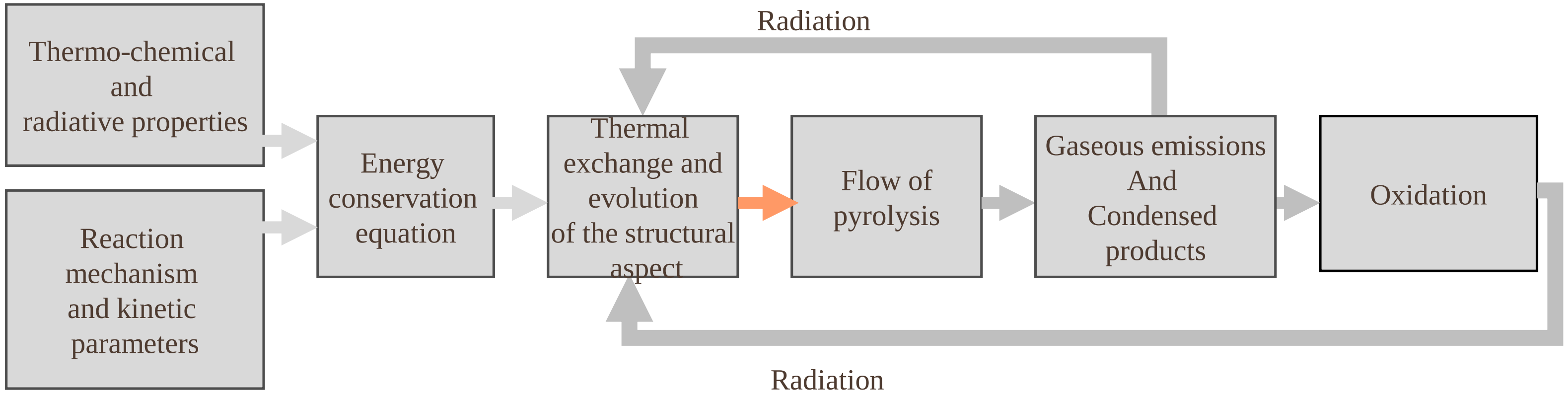


The **objective** is to describe the thermal decomposition and what happen at this scale:

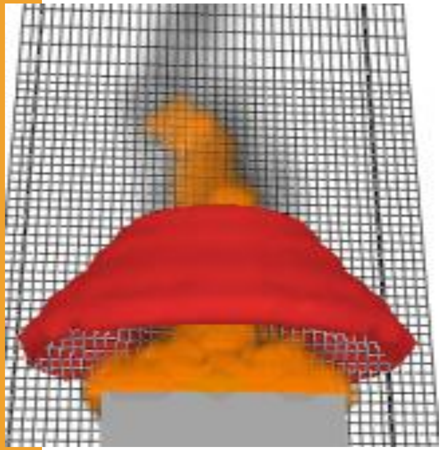
- CFD modeling
- Integration of the « model of pyrolysis » and the « thermochemical data » in the CFD model (as FDS or Firefoam)
- Modeling of the cone calorimeter or the FPA experiments

Validation of the model of pyrolysis at this scale, with thermal and mass transfers

Thermal decomposition 4th phase: validation of the thermal decomposition model



Validation of the thermal decomposition model

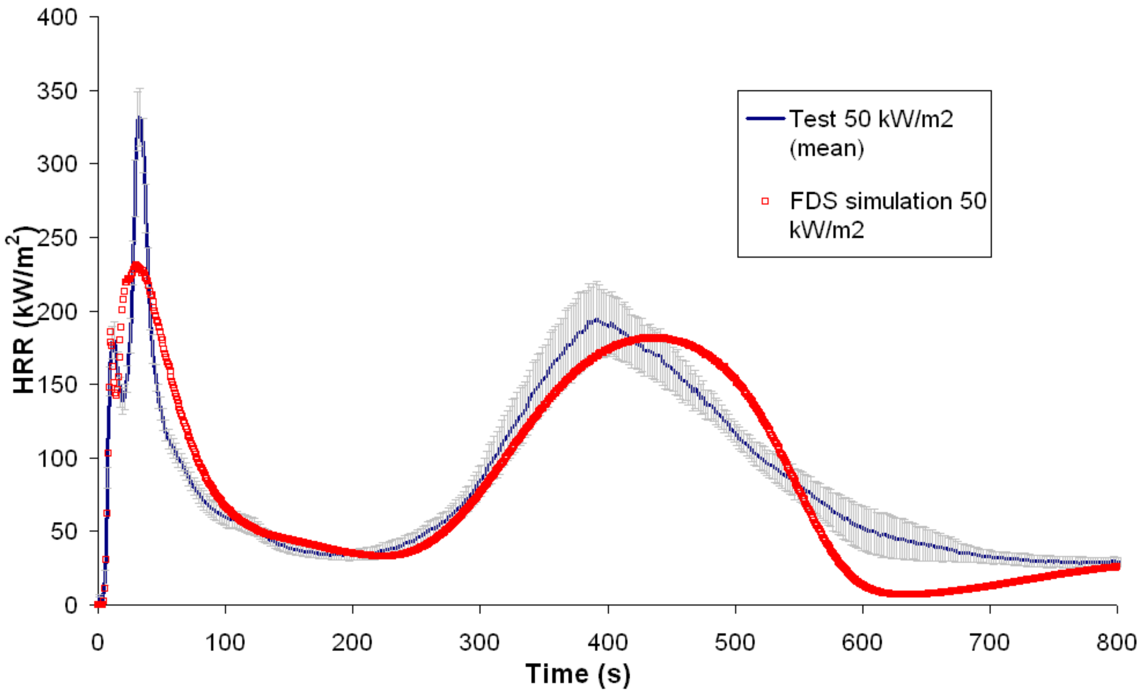


Small Scale

Model of pyrolysis

Thermo-chemical properties

Cone calorimeter results ⟺ Prediction of thermal decomposition at small scale



[A. Camillo & al.]

Focus: Experimental Investigations – Thermal decomposition

Matter Scale	Small Scale	Product scale / Real Scale
0D process	**1D process**	**2D or 3D process**
▪ No thermal diffusion ▪ No mass transfer ▪ « Perfect » control of the boundary conditions ▪ Controlled conditions: imposed by the user ▪ Intrinsic properties	▪ 1D Thermal transfer ▪ Radiative properties ▪ 1D mass transfer ▪ We try to control the boundary and the experimental conditions ▪ No flame transfer-propagation ▪ Simplified characterization of the diffusion aspects	▪ Thermal transfer ▪ Radiative properties ▪ Mass transfer ▪ Uncontrolled or difficulty to control the boundary conditions and the experimental ones ▪ Flame transfer-propagation ▪ Real diffusion aspects but… many process invoved and difficulty to determine the « conditions »

Approach: to simplify the problem and to add complexity as we go along

Focus: Experimental Investigations – Thermal decomposition

Matter Scale

- Mass Loss
- Mass loss rate

- Enthalpy of the reactions
- Heat of combustion
- Specific heat

- Model of pyrolysis (mechanisms and parameters)

Small Scale

- Delays of inflammation, extinction
- Mass loss
- Mass loss rate
- HRR
- Temperatures
- Gaseous emissions concentrations

- Thermal inertia
- Effusivity

Product scale

Real Scale

- Delays of inflammation, extinction, flash-over
- Mass loss
- Mass loss rate
- Gaseous emissions
- HRR
- Flame propagation
- Temperatures

Focus: Experimental Investigations – Thermal decomposition

Product scale

Classical experimental investigations:

- IMO LIFT
- Medium Burning Item
- Single Burning Item

→ **HRR**
Rate of flame
Temperatures
Flame structure

- Open calorimeter (Nordtest NT Fire 032 calorimeter)
- Combustion chamber (ISO 9705 – room corner test)

→ **HRR**

[V. Brabauskas, Heat release rate, The SFPE Handbook of fire protection engineering 4th edition]

Focus: Experimental Investigations – Thermal decomposition

Product scale

Single Burning Item (SBI), Medium and LIFT :

- Mass: kg
- Atmosphere: under air
- Piloted ignition by a burner or a piloted flame
- Gaseous analysis possible

Medium or LIFT: vertical or horizontal samples

SBI: possibility to « create » structures

Controlled parameters

- LIFT: incident heat flux
- SBI and medium : P_{burner}

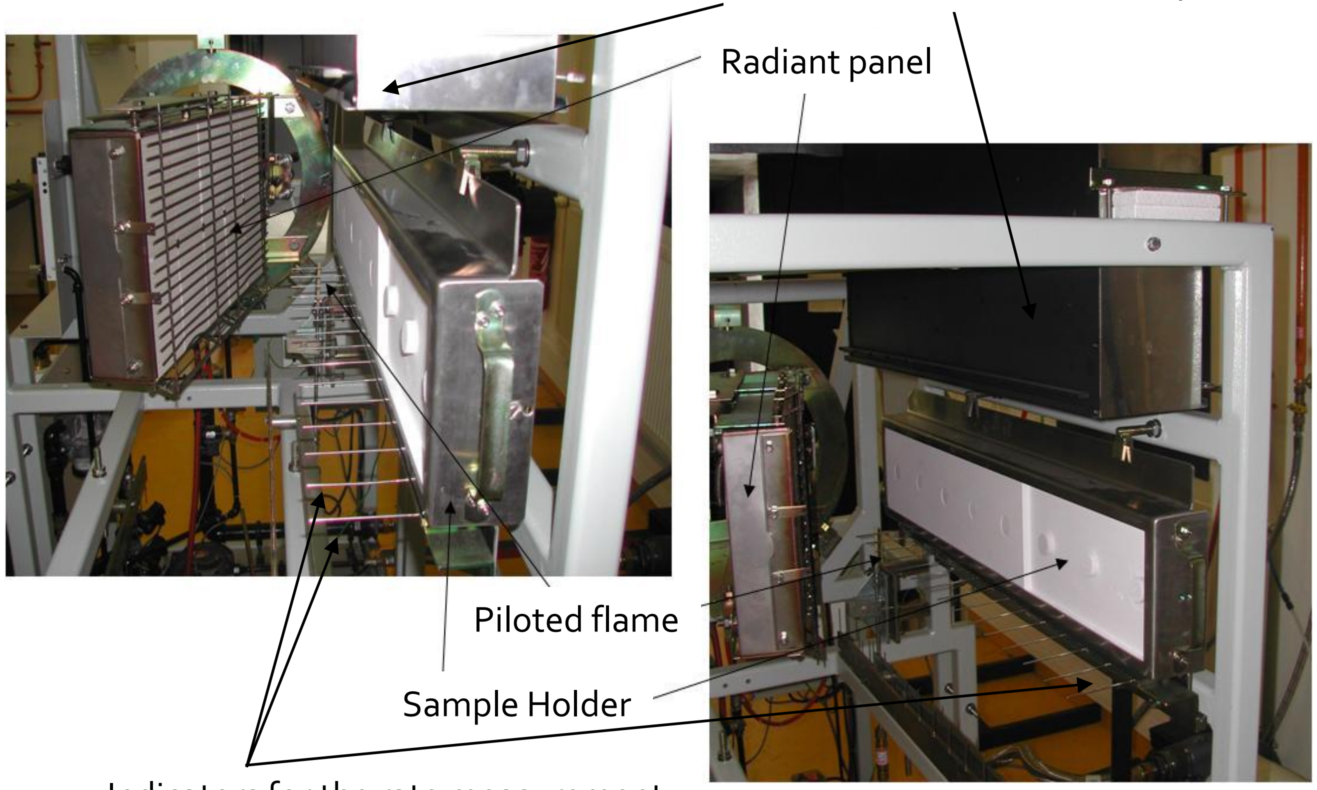
Unknown parameters

- Temperature
- Heating rate
- Local mass fractions of combustible and oxygen
- Permeability
- Gradient of oxygen (diffusion) in the condensed phase
- Gradient of temperature
- Reactive zone thickness

Focus: Experimental Investigations – Thermal decomposition

IMO – LIFT :

Product scale



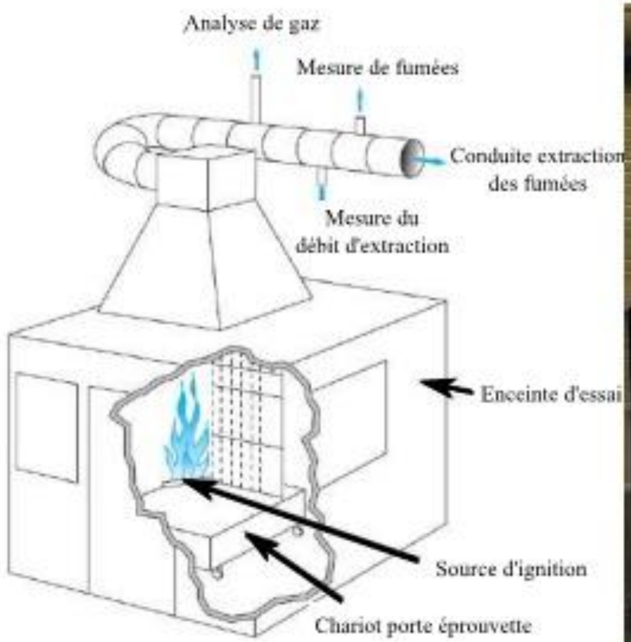
IMO/LIFT Spread Of Flame Apparatus (ISO 5658)

Focus: Experimental Investigations – Thermal decomposition

Medium and Single Burning Items:

Product scale

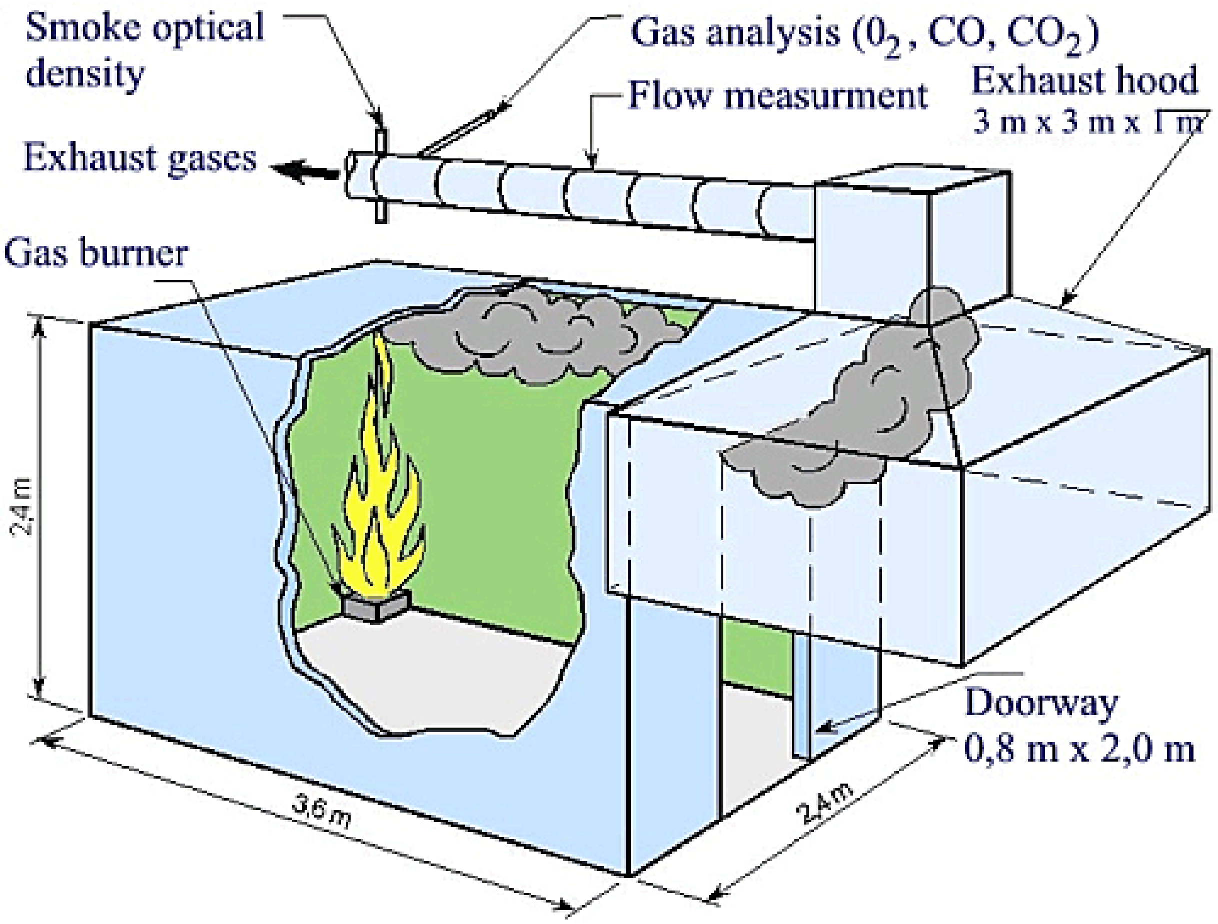
Medium Burning Item

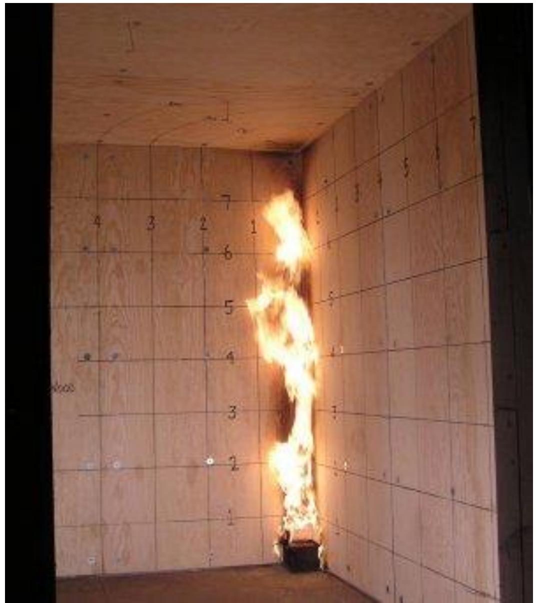


Single Burning Item

ISO 9705 - Room corner test:

Product scale



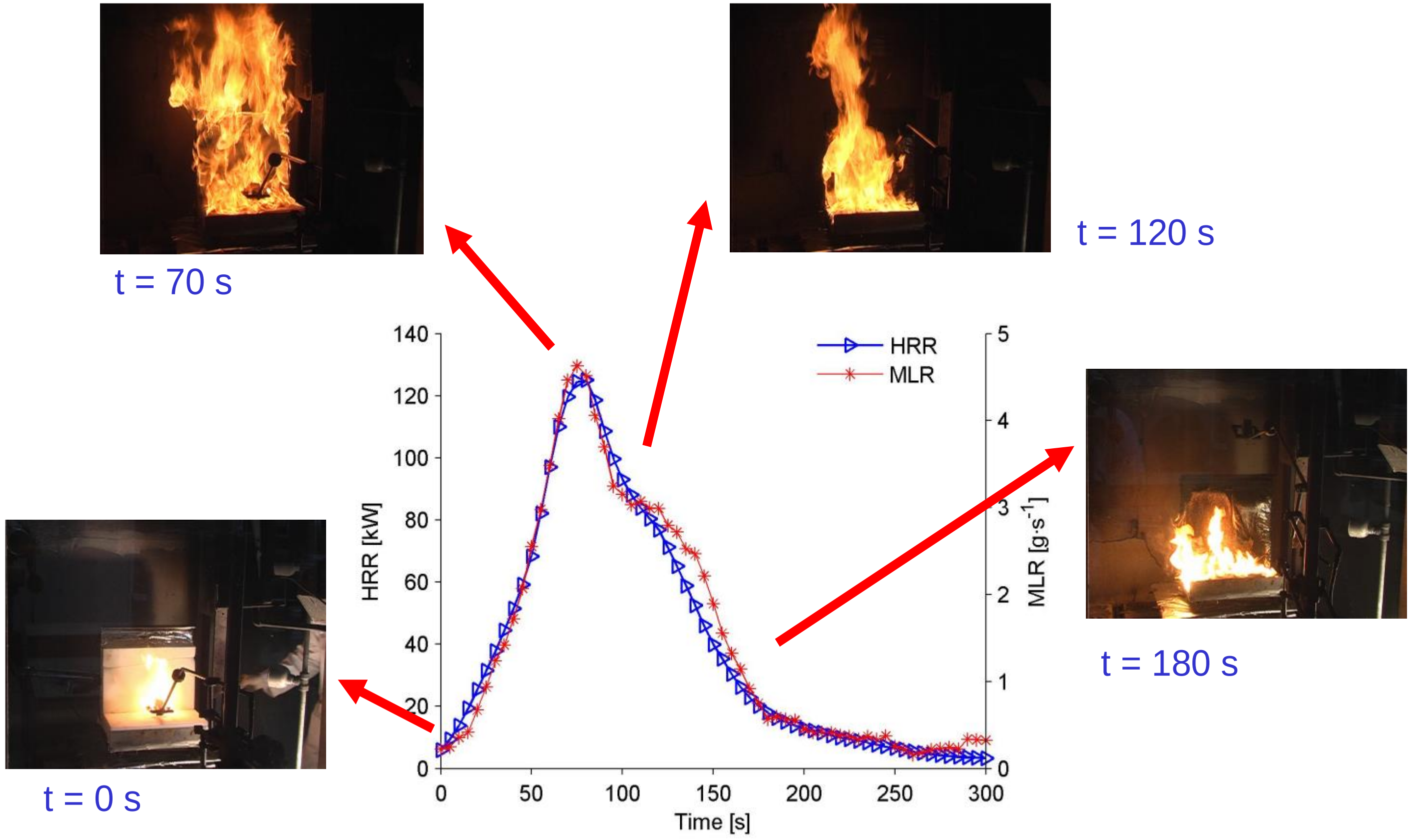


Focus: Experimental Investigations – Thermal decomposition

SBI et Medium:

Product scale

- Example of results



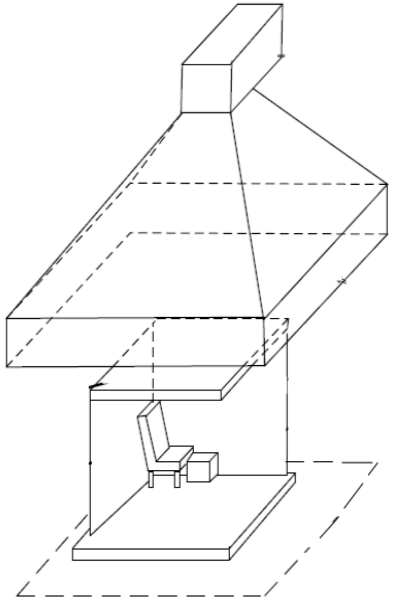
Thermal decomposition and combustion of a PU Foam into a SBI

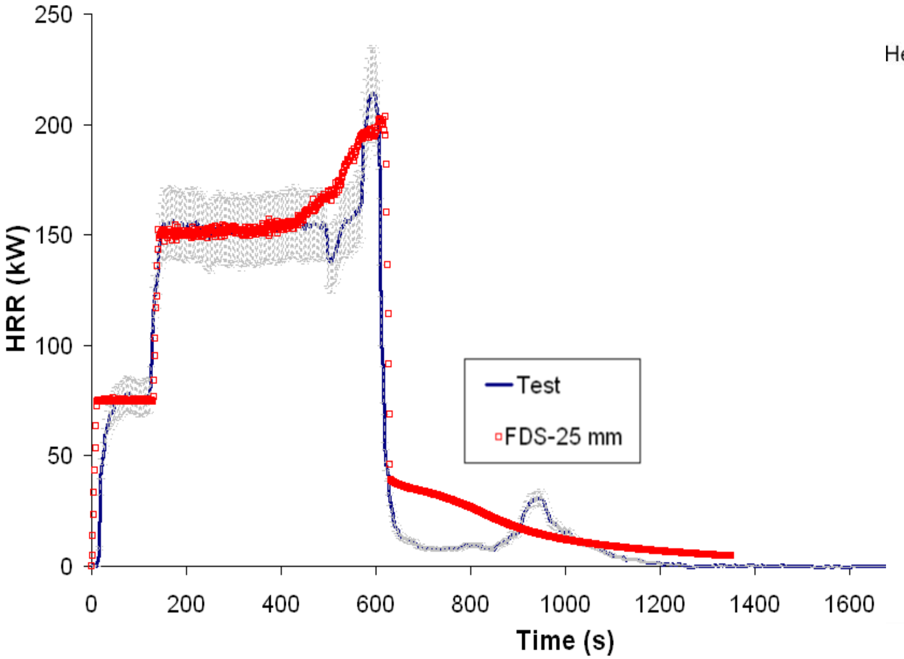
[L. Bustamante Valencia& al. Analysis of principal gas products during combustion of polyether polyurethane foam at different irradiance levels. Fire Safety Journal, 2009.]

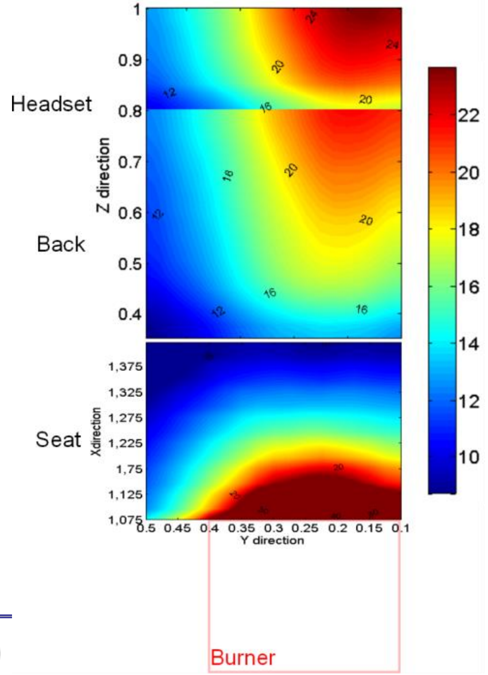
Focus: Experimental Investigations – Thermal decomposition

A : 30 s | B : 270 s | C : 480 s | D : 510 s | E : 600 s

SBI:







Heat flux density distribution

[E. Guillaume & al. Application and Limitations of a Method Based on Pyrolysis Models to Simulate Railway Rolling Stock Fire Scenarios. Fire Technology, 50, pp. 317-348, 2014]

Focus: Experimental Investigations – Thermal decomposition

Training device of the firemen (container)

Real scale

[Colloque national – Contraintes thermiques et performance des EPI des sapeurs-pompiers en caisson d'observation et d'entrainement aux phénomènes thermiques, Niort, 2014]

Focus: Experimental Investigations – Thermal decomposition

Rooms and building

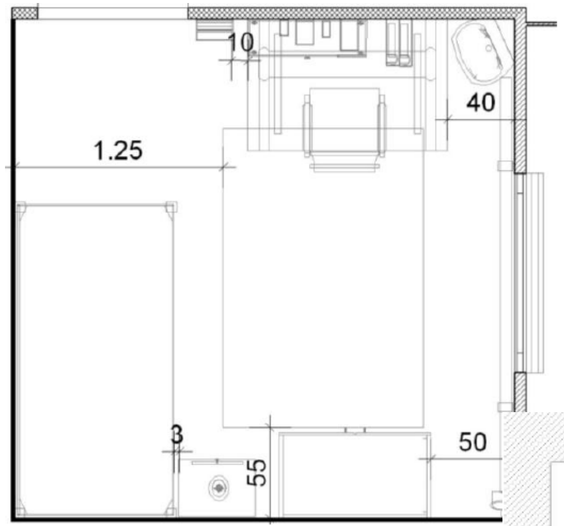
Real Scale

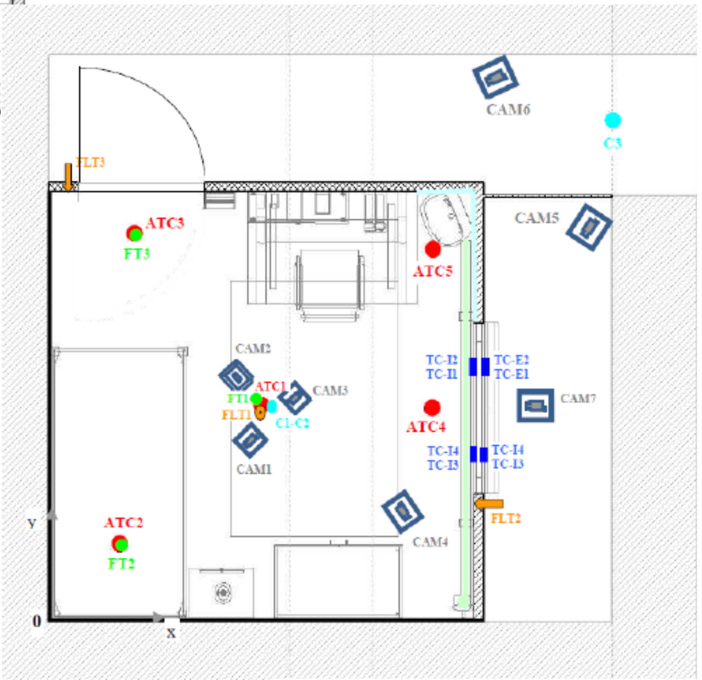
[Essais Dalmarnock, Université d'Edinburgh]

Focus: Experimental Investigations – Thermal decomposition

<u>Rooms and building</u>

Real Scale



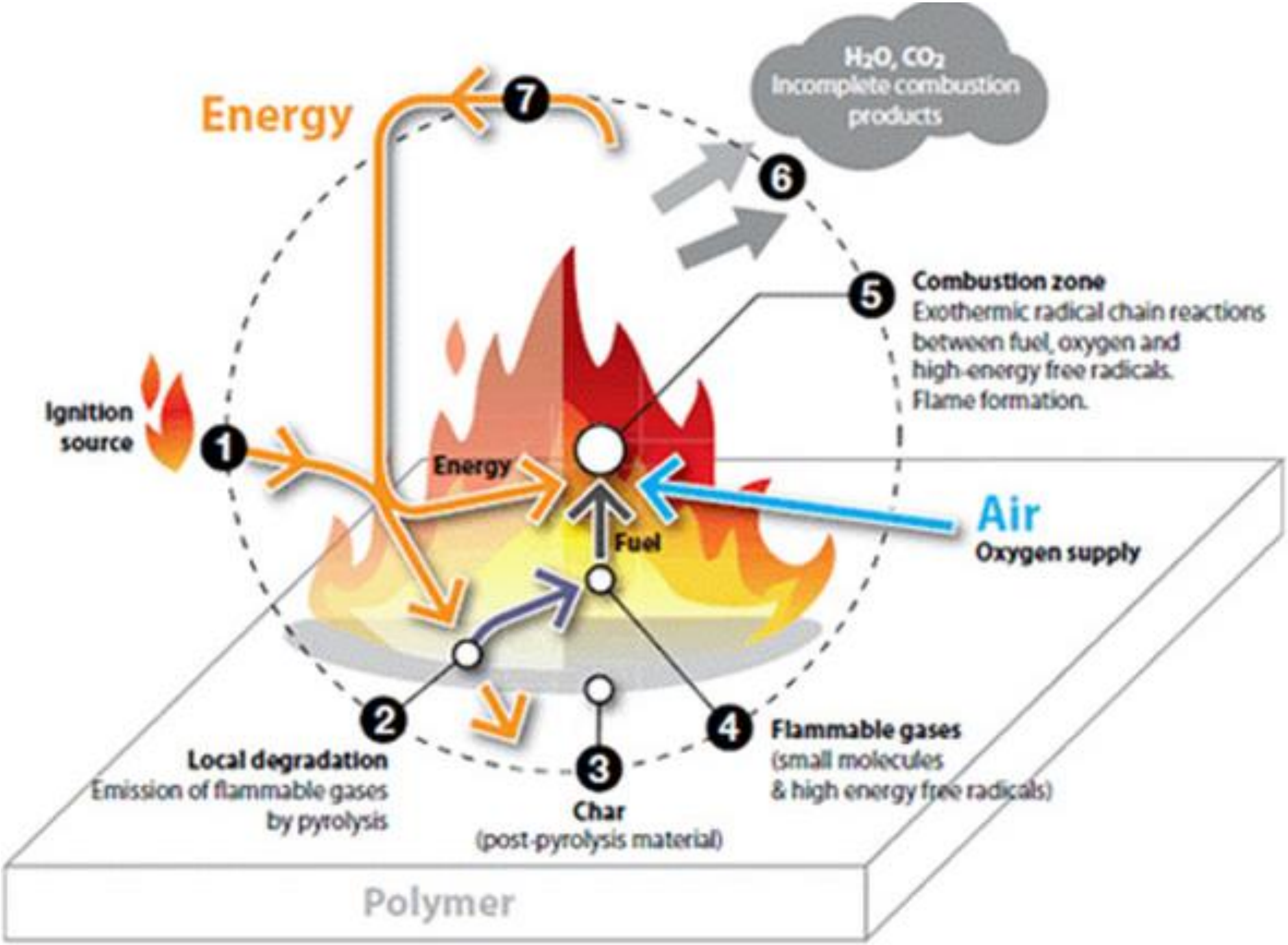


[E. Guillaume & al. Real-scale Fire Tests of One Bedroom Apartments with Regard to Tenability Assessment, Fire safety journal, 2015]

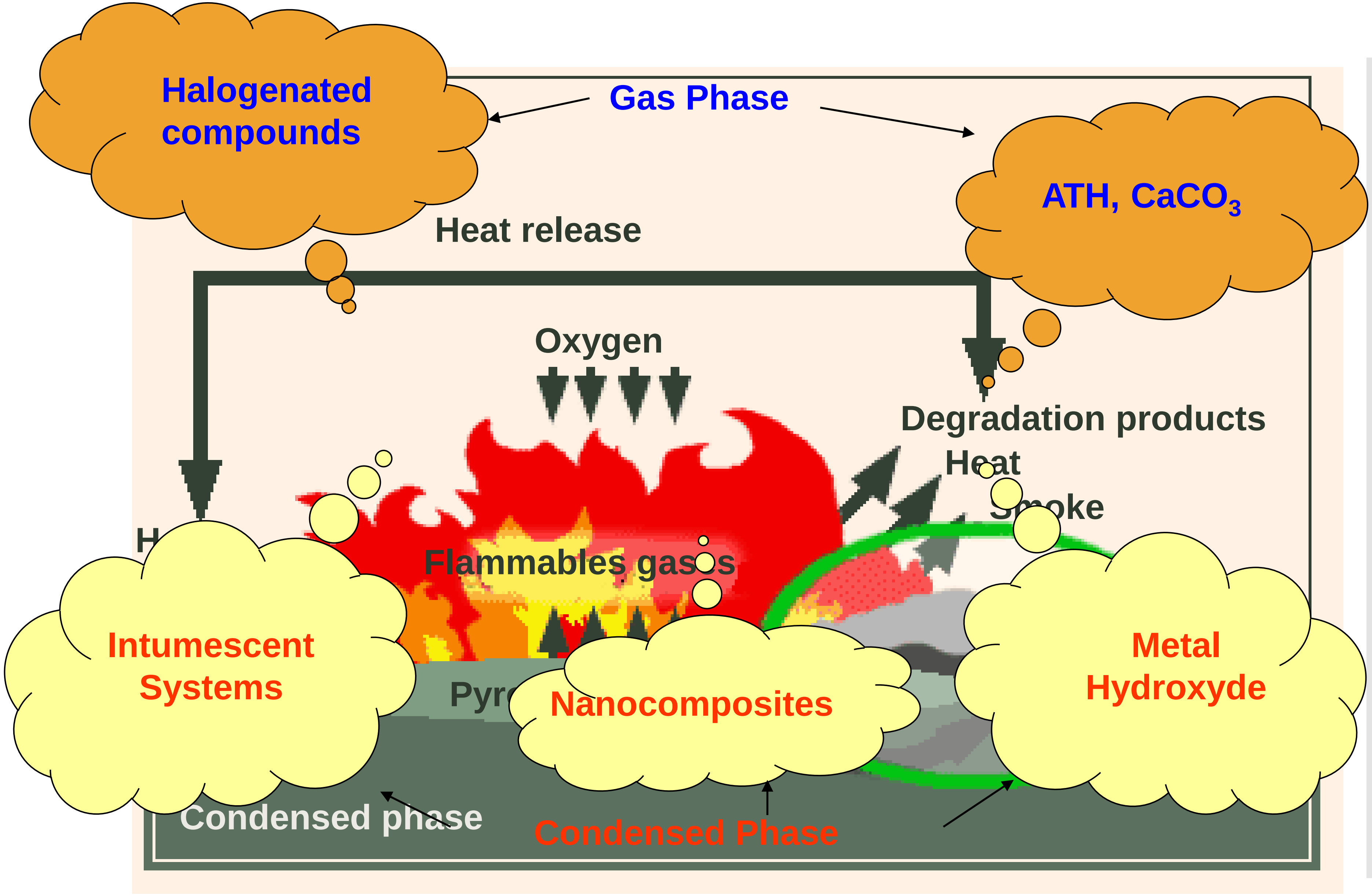
Fire Retardancy of Polymers

Combustion cycle

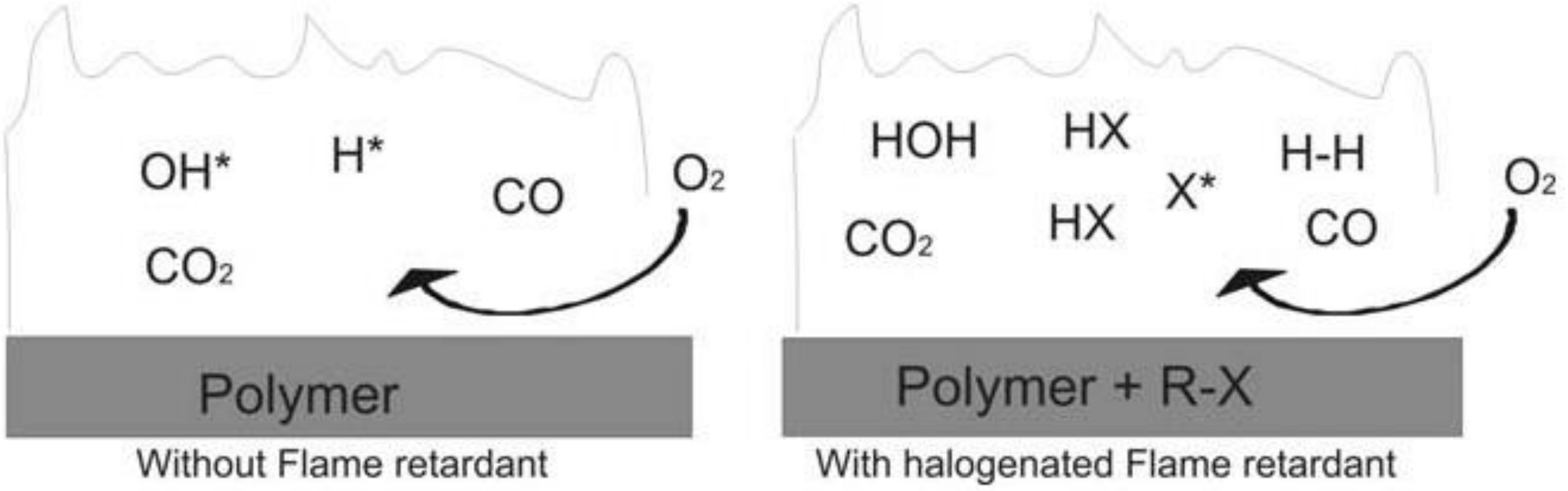
http://www.enteknomaterials.com



Flame retardancy



Fire Retardancy of Polymers : Halogenated Compounds



Toxicology and Applied Pharmacology 216(2):274-81 · November 2006

$$X^{\cdot} + RH \longrightarrow HX + R^{\cdot}$$

$$H^{\cdot} + HX \longrightarrow H_2 + X^{\cdot}$$

$$HO^{\cdot} + HX \longrightarrow H_2O + X^{\cdot}$$

$$Sb_2O_3 + 6\,HX \rightleftharpoons 2\,SbX_3 + 3\,H_20$$

$$SbX_3 + H^{\cdot} \rightleftharpoons SbX_2 + HX$$

$$SbX_2 + H^{\cdot} \rightleftharpoons SbX + HX$$

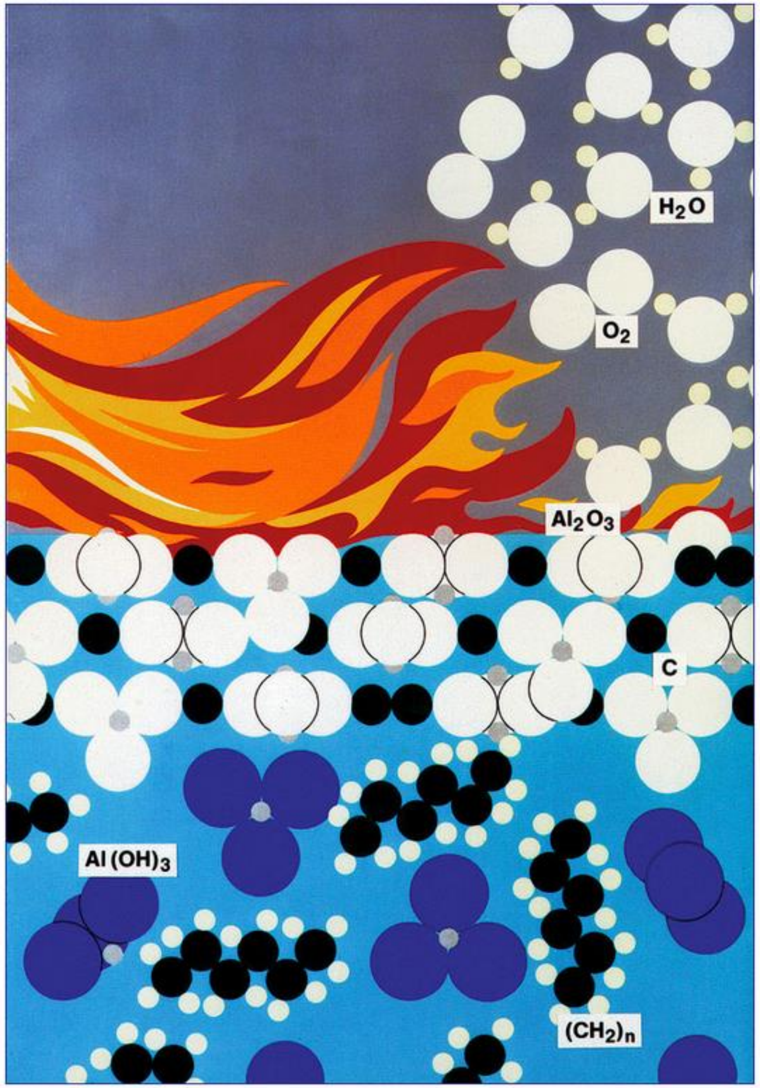
$$SbX + H^{\cdot} \rightleftharpoons Sb + HX$$

$$Sb + O^{\cdot} \rightleftharpoons SbO^{\cdot}$$

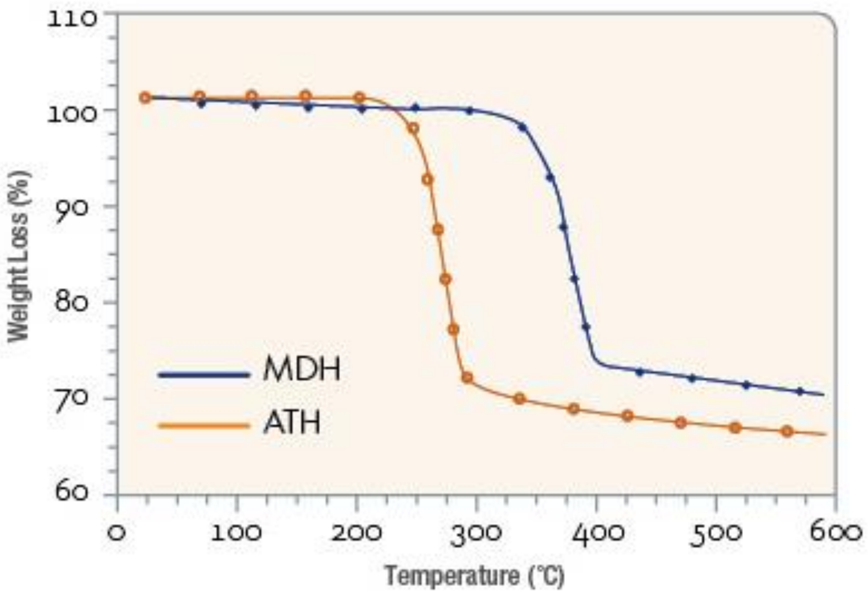
$$SbO^{\cdot} + H^{\cdot} \rightleftharpoons SbOH$$

$$SbOH + H^{\cdot} \rightleftharpoons SbO^{\cdot} + H_2$$

Fire Retardancy of Polymers : Metal Hydroxides



https://www.martinswerk.de/



https://www.hubermaterials.com/

$$2\,Al(OH)_3 \xrightarrow[+1050\ kJ/kg]{200\ °C} 3\,H_2O + Al_2O_3$$

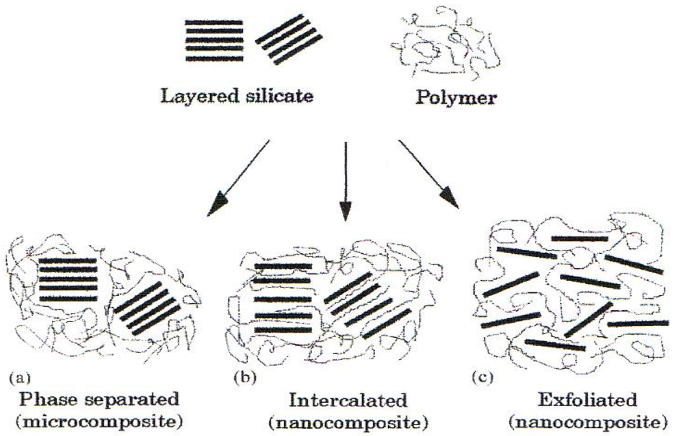
$$Mg(OH)_2 \xrightarrow[+1300\ kJ/kg]{300°C} H_2O + MgO$$

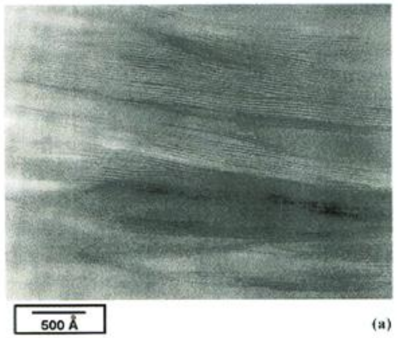
- « To intumesce » was used by the tragedian John Webster in the 16th century with two meanings
  - To grow and to increase in volume against the heat
  - To show an expanding effect by bubbling

- The result of this process is a *foamed charred layer* on the surface which protects the underlying material from the action of the heat flux or the flame

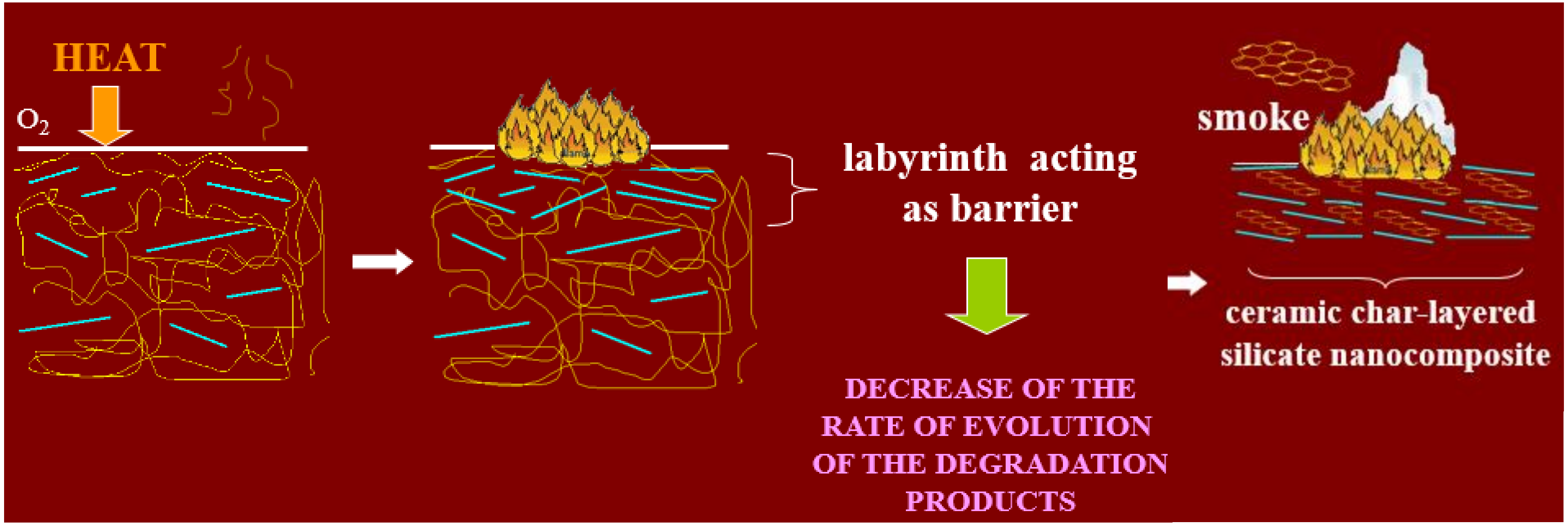
- Flame retarding polymers by intumescence is essentially a special case of a *condensed phase mechanism*

Fire Retardancy of Polymers : Nanocomposites





Intercalated PS**



**Doh et al.
Polym. Bull. 41:511-517 (1998)

Intumescent Polyurethane

Intumescent additive

Ammonium polyphosphate (APP)

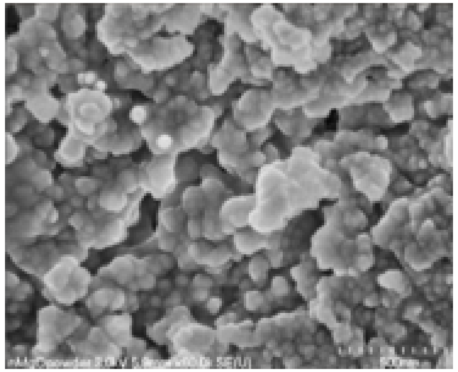
Improve fire properties by formation of a protective charred layer

Synergistic agent

Nano magnesium oxide (nano MgO)

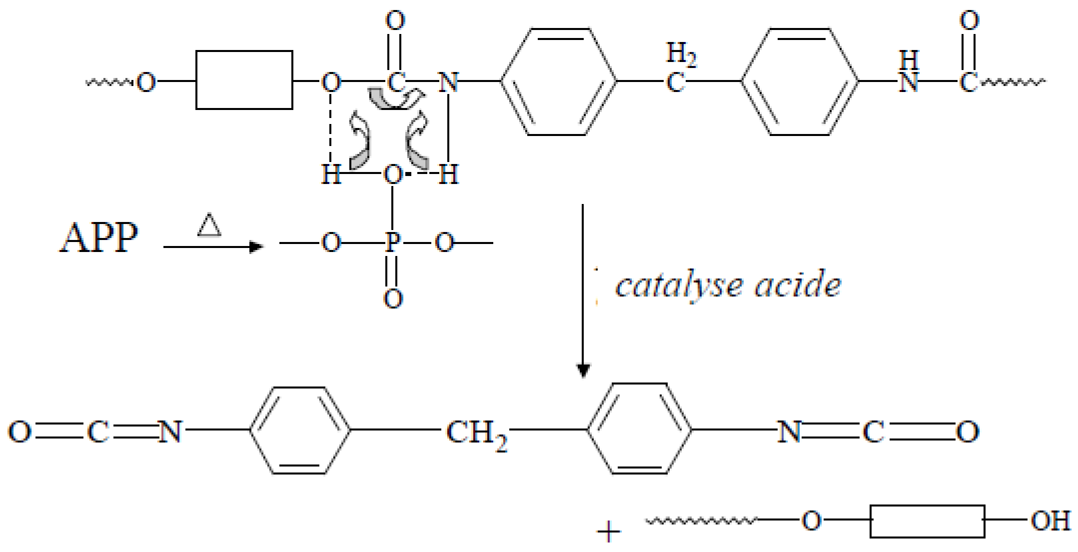
Better fire performances for the same additive loading

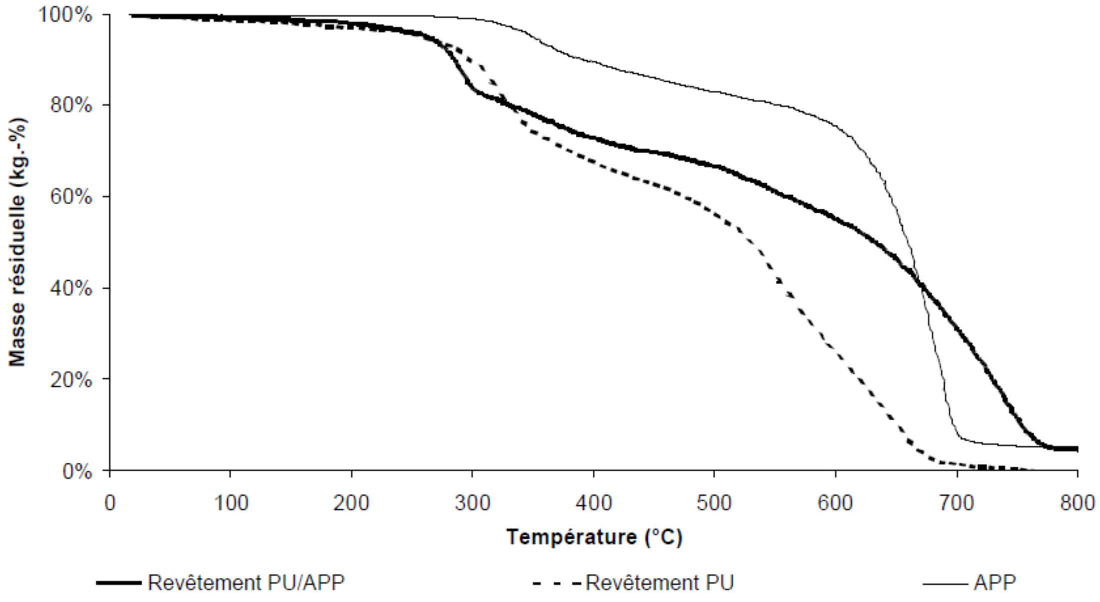
POLYURETHANE MATRIX (PU) = Polyol + isocyanate



PU with improved flame retardancy

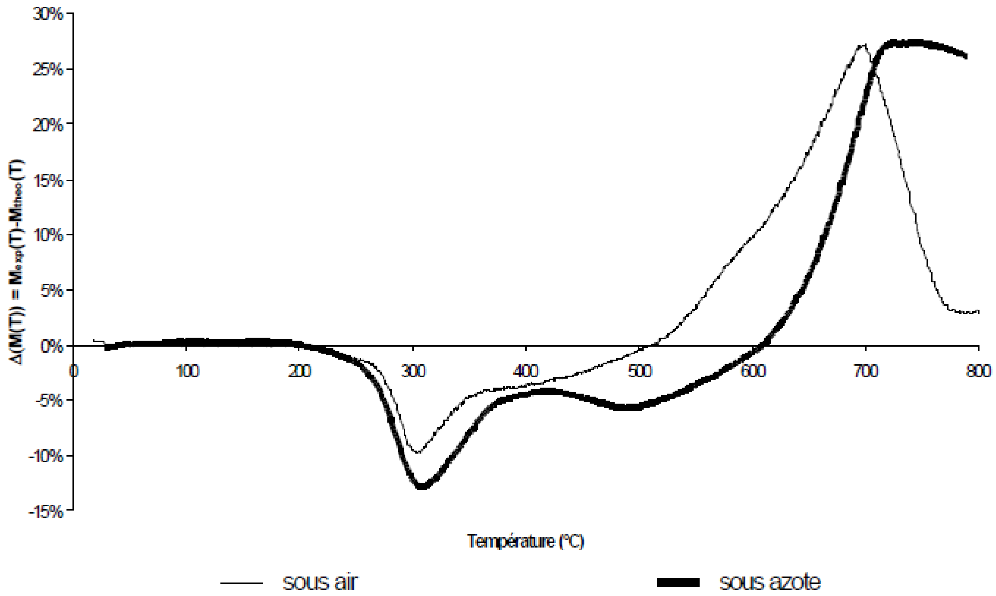
Intumescent Polyurethane





$$\Delta(M(T)) = M_{exp}(T) - M_{theo}(T)$$

$$M_{theo}(T) = (1-x) \times M_{poly}(T) + x \times M_{add}(T)$$

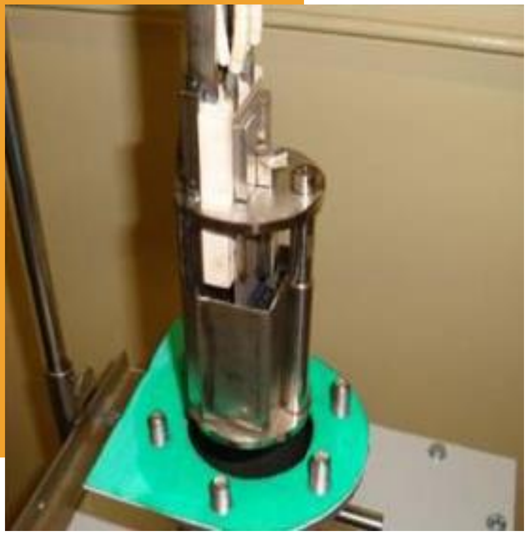


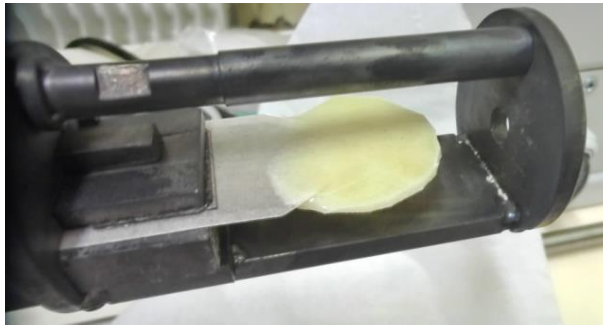
Intumescent Polyurethane

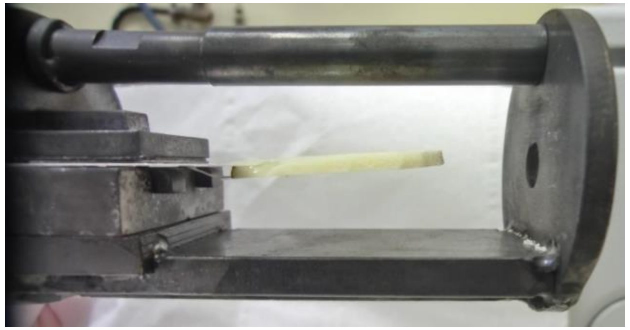
✓ Method to measure the heat conductivity of PU chars

- Transient plane source method
- Sensor in mica (resists at high T), r = 3.189 mm
- Measurement at **ambient T** and as a **function of T**

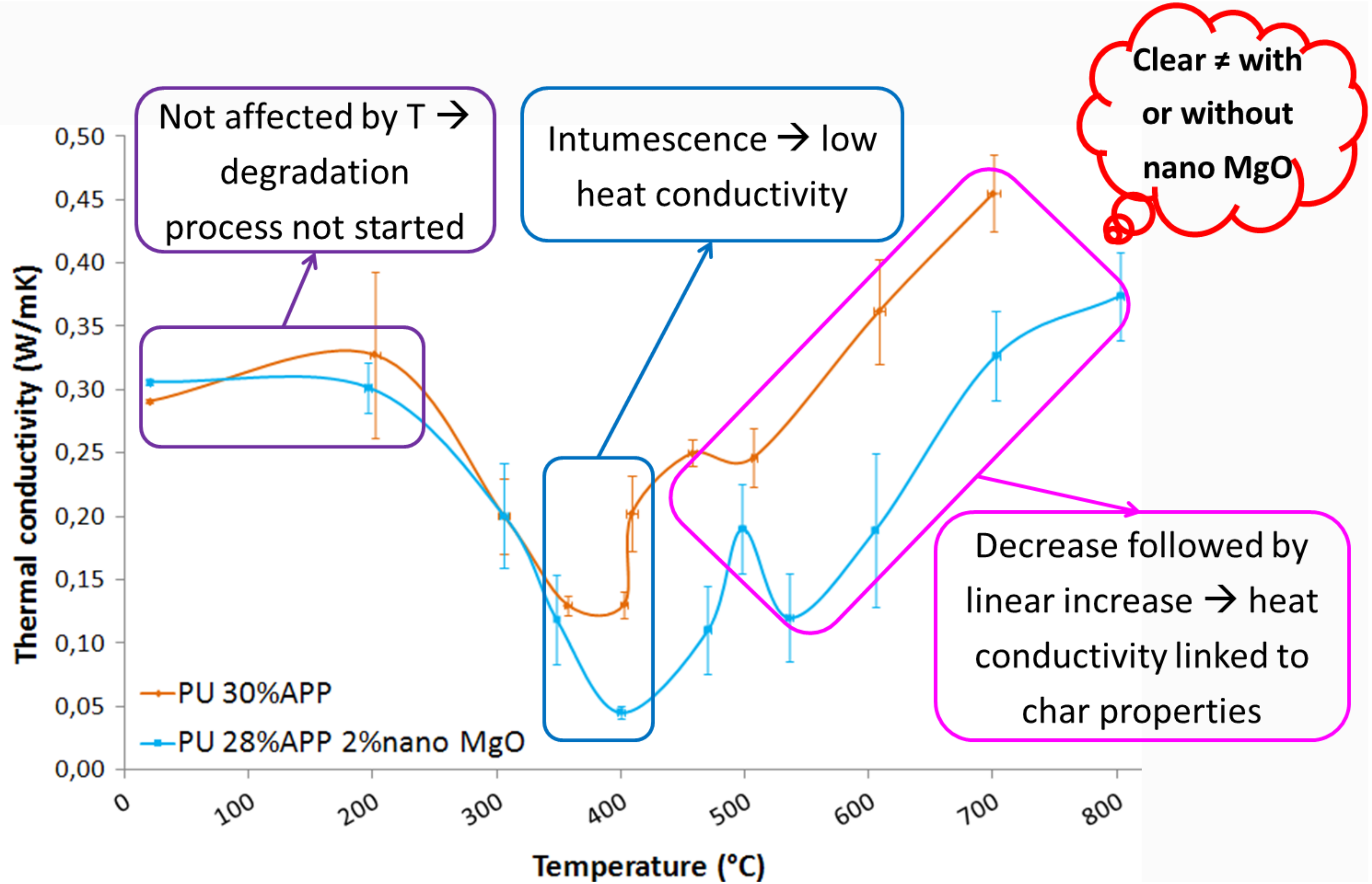
→ sample heated in a furnace controlled by the Hot Disk software





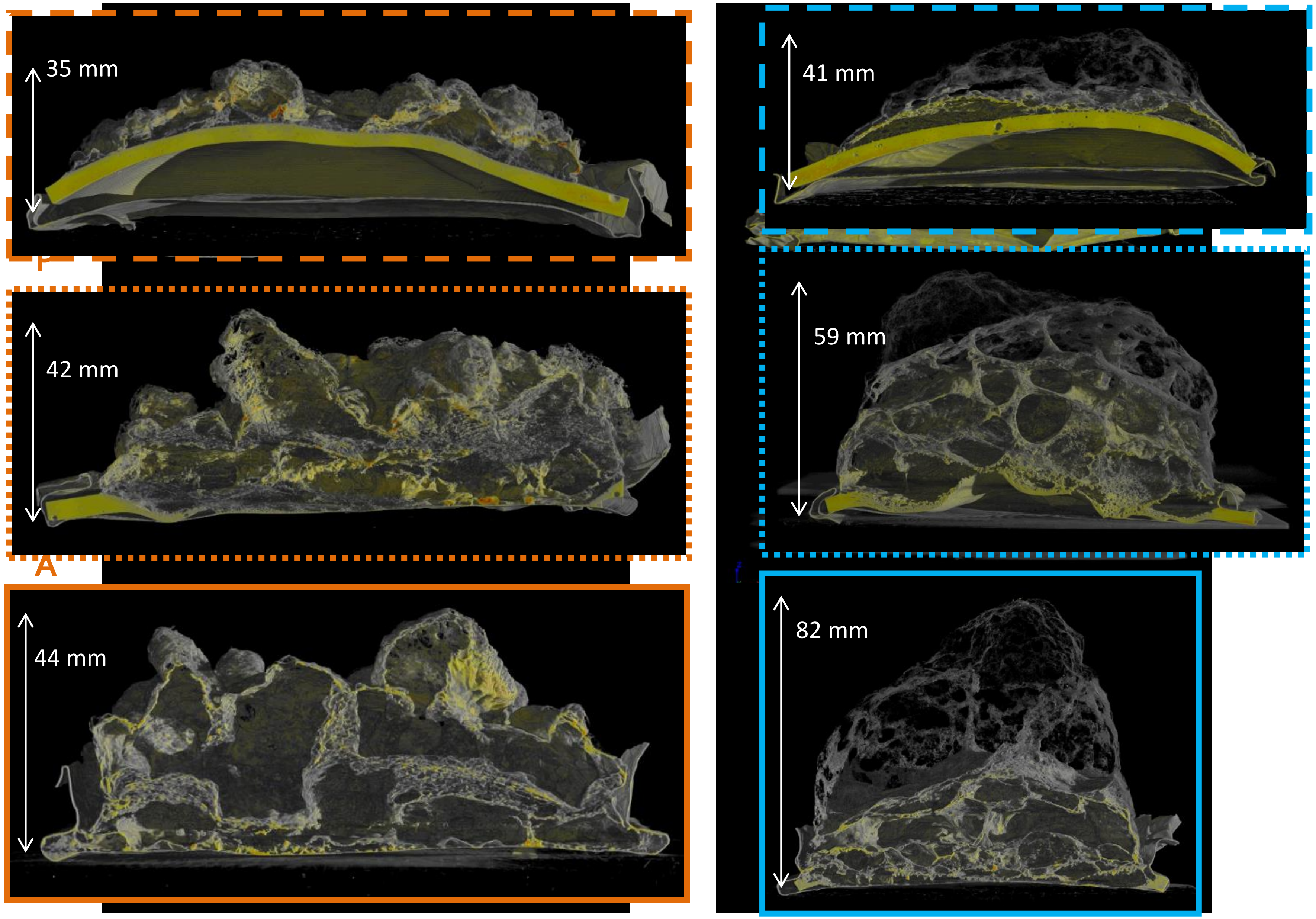


Intumescent Polyurethane



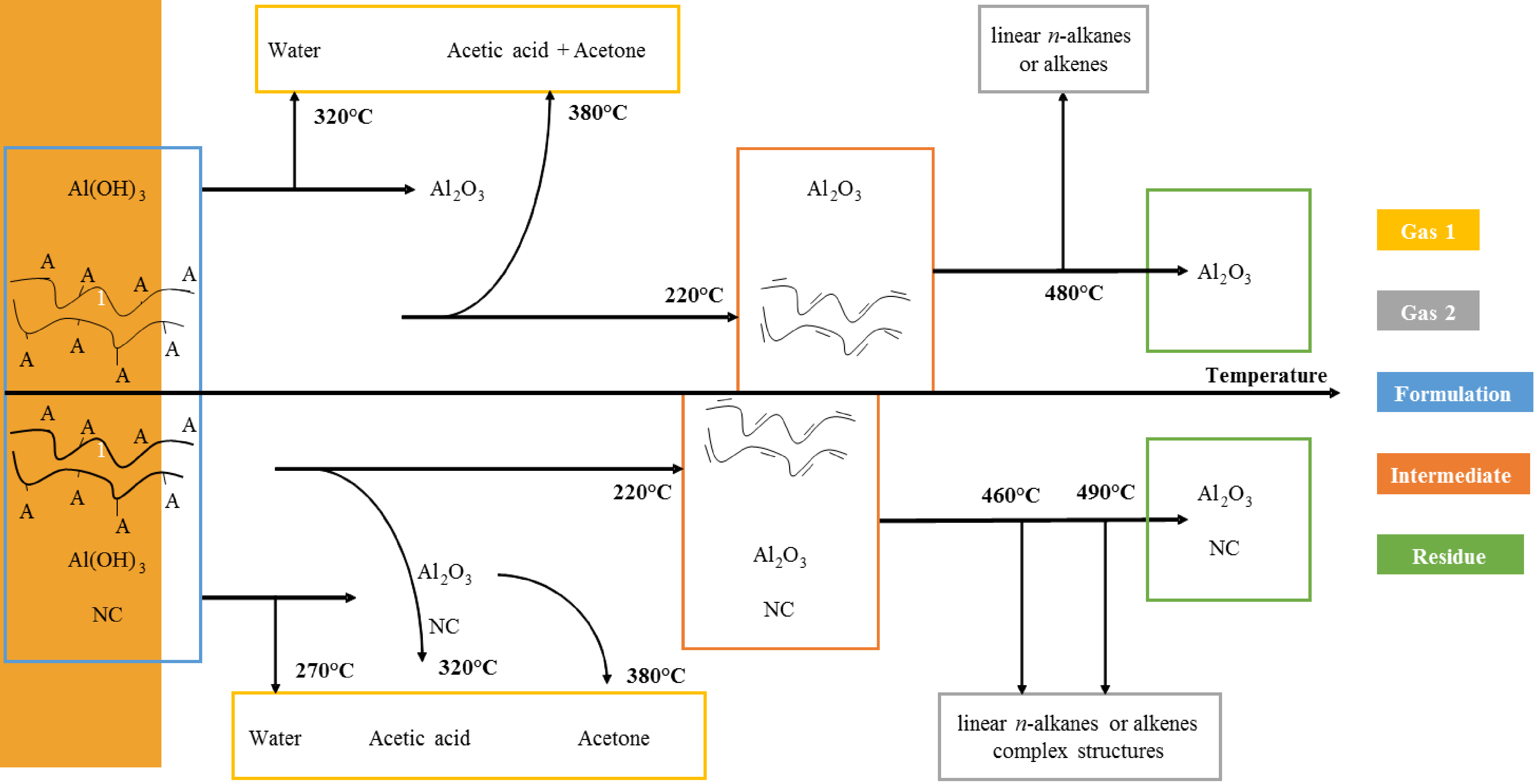
Intumescent Polyurethane

PU 30% APP

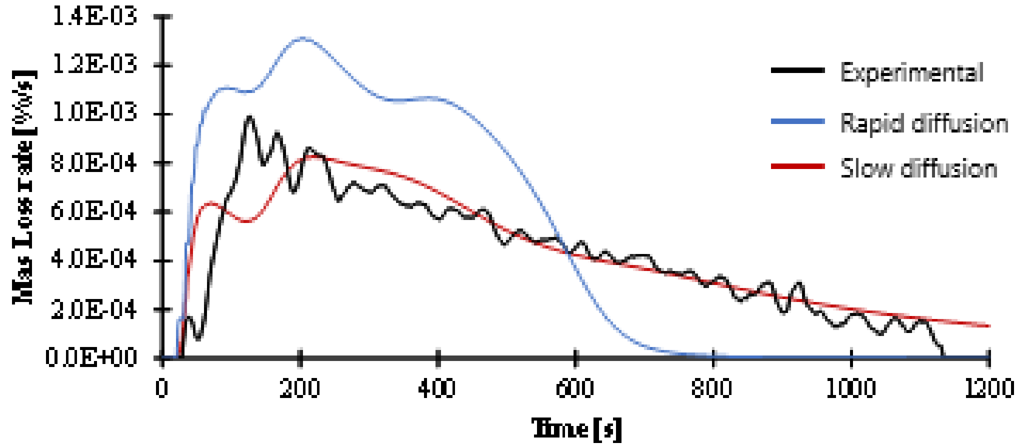


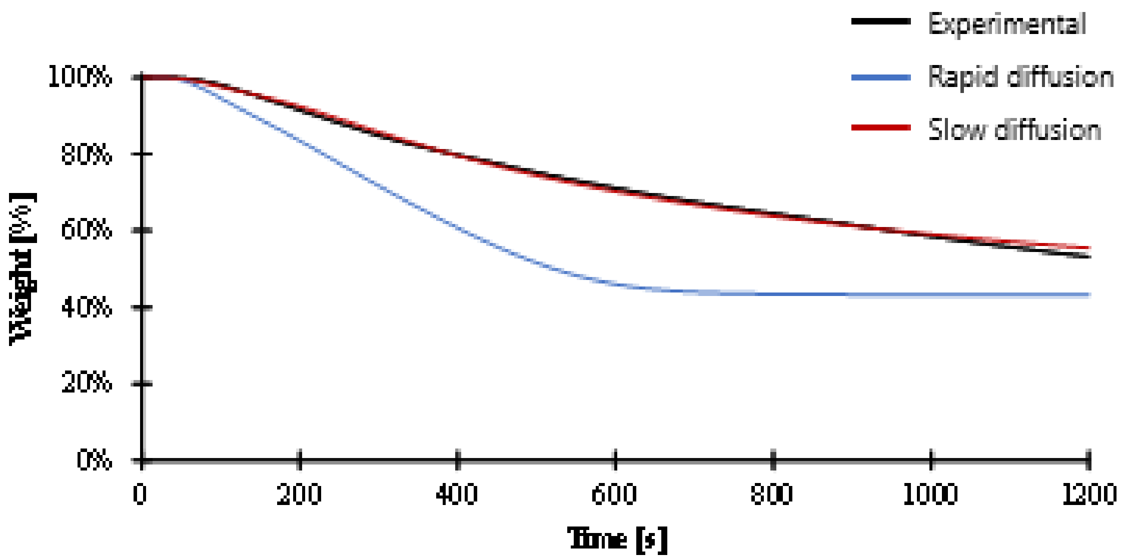
PU 28% APP 2% nMgO

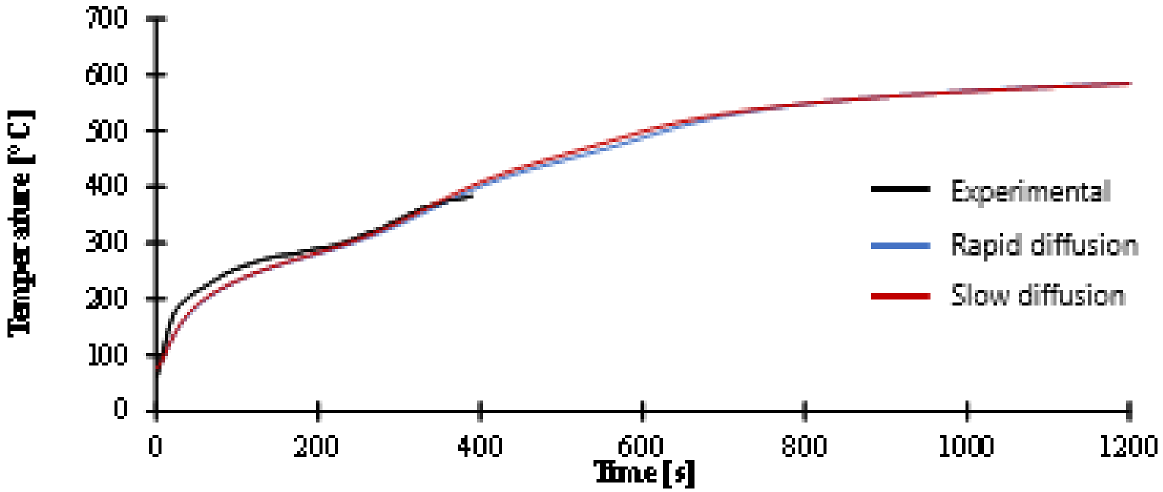
Degradation pathway of EVA/ATH and EVA/ATH/NC



Gasification experiments of EVA/ATH







- Good prediction of the temperature
- Prediction of MLR is achieved by tuning the mass transfer coefficient

→ Assumption of slow diffusion of the gases inside the decomposing materials

Conclusion and perspective

Conclusion and perspectives

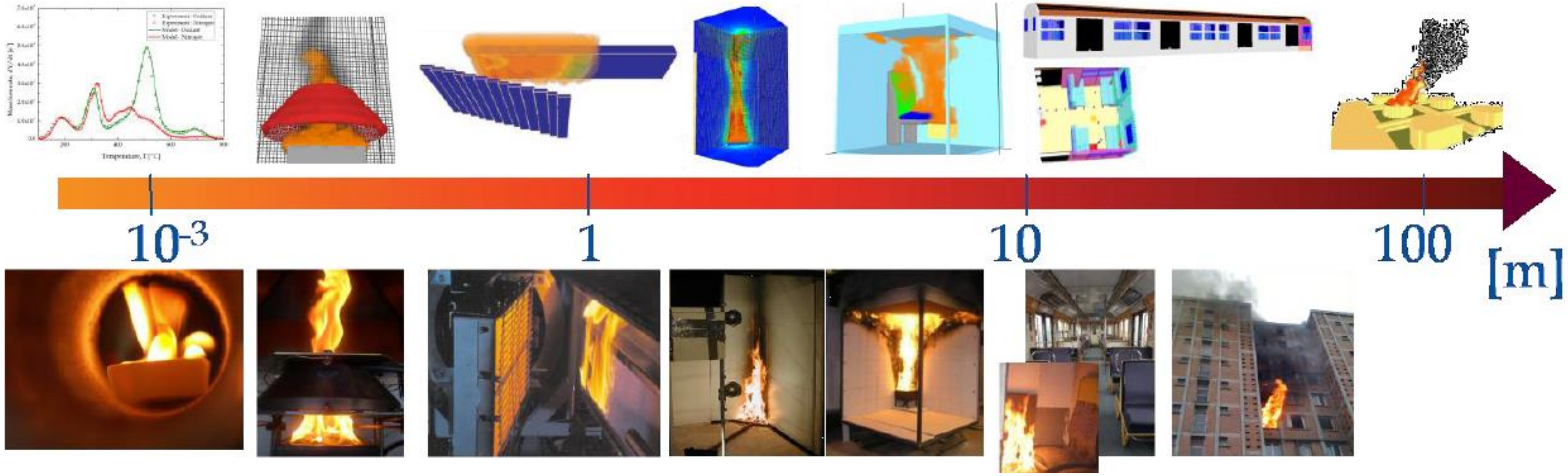
One says (to do not cited him… Pr Torero) that the research on thermal decomposition will even require Hundreds of years !!!!

So, just a perspective: WE NEED YOU !!!!!

ANNEXES

Focus: Experimental Investigations – Thermal decomposition

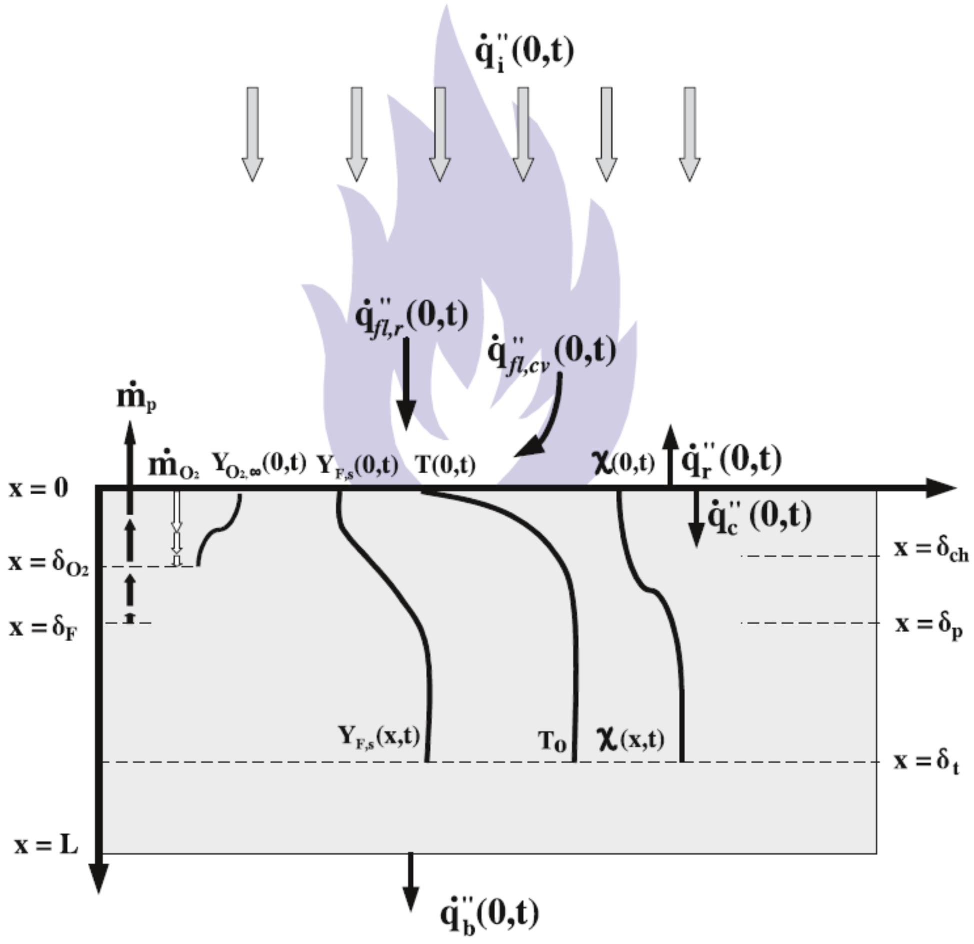
Multi scale approach



Matter scale | Material | Product | Large | Real | SCALE

Determination of a model of pyrolysis - Matter scale investigations

Mass loss rate (MLR)



$$\dot{m}_c'' = \frac{\dot{q}_i'' + \dot{q}_{fl}'' - \dot{q}_{rr}''}{\Delta h_g}$$

$\dot{q}_i''$ *Incident heat flux [kW.m-2]*

$\dot{q}_{fl}'' = \dot{q}_{fl,c}'' + \dot{q}_{fl,c}''$ *Heat flux from the flame [kW.m-2]*

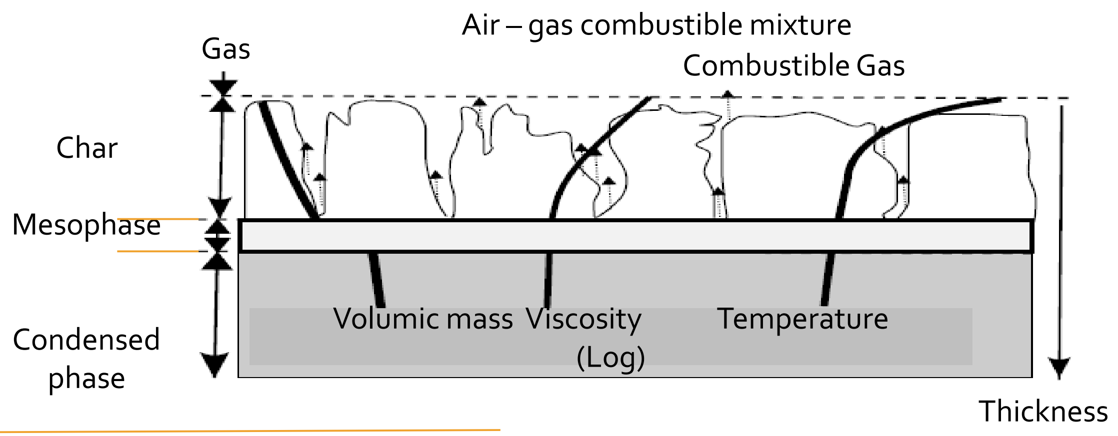
$\dot{q}_{rr}''$ *Radiative heat loss [kW.m-2]*

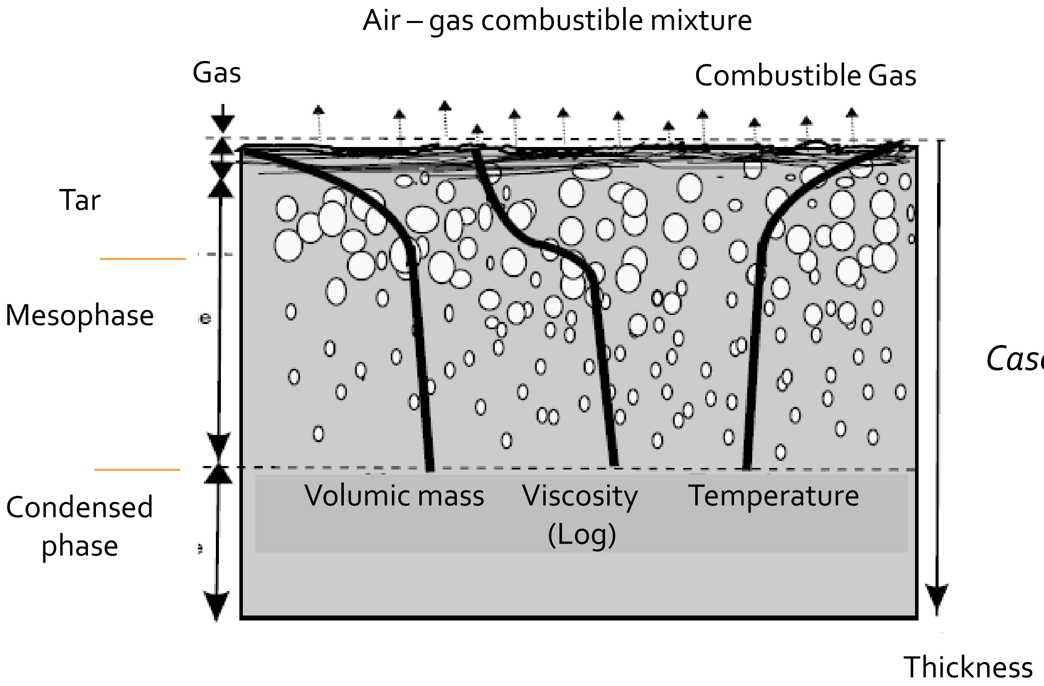
Δh_g *Enthalpy of gaseification [kJ.kg]*

Problems into the Condensed phase

Different kinds of solid polymers

Case of a charring material – ex: wood.





Case of a solid decomposition with liquid phase - tar